



Oxford Cambridge and RSA

Tuesday 13 May 2025 – Morning

AS Level Chemistry A

H032/01 Breadth in chemistry

Time allowed: 1 hour 30 minutes



You must have:

- the Data Sheet for Chemistry A

You can use:

- a scientific or graphical calculator
- an HB pencil



Please write clearly in black ink. **Do not write in the barcodes.**

Centre number

--	--	--	--	--

Candidate number

--	--	--	--

First name(s)

Last name

StemUp Model Solution

INSTRUCTIONS

- Use black ink. You can use an HB pencil, but only for graphs and diagrams.
- Write your answer to each question in the space provided. If you need extra space use the lined pages at the end of this booklet. The question numbers must be clearly shown.
- Answer **all** the questions.
- Where appropriate, your answer should be supported with working. Marks might be given for using a correct method, even if your answer is wrong.

INFORMATION

- The total mark for this paper is **70**.
- The marks for each question are shown in brackets [].
- This document has **24** pages.

ADVICE

- Read each question carefully before you start your answer.

2
SECTION A

You should spend a **maximum** of **25 minutes** on this section.

Write your answer to each question in the box provided.

- 1** The burette readings from a titration are shown below.

Initial reading / cm³	4.95
Final reading / cm³	23.10

The burette used has an uncertainty of 0.05 cm³ for each reading.

What is the percentage uncertainty of the resulting titre?

- A** 0.22%
B 0.28%
C 0.44%
D 0.55%

Your answer

D

[1]

- 2** How many p electrons does an atom of chlorine have?

- A** 5
B 7
C 11
D 17

Your answer

C

[1]

- 3** What is the number of hydrogen atoms in 6.8 g NH₃?

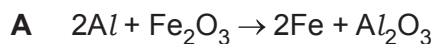
- A** 2.408×10^{23}
B 3.612×10^{23}
C 7.224×10^{23}
D 1.806×10^{24}

Your answer

C

[1]

4 Which reaction to extract iron (Fe) has the highest atom economy?



Your answer

D

[1]

5 Phosphorus is burnt in oxygen to produce a mixture of oxides.

Which oxide has the closest to 50% by mass of phosphorus?



Your answer

A

[1]

6 What is the oxidation number of S in $Na_2S_2O_3$?

A -4

B -2

C +2

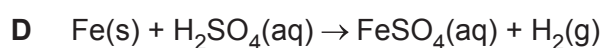
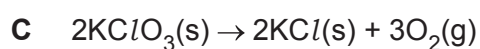
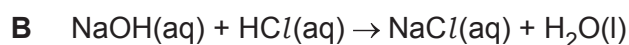
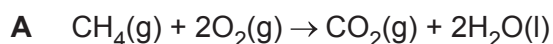
D +4

Your answer

C

[1]

7 Which of the following is **not** a redox reaction?

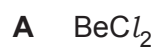


Your answer

B

[1]

8 Which compound has molecules with a trigonal planar shape?

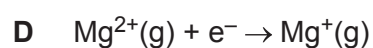
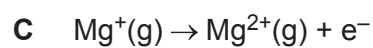
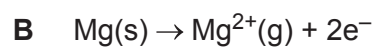
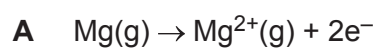


Your answer

B

[1]

9 Which equation describes the second ionisation energy of magnesium?



Your answer

C

[1]

10 Which molecule is polar?

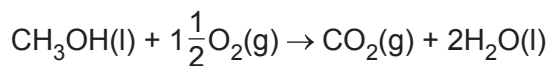


Your answer

A

[1]

- 11 The equation for the complete combustion of methanol, CH_3OH , is shown below.



The table shows standard enthalpy changes of formation, $\Delta_f H^\ominus$, in kJ mol^{-1} .

Substance	$\text{CH}_3\text{OH(l)}$	$\text{O}_2\text{(g)}$	$\text{CO}_2\text{(g)}$	$\text{H}_2\text{O(l)}$
$\Delta_f H^\ominus / \text{kJ mol}^{-1}$	-239	0	-393	-286

What is the enthalpy change of combustion of $\text{CH}_3\text{OH(l)}$, in kJ mol^{-1} ?

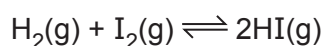
- A +726
B +440
C -440
D -726

Your answer

D

[1]

- 12 Hydrogen and iodine react in the reaction below.



