

GCE

Chemistry A

H032/01: Breadth in chemistry

AS Level

Mark Scheme for June 2022

OCR (Oxford Cambridge and RSA) is a leading UK awarding body, providing a wide range of qualifications to meet the needs of candidates of all ages and abilities. OCR qualifications include AS/A Levels, Diplomas, GCSEs, Cambridge Nationals, Cambridge Technicals, Functional Skills, Key Skills, Entry Level qualifications, NVQs and vocational qualifications in areas such as IT, business, languages, teaching/training, administration and secretarial skills.

It is also responsible for developing new specifications to meet national requirements and the needs of students and teachers. OCR is a not-for-profit organisation; any surplus made is invested back into the establishment to help towards the development of qualifications and support, which keep pace with the changing needs of today's society.

This mark scheme is published as an aid to teachers and students, to indicate the requirements of the examination. It shows the basis on which marks were awarded by examiners. It does not indicate the details of the discussions which took place at an examiners' meeting before marking commenced.

All examiners are instructed that alternative correct answers and unexpected approaches in candidates' scripts must be given marks that fairly reflect the relevant knowledge and skills demonstrated.

Mark schemes should be read in conjunction with the published question papers and the report on the examination.

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Longer Answer Questions (requiring a developed response)

Where candidates have provided two (or more) responses to a medium or high tariff question which only required a single (developed) response and not crossed out the first response, then only the first response should be marked. Examiners will need to apply professional judgement as to whether the second (or a subsequent) response is a 'new start' or simply a poorly expressed continuation of the first response.

6. Always check the pages (and additional objects if present) at the end of the response in case any answers have been continued there. If the candidate has continued an answer there, then add a tick to confirm that the work has been seen.
7. Award No Response (NR) if:
 - there is nothing written in the answer space

Award Zero '0' if:

- anything is written in the answer space and is not worthy of credit (this includes text and symbols).

Team Leaders must confirm the correct use of the NR button with their markers before live marking commences and should check this when reviewing scripts.

8. The RM Assessor **comments box** is used by your team leader to explain the marking of the practice responses. Please refer to these comments when checking your practice responses. **Do not use the comments box for any other reason.** If you have any questions or comments for your team leader, use the phone, the RM Assessor messaging system, or e-mail.
9. *Assistant Examiners will send a brief report on the performance of candidates to their Team Leader (Supervisor) via email by the end of the marking period. The report should contain notes on particular strengths displayed as well as common errors or weaknesses. Constructive criticism of the question paper/mark scheme is also appreciated.*

10. For answers marked by levels of response: Not applicable in F501
- To determine the level** – start at the highest level and work down until you reach the level that matches the answer
 - To determine the mark within the level**, consider the following

<i>Descriptor</i>	<i>Award mark</i>
On the borderline of this level and the one below	At bottom of level
Just enough achievement on balance for this level	Above bottom and either below middle or at middle of level (depending on number of marks available)
Meets the criteria but with some slight inconsistency	Above middle and either below top of level or at middle of level (depending on number of marks available)
Consistently meets the criteria for this level	At top of level

11. Annotations available in RM Assessor

Annotation	Meaning
	Correct response
	Incorrect response
	Omission mark
	Benefit of doubt given
	Contradiction
	Rounding error
	Error in number of significant figures
	Error carried forward
	Level 1
	Level 2
	Level 3
	Benefit of doubt not given
	Noted but no credit given
	Ignore
	Blank page

12. Abbreviations, annotations and conventions used in the detailed Mark Scheme (to include abbreviations and subject-specific conventions).

Annotation	Meaning
DO NOT ALLOW	Answers which are not worthy of credit
IGNORE	Statements which are irrelevant
ALLOW	Answers that can be accepted
()	Words which are not essential to gain credit
—	Underlined words must be present in answer to score a mark
ECF	Error carried forward
AW	Alternative wording
ORA	Or reverse argument

13. Subject-specific Marking Instructions

INTRODUCTION

Your first task as an Examiner is to become thoroughly familiar with the material on which the examination depends. This material includes:

- the specification, especially the assessment objectives
- the question paper
- the mark scheme.

You should ensure that you have copies of these materials.