Which row is correct when pressure is **increased**?

	Rate of reaction	Equilibrium yield of HI
A	Increases	Increases
B	Increases	No effect
C	Increases	Decreases
D	Decreases	No effect

Your answer

B

[1]

13 Which statement is **not** a sustainability benefit of an industrial process using a catalyst?

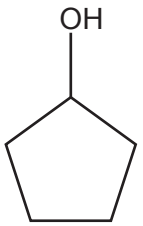
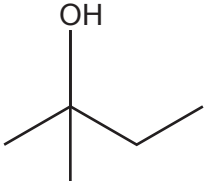
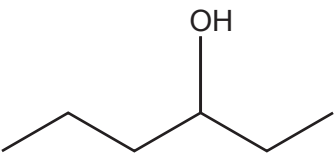
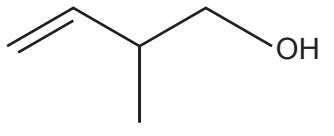
- A Lower temperatures can be used.
- B The position of equilibrium moves towards the products.
- C The reaction takes place in less time.
- D There is reduced energy demand.

Your answer

B

[1]

14 The skeletal formulae of four alcohols, **W**, **X**, **Y**, **Z**, are shown below.

W	
X	
Y	
Z	

Which pair of alcohols are structural isomers of each other?

- A **W** and **X**
- B **W** and **Y**
- C **W** and **Z**
- D **X** and **Y**

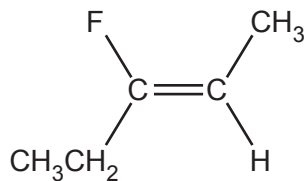
Your answer

C

[1]

15 A compound is shown below.

What is the systematic name of this compound?



- A *E*-3-fluoropent-2-ene
- B *E*-3-fluoropent-3-ene
- C *Z*-3-fluoropent-2-ene
- D *Z*-2-fluoropent-3-ene

Your answer

C

[1]

16 Which equation is a **propagation** step in the catalytic breakdown of ozone?

- A $\text{C}_2\text{F}_2\text{Cl}_2 \rightarrow \text{C}_2\text{F}_2\text{Cl}\cdot + \text{Cl}\cdot$
- B $\text{O}_3 + \text{O} \rightarrow 2\text{O}_2$
- C $\text{Cl}\cdot + \text{Cl}\cdot \rightarrow \text{Cl}_2$
- D $\text{ClO}\cdot + \text{O} \rightarrow \text{Cl}\cdot + \text{O}_2$

Your answer

D

[1]

17 How many structural isomers of $\text{C}_5\text{H}_{12}\text{O}$ are secondary alcohols?

- A 1
- B 2
- C 3
- D 4

Your answer

C

[1]

- 18 The hydrolysis of 1-iodopropane by aqueous potassium hydroxide takes place at a faster rate than 1-bromopropane.

Which statement explains the different rates of reaction?

- A Iodine is less reactive than bromine.
- B The C–I bond enthalpy is less than the C–Br bond enthalpy.
- C The C–I bond is less polar than a C–Br bond.
- D The C–I bond length is shorter than the C–Br bond length.

Your answer

B

[1]

- 19 In a mass spectrum, which alkene is likely to have a fragment ion at $m/z = 29$?

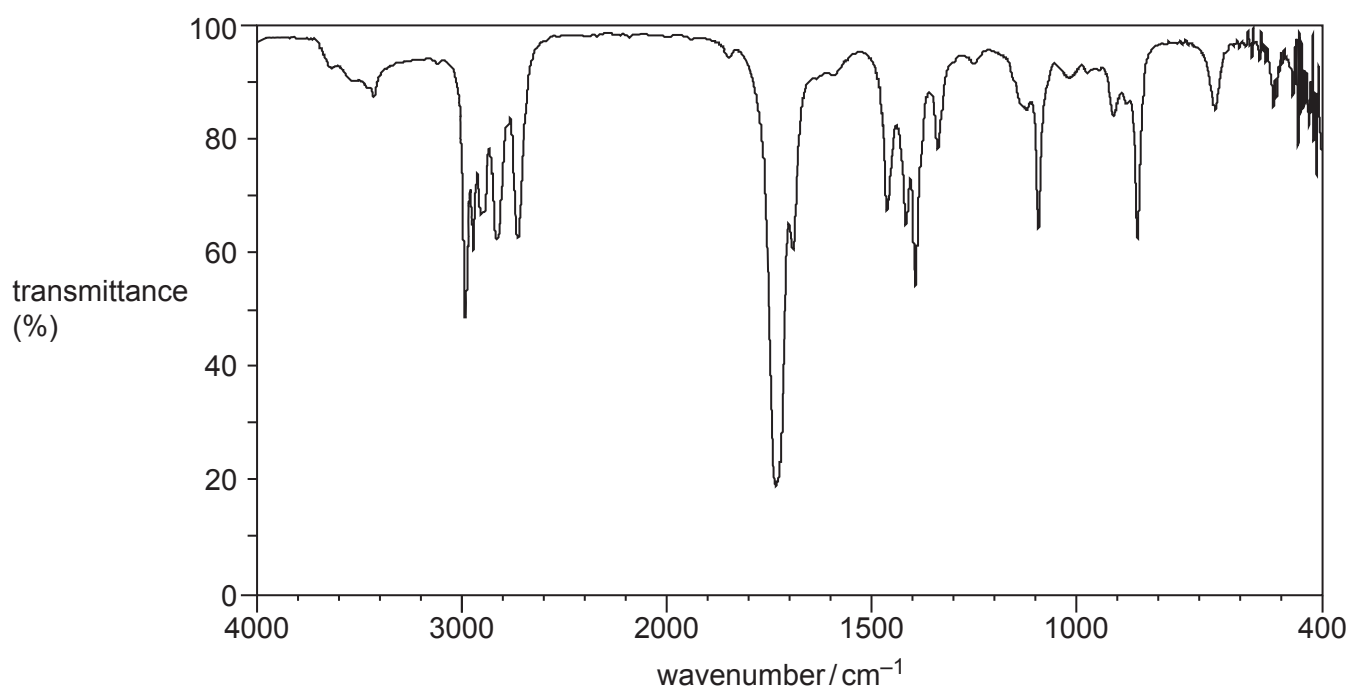
- A $\text{H}_2\text{C}=\text{CHCH}=\text{CH}_2$
- B $\text{H}_2\text{C}=\text{C}(\text{CH}_3)_2$
- C $\text{H}_3\text{CCH}=\text{CHCH}_3$
- D $\text{H}_3\text{CCH}=\text{CHCH}_2\text{CH}_3$

Your answer

D

[1]

20 The IR spectrum of a compound is shown below.



Which compound could have produced the IR spectrum?

- A $\text{CH}_3\text{CH}=\text{CH}_2$
- B $\text{CH}_3\text{CH}_2\text{CHO}$
- C $\text{CH}_3\text{CH}_2\text{CH}_2\text{OH}$
- D $\text{CH}_3\text{CH}_2\text{COOH}$

Your answer

B

[1]

10
SECTION B

21 This question is about atomic structure and isotopes of two elements in Period 3.

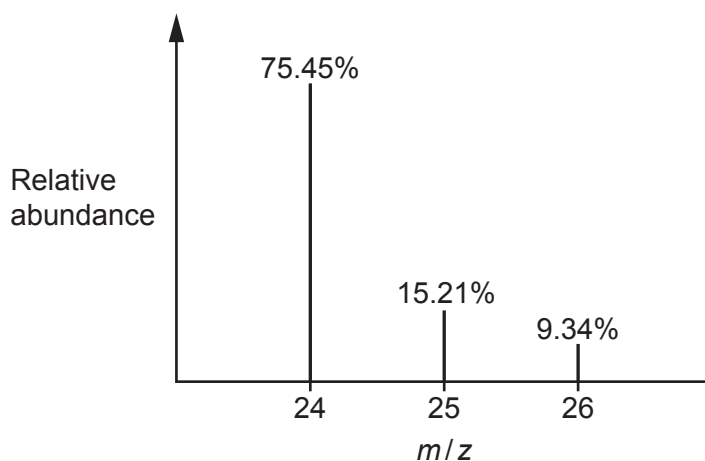
(a) This part of the question is about sulfur.

Complete the table below.

Species	Protons	Neutrons	Electrons
^{34}S	16	18	16
$^{32}\text{S}^{2-}$	16	16	18

[2]

(b) The mass spectrum of a sample of magnesium is shown below.



Calculate the relative atomic mass of magnesium in the sample.

Give your answer to 2 decimal places.