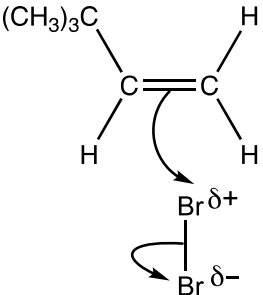
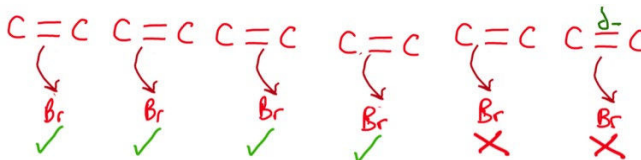
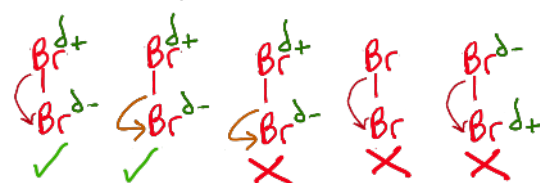
You should ensure also that you are familiar with the administrative procedures related to the marking process. These are set out in the OCR booklet **Instructions for Examiners**. If you are examining for the first time, please read carefully **Appendix 5 Introduction to Script Marking: Notes for New Examiners**.

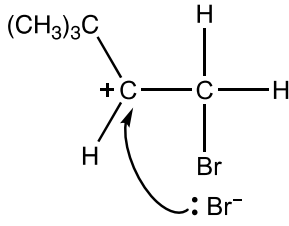
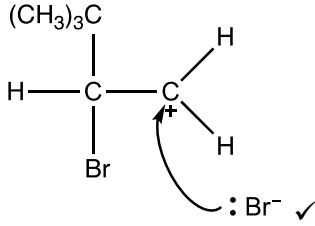
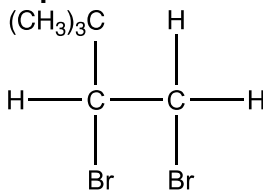
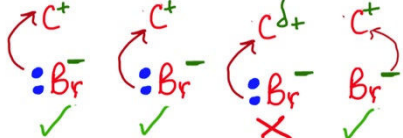
Please ask for help or guidance whenever you need it. Your first point of contact is your Team Leader.

SECTION A

Question	Answer	Marks	AO element	Guidance
1	B	1	AO1.1	
2	A	1	AO1.1	
3	D	1	AO1.1	
4	C	1	AO2.1	ALLOW 7
5	D	1	AO1.2	ALLOW Ar
6	B	1	AO2.1	
7	C	1	AO2.2	
8	B	1	AO2.4	
9	D	1	AO2.6	
10	D	1	AO1.1	
11	A	1	AO2.2	
12	D	1	AO1.1	
13	A	1	AO2.6	
14	D	1	AO1.2	ALLOW 9
15	C	1	AO1.2	ALLOW 12
16	D	1	AO2.6	
17	D	1	AO1.2	
18	B	1	AO1.1	
19	A	1	AO2.5	
20	C	1	AO2.5	
	Total	20		

SECTION B

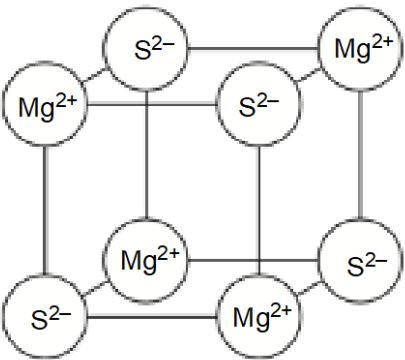
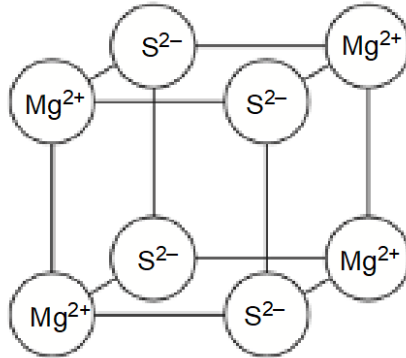
Question			Answer	Marks	AO element	Guidance
21	(a)	(i)	3,3-dimethylbut-1-ene ✓ CARE: Look for dimethyl	1	AO1.2 ×1	IGNORE lack of hyphens, or addition of commas or spaces ALLOW full stops or spaces between numbers e.g. 3.3 dimethyl but-1-ene DO NOT ALLOW meth OR methy
		(ii)	ANNOTATE ANSWER WITH TICKS AND CROSSES  1st curly arrow (from ANY alkene) Curly arrow from double bond to Br of Br-Br ✓ DO NOT ALLOW partial charge on C=C 2nd curly arrow Correct dipole on Br-Br AND curly arrow for breaking of Br-Br bond ✓	5	For curly arrows, ALLOW straight or snake-like arrows and small gaps (see examples): 1st curly arrow must - go to a Br atom of Br-Br AND - start from, OR be traced back to any point across width of C=C  2nd curly arrow must - start from, OR be traced back to, any part of $\delta^+ \text{Br}-\text{Br}^{\delta-}$ bond AND go to Br δ^- 	AO1.2 AO1.2