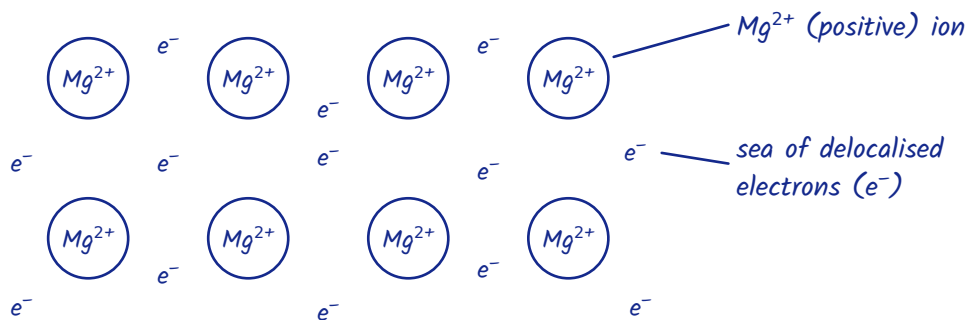
$$\begin{aligned}
 A_r &= (75.45 \times 24 + 15.21 \times 25 + 9.34 \times 26) \div 100 \\
 &= (1810.8 + 380.25 + 242.84) \div 100 \\
 &= 2433.89 \div 100 = 24.3389
 \end{aligned}$$

Relative atomic mass = ...24.34..... [2]

(c) Magnesium has a giant metallic structure.

- Draw a labelled diagram to show the bonding and structure in magnesium. Include the correct charges on the metal particles in your diagram.
- Explain how the structure of magnesium allows it to conduct electricity.

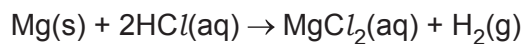
Diagram



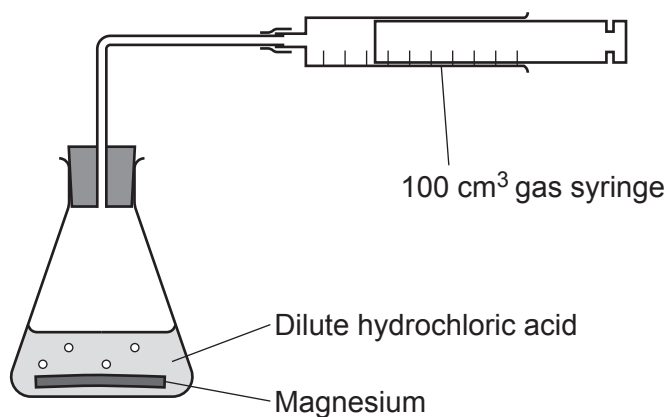
Explanation *The delocalised electrons are mobile, so they can move (flow) through the lattice and carry charge when a potential difference is applied.*

[3]

(d) Magnesium reacts with dilute hydrochloric acid, as shown in the equation.