Question			Answer	Marks	AO element	Guidance
			3rd curly arrow Correct carbocation with + charge on C with 3 bonds AND curly arrow from Br⁻ to C⁺ of carbocation DO NOT ALLOW δ+ on C of carbocation  OR  <i>i.e. ALLOW carbonium + on either C atom</i> Correct product to match mechanism/intermediate ✓  DO NOT ALLOW half headed or double headed arrows but allow ECF if seen more than once Name of mechanism: Electrophilic addition ✓			IGNORE connectivity of CH₃ groups in carbocation and product and ALLOW C₄H₉ 3rd curly arrow must - go to the C⁺ of carbocation AND - start from, OR be traced back to any point across width of lone pair on :Br⁻ OR start from – charge on Br⁻ ion  <i>(Lone pair NOT needed if curly arrow shown from – charge on Br⁻)</i> ALLOW bromonium ion ALLOW any combination of skeletal OR structural OR displayed formula as long as unambiguous NOTE: For a mechanism with HBr, ALLOW all marks EXCEPT for final product mark
(b)	(i)			1		For repeat unit,

Question			Answer	Marks	AO element	Guidance
			$ \begin{array}{c} (\text{CH}_3)_3\text{C} \quad \text{H} \\ \quad \\ \text{---C---C---} \\ \quad \\ \text{H} \quad \text{H} \end{array} $ Correct polymer with side links ✓		AO2.5 ×1	• 'side bonds' required on either side of repeat unit from C atoms • ALLOW more than one repeat unit • ALLOW C₄H₉ for C(CH₃)₃ IGNORE brackets • IGNORE <i>n</i> IGNORE connectivity of C(CH₃)₃ group
	(b)	(ii)	Advantage: Energy/electricity (produced) AND Disadvantage: CO₂ produced OR gases causing global warming/climate change OR greenhouse gases, e.g CO₂ BOTH advantage and disadvantage ✓	1	AO1.1 ×1	ALLOW reduced use of fossil fuels IGNORE produced CO₂ and H₂O ALLOW less landfill / less harm to wildlife or environment (<i>not just harmful</i>) ALLOW toxic/poisonous (waste) products/gases, e.g. CO IGNORE harmful/dangerous

Question		Answer	Marks	AO element	Guidance																		
22	(a)	FIRST CHECK ANSWER ON THE ANSWER LINE IF answer = 190.47 (to 2 DP) award 2 marks $\frac{(188 \times 12.13) + (189 \times 16.75) + (190 \times 27.23) + (192 \times 43.89)}{100}$ OR 190.4677 OR 190.468 ✓ = 190.47 (to 2 DP) ✓	2	AO1.2 ×2	For 1 mark: ALLOW ECF → to 2 DP if: %s used with wrong isotopes ONCE OR - transposed decimal places for ONE %																		
	(b)	<table border="1"> <thead> <tr> <th>Element</th><th>Mass number</th><th>Protons</th><th>Neutrons</th><th>Electrons</th><th>Charge</th></tr> </thead> <tbody> <tr> <td>Ni</td><td>62</td><td>28</td><td>34</td><td>1s²2s²2p⁶3s²3p⁶3d⁸4s²</td><td>0</td></tr> <tr> <td>P</td><td>33</td><td>15</td><td>18</td><td>1s²2s²2p⁶3s²3p⁶</td><td>3–</td></tr> </tbody> </table> ✓ ✓ Mark by row	Element	Mass number	Protons	Neutrons	Electrons	Charge	Ni	62	28	34	1s ² 2s ² 2p ⁶ 3s ² 3p ⁶ 3d ⁸ 4s ²	0	P	33	15	18	1s ² 2s ² 2p ⁶ 3s ² 3p ⁶	3–	2	AO1.2 ×2	<i>Easiest to check element first</i> ALLOW P^{3–} ALLOW names for elements IGNORE charges with element in 1st column, even if wrong. For electron configuration, ALLOW 4s² before 3d⁸ i.e. 1s²2s²2p⁶3s²3p⁶4s²3d⁸ ALLOW upper case D, etc and subscripts, e.g. 4S₂3D₁ ALLOW [Ar]3d⁸4s²
Element	Mass number	Protons	Neutrons	Electrons	Charge																		
Ni	62	28	34	1s ² 2s ² 2p ⁶ 3s ² 3p ⁶ 3d ⁸ 4s ²	0																		
P	33	15	18	1s ² 2s ² 2p ⁶ 3s ² 3p ⁶	3–																		

Question		Answer	Marks	AO element	Guidance
	(c)	Molar ratios $\begin{array}{cccc} \text{Zn} & : & \text{H} & : & \text{N} & : & \text{O} \\ = & \frac{21.99}{65.4} & : & \frac{4.04}{1.0} & : & \frac{9.41}{14.0} & : & \frac{64.56}{16.0} \\ \text{OR} & 0.336 & : & 4.04 & : & 0.672 & : & 4.04 \\ \text{OR} & 1 & : & 12 & : & 2 & : & 12 \checkmark \end{array}$ Empirical formula $\text{ZnH}_{12}\text{N}_2\text{O}_{12} \checkmark$ Any order With water of crystallisation $\text{ZnN}_2\text{O}_6 \cdot 6\text{H}_2\text{O}$ OR $\text{Zn}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O} \checkmark$ Inverse fractions → NO MARKS	3	AO1.2 ×2 AO2.2 ×1	NOTE: If only the correct answer of $\text{ZnN}_2\text{O}_6 \cdot 6\text{H}_2\text{O}$ OR $\text{Zn}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ is seen with no working, award 1 mark only ALLOW ECF from incorrect molar ratios of Zn : H : N : O e.g. from use of atomic number(s) ALLOW $\text{Zn}(\text{NO}_3)_2(\text{H}_2\text{O})_6$ ALLOW ECF from incorrect empirical formula e.g. $\text{ZnNO}_3 \cdot 3\text{H}_2\text{O}$ from ZnH_6NO_6

Question			Answer	Marks	AO element	Guidance
23	(a)	(i)	(Electrostatic) attraction between oppositely charged OR + and – ions ✓	1	AO1.1 ×1	Attraction is essential IGNORE references to metal and non-metal
		(ii)	 Mg²⁺ shown alternately in FOUR circles ✓ S²⁻ shown alternately in FOUR circles ✓	2	AO1.2 ×2	ALLOW labels if seen outside circles provided it clear which circle the label applies to ALLOW 1 mark for Mg AND S shown alternately, each in FOUR circles <i>i.e. with no charges or incorrect charges</i> ALLOW 1 mark for 2+/-+2 AND 2-/-2 shown alternately in FOUR circles (with no Mg and S) DO NOT ALLOW All circles with same ion, <i>i.e. all Mg²⁺ OR all S²⁻</i> ALLOW 1 mark for 4 Mg²⁺ AND 4S²⁻ but NOT shown alternately e.g.  ✓