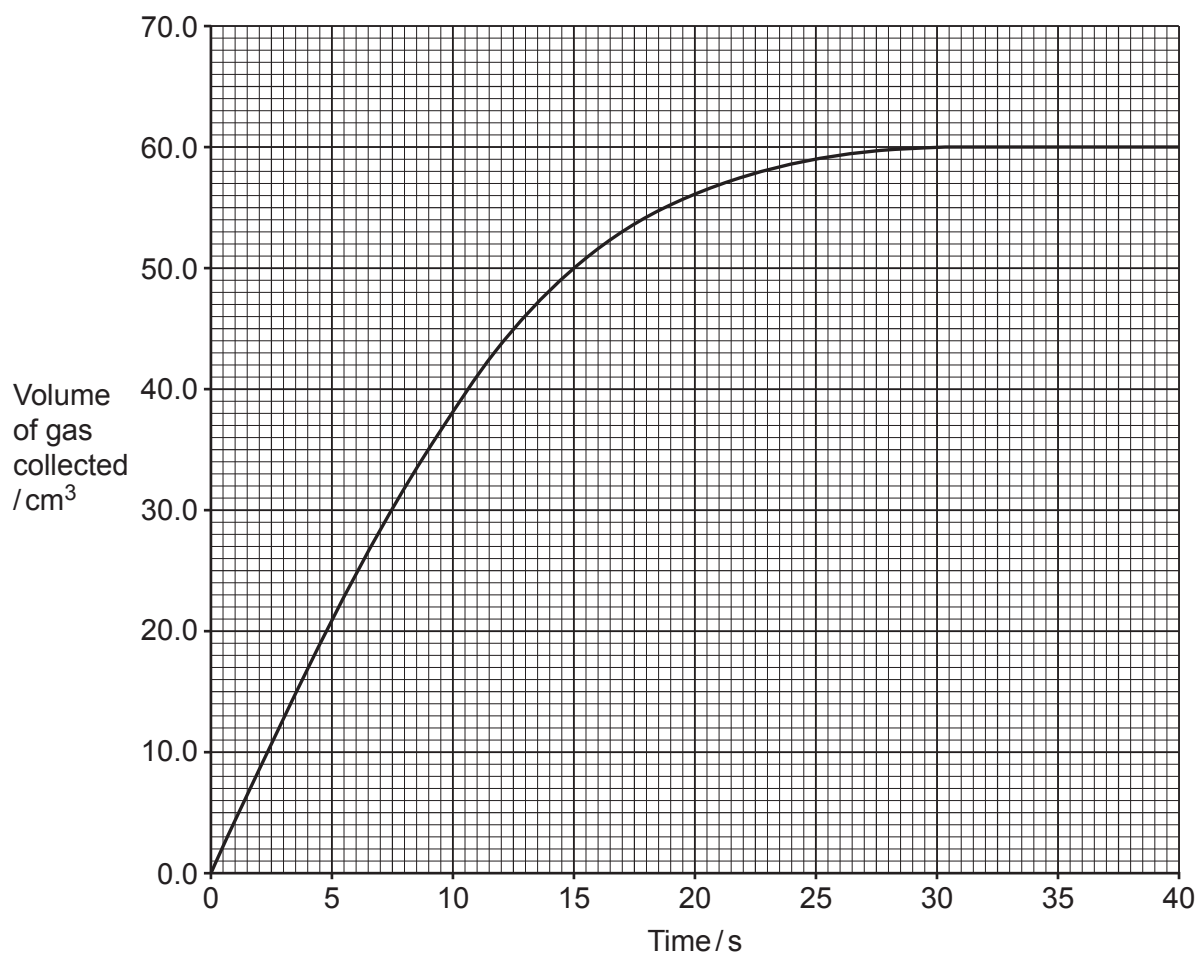
A student reacts an excess of Mg with 24.5 cm^3 HCl(aq) using the apparatus below.



The student measures the volume of gas collected over time. All gas volumes are measured at RTP.

A graph of the student's results is shown in **Fig. 21.1** below.

Fig. 21.1



- (i) Calculate the concentration of HCl that the student used.

Give your answer to 2 significant figures and in standard form.

From Fig. 21.1, final volume of $H_2 = 60.0 \text{ cm}^3$

$$n(H_2) = 60.0 \div 24000 = 2.50 \times 10^{-3} \text{ mol}$$

$$n(HCl) = 2 \times 2.50 \times 10^{-3} = 5.00 \times 10^{-3} \text{ mol (2 : 1 ratio)}$$

$$c(HCl) = 5.00 \times 10^{-3} \div (24.5 \div 1000)$$

$$= 5.00 \times 10^{-3} \div 0.0245 = 0.204 \text{ mol dm}^{-3}$$

Concentration of HCl = ... 2.0×10^{-1} mol dm^{-3} [3]

- (ii) Explain the shape of the graph in **Fig. 21.1**, in terms of concentration, collisions and time.

Concentration: as the reaction proceeds the HCl is used up, so the acid

concentration decreases (fewer particles in a given volume).

Collisions: so there are fewer collisions per second (less frequent collisions), and

the rate decreases, so the gradient of the graph decreases with time.

Time: after about 30 s the graph plateaus because the reaction stops: all of the

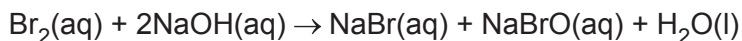
HCl (the limiting reagent) has reacted, as the Mg is in excess.

..... [3]

22 This question is about halogens and their compounds.

- (a) An aqueous solution of bromine reacts with cold aqueous sodium hydroxide in a disproportionation reaction.

The equation is shown below.



Explain what is meant by the term **disproportionation**.

Use the equation above and oxidation numbers.

Disproportionation is the oxidation and reduction of the same element (bromine) in the same reaction.

Br is oxidised from 0 in Br₂ to +1 in NaBrO.

Br is reduced from 0 in Br₂ to -1 in NaBr.

[3]

- (b) An aqueous solution of chlorine also reacts with cold, aqueous sodium hydroxide. The reaction is more vigorous than the reaction of Br₂(aq) with NaOH(aq) in (a).

- (i) Explain why chlorine reacts more vigorously than bromine.

Chlorine has a smaller atomic radius than bromine (fewer shells, less shielding).

So the nuclear attraction on the outer electrons (and on an incoming electron) is greater in Cl than in Br.

So it is easier for chlorine to gain an electron, i.e. chlorine is more easily reduced, so it is a stronger oxidising agent and reacts more vigorously.

[3]

- (ii) Chlorine is used in water treatment.

State **one** disadvantage of treating water with chlorine.

Chlorine is toxic, and it can react with organic matter to form chlorinated hydrocarbons, which are carcinogenic. [1]

- (c) A group of students is provided with aqueous solutions of sodium chloride, sodium bromide and sodium iodide in three unlabelled bottles.

Explain how the students can identify which bottle contains each aqueous solution using precipitation reactions and the solubility of any precipitate formed.

Include an **ionic** equation in your answer.

Add aqueous silver nitrate, $\text{AgNO}_3(\text{aq})$, to a sample from each bottle.

Chloride gives a white precipitate, bromide gives a cream precipitate and iodide gives a yellow precipitate.

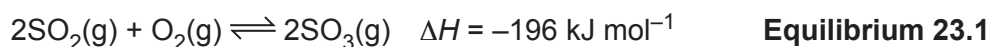
$\text{Ag}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{AgCl}(\text{s})$ ($\text{Ag}^+ + \text{X}^- \rightarrow \text{AgX}$)

Then add aqueous ammonia to each precipitate: AgCl dissolves in dilute $\text{NH}_3(\text{aq})$, AgBr dissolves only in concentrated $\text{NH}_3(\text{aq})$, and AgI does not dissolve even in concentrated $\text{NH}_3(\text{aq})$.

[4]

23 This question is about equilibrium and bond enthalpies.

The chemical industry manufactures sulfur trioxide, SO_3 , by reacting sulfur dioxide, SO_2 , and oxygen. This is a reversible reaction which can reach dynamic equilibrium.



(a) State **two** features of a **dynamic equilibrium**.

- 1 *The rate of the forward reaction equals the rate of the reverse reaction.*
- 2 *The concentrations of reactants and products do not change (are constant), in a closed system.*

[2]

(b) The displayed formulae of SO_2 and SO_3 are shown below.



The bond enthalpies for the $\text{S}=\text{O}$ bond in SO_2 and SO_3 are **not** the same.

The table shows bond enthalpies.

Bond	Bond enthalpy / kJ mol^{-1}
$\text{O}=\text{O}$	+494
$\text{S}=\text{O}$ in SO_2	+522

Calculate the $\text{S}=\text{O}$ bond enthalpy in SO_3 .

Use the information in the table and **Equilibrium 23.1**.

$$\text{bonds broken: } (4 \times \text{S}=\text{O in SO}_2) + (1 \times \text{O}=\text{O})$$

$$= (4 \times 522) + 494 = 2088 + 494 = 2582 \text{ kJ}$$

$$\Delta H = \text{bonds broken} - \text{bonds formed, so}$$

$$\text{bonds formed } (6 \times \text{S}=\text{O in SO}_3) = 2582 - (-196) = 2582 + 196$$

$$= 2778 \text{ kJ}$$

$$\text{S}=\text{O in SO}_3 = 2778 \div 6 = +463 \text{ kJ mol}^{-1}$$

$$\text{Bond enthalpy of S}=\text{O in SO}_3 = \dots\dots\dots +463 \dots\dots\dots \text{ kJ mol}^{-1} \quad [3]$$

(c) The table shows equilibrium concentrations for **Equilibrium 23.1**.

$[\text{SO}_2(\text{g})]/\text{mol dm}^{-3}$	$[\text{O}_2(\text{g})]/\text{mol dm}^{-3}$	$[\text{SO}_3(\text{g})]/\text{mol dm}^{-3}$
25.0	108	42.0

For **Equilibrium 23.1**

- write the expression for the equilibrium constant, K_c
- calculate K_c to 3 significant figures and in standard form.

$$K_c = \frac{[\text{SO}_3(\text{g})]^2}{[\text{SO}_2(\text{g})]^2 [\text{O}_2(\text{g})]}$$

$$K_c = \frac{42.0^2}{25.0^2 \times 108} = 1764 \div 67500 = 0.026133\ldots$$

$$= 2.61 \times 10^{-2} \text{ dm}^3 \text{ mol}^{-1}$$

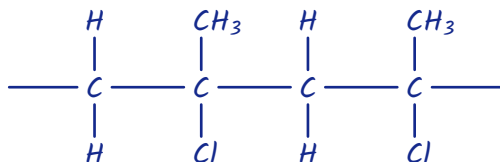
$$K_c = \dots\dots\dots 2.61 \times 10^{-2} \dots\dots\dots \text{ dm}^3 \text{ mol}^{-1}$$

[3]

24 This question is about alkenes.

(a) 2-Chloropropene can form an addition polymer.

(i) Draw **two** repeat units of the polymer formed by 2-chloropropene.



[1]

(ii) 2-Chloropropene can be disposed of by combustion. This process releases the greenhouse gas carbon dioxide.

State **one** other problem with this method of disposal.

Toxic hydrogen chloride (HCl) gas is formed, as the polymer contains chlorine.

.....

 [1]

(b) Pentene, C_5H_{10} , can undergo complete combustion to produce the same products as the combustion of an alkane.

Write the equation for the complete combustion of C_5H_{10} .

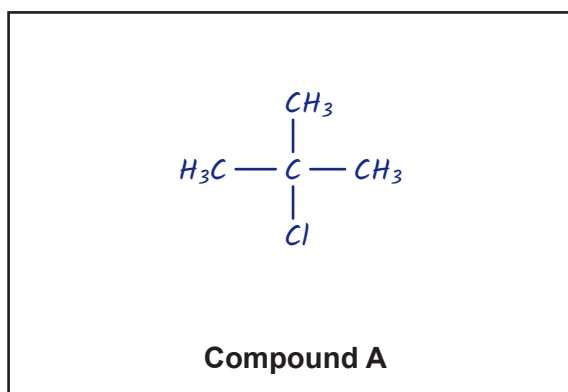
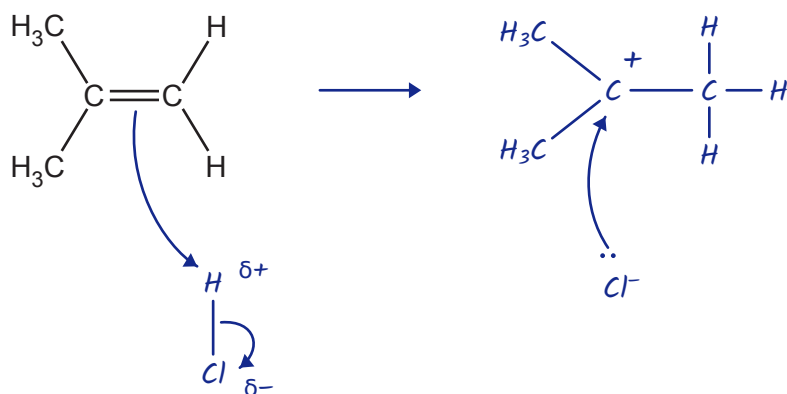
$C_5H_{10} + 7\frac{1}{2}O_2 \rightarrow 5CO_2 + 5H_2O$ [1]

(c) Methylpropene, $(\text{CH}_3)_2\text{C}=\text{CH}_2$, reacts with HCl to form a mixture of two products. Compound **A** is the major product.

(i) Outline the mechanism for the formation of the major product, compound **A**.

The structure of $(\text{CH}_3)_2\text{C}=\text{CH}_2$ has been provided.

Include curly arrows and relevant dipoles, the structure of compound **A** and the name of the mechanism.



Name of mechanism electrophilic addition

[5]

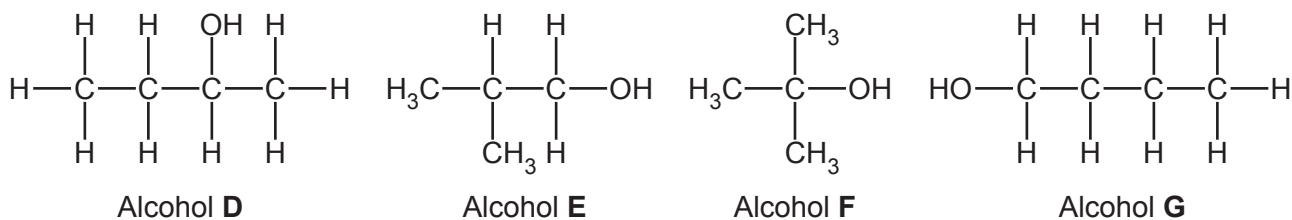
(ii) Explain why compound **A** is the major product.

The intermediate in its formation is a tertiary carbocation, which is more stable than the primary carbocation that would lead to the minor product.

[1]

25 This question is about alcohols.

Alcohols **D**, **E**, **F** and **G**, shown below, are structural isomers of $C_4H_{10}O$.



(a) Alcohols **D**, **E**, **F** and **G** are each heated under reflux with $H^+/Cr_2O_7^{2-}$.

Draw the structure of the organic product that is formed from each alcohol.

If there is no reaction, state 'No reaction'.

Alcohol	Organic product
D	$ \begin{array}{ccccccc} & H & & O & & H & H \\ & & & & & & \\ H - & C & - & C & - & C & - C - H \\ & & & & & & \\ & H & & & & H & H \end{array} $
E	$ \begin{array}{ccccc} & H & & O & \\ & & & & \\ H_3C - & C & - & C & - OH \\ & & & & \\ & CH_3 & & & \end{array} $
F	No reaction
G	$ \begin{array}{ccccccc} & H & & H & & H & O \\ & & & & & & \\ H - & C & - & C & - & C & - C - OH \\ & & & & & & \\ & H & & H & & H & \end{array} $

[3]

(b) Alcohols **D**, **E**, **F** and **G** are each heated with concentrated H_2SO_4 , which acts as an acid catalyst.

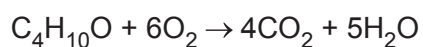
Which alcohol forms a mixture of two alkenes that are structural isomers of one another?

Alcohol**D**.....

[1]

- (c) 25.9 g of alcohol **F** is completely combusted in oxygen at a temperature of 100 °C and 1.50×10^5 Pa pressure.

The equation for this reaction is:



Calculate the volume of $\text{CO}_2(\text{g})$ produced, in cm^3 .

Give your answer to an **appropriate** number of significant figures.

$$M(\text{C}_4\text{H}_{10}\text{O}) = (4 \times 12.0) + (10 \times 1.0) + 16.0 = 74.0 \text{ g mol}^{-1}$$

$$n(\text{F}) = 25.9 \div 74.0 = 0.350 \text{ mol}$$

$$n(\text{CO}_2) = 4 \times 0.350 = 1.40 \text{ mol}$$

$$pV = nRT, \text{ so } V = nRT \div p$$

$$T = 100 + 273 = 373 \text{ K}; \quad p = 1.50 \times 10^5 \text{ Pa}$$

$$V = (1.40 \times 8.314 \times 373) \div 1.50 \times 10^5$$

$$= 0.028944 \text{ m}^3$$

$$= 0.028944 \times 10^6 \text{ cm}^3 = 28944 \text{ cm}^3$$

$$= 2.89 \times 10^4 \text{ cm}^3 \text{ (3 s.f.)}$$

Volume of $\text{CO}_2(\text{g}) = \dots\dots\dots 2.89 \times 10^4 \dots\dots\dots \text{ cm}^3$ [5]

END OF QUESTION PAPER

This image shows a blank sheet of white paper designed for handwriting practice. It features a solid vertical line on the left side, creating a narrow margin. The rest of the page is filled with evenly spaced horizontal dashed lines, providing guides for letter height and placement. There are no other markings, text, or illustrations on the page.

OCR

Oxford Cambridge and RSA

Copyright Information

OCR is committed to seeking permission to reproduce all third-party content that it uses in its assessment materials. OCR has attempted to identify and contact all copyright holders whose work is used in this paper. To avoid the issue of disclosure of answer-related information to candidates, all copyright acknowledgements are reproduced in the OCR Copyright Acknowledgements Booklet. This is produced for each series of examinations and is freely available to download from our public website (www.ocr.org.uk) after the live examination series.

If OCR has unwittingly failed to correctly acknowledge or clear any third-party content in this assessment material, OCR will be happy to correct its mistake at the earliest possible opportunity.

For queries or further information please contact The OCR Copyright Team, The Triangle Building, Shaftesbury Road, Cambridge CB2 8EA.

OCR is part of Cambridge University Press & Assessment, which is itself a department of the University of Cambridge.

© OCR 2025