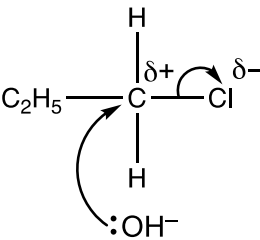
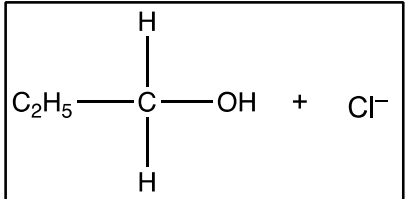
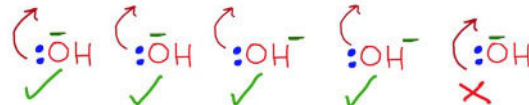
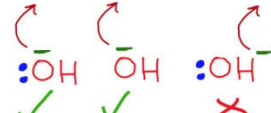
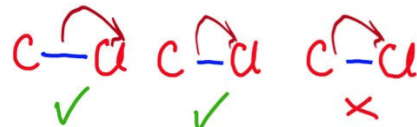
Question			Answer	Marks	AO element	Guidance												
	(b)		<table><tr><th>Name of oxyanion</th><th>Ionic charge</th><th>Formula of oxyanion</th></tr><tr><td>Bromate(III) ✓</td><td>1–</td><td>BrO₂[–]</td></tr><tr><td>Sulfate(VI)</td><td>2–</td><td>SO₄^{2–}</td></tr><tr><td>Phosphate(V)</td><td>3–</td><td>PO₄^{3–} ✓</td></tr></table>	Name of oxyanion	Ionic charge	Formula of oxyanion	Bromate(III) ✓	1–	BrO ₂ [–]	Sulfate(VI)	2–	SO ₄ ^{2–}	Phosphate(V)	3–	PO ₄ ^{3–} ✓	2	AO3.1 ×2	 ALLOW PO ₄ ^{–3}
Name of oxyanion	Ionic charge	Formula of oxyanion																
Bromate(III) ✓	1–	BrO ₂ [–]																
Sulfate(VI)	2–	SO ₄ ^{2–}																
Phosphate(V)	3–	PO ₄ ^{3–} ✓																
	(c)		Structure Giant ✓ Bonding Metallic (bonding) ✓ Particles 2+ /Ca²⁺ ions and delocalised electrons ✓ Conductivity (Delocalised) electrons move/flow ✓ <i>Idea of movement required</i> <i>Delocalised can be seen anywhere</i>	4	AO1.1 ×4	ALLOW marks from labelled diagram ‘Giant metallic’ gains BOTH structure and bonding marks ALLOW attraction between cations and electrons Attraction between nucleus and electrons is CON Watch for ‘metallic’ being CONNed within overall response ALLOW charge flows ONLY when linked to electrons IGNORE electrons carry charge IGNORE electrons are free BUT ALLOW mobile electrons carry charge												

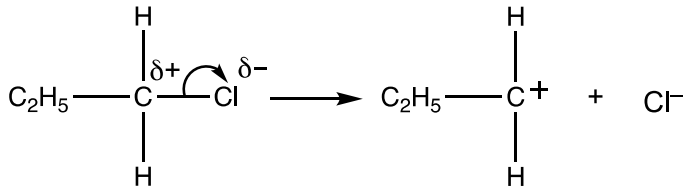
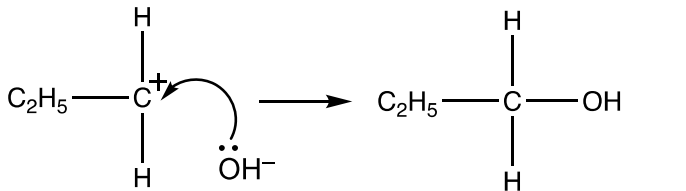
Question			Answer	Marks	AO element	Guidance
24	(a)	(i)	Oxidation and reduction of the same element ✓ 'Atom' is insufficient for element	1	AO1.1 ×1	ALLOW 'chlorine' OR 'Cl' for same element IGNORE 'species' for 'element'
		(ii)	Equation $\text{Cl}_2 + 2\text{NaOH} \rightarrow \text{NaClO} + \text{NaCl} + \text{H}_2\text{O} \quad \checkmark$ Redox: Cl is oxidised from 0 (in Cl ₂) to +1 in NaClO ✓ Cl is reduced from 0 (in Cl ₂) to –1 in NaCl/HCl ✓ IGNORE oxidation numbers shown in equation <i>(treat as rough working)</i> BUT If no oxidation numbers in explanation, <i>look at equation for oxidation numbers</i>	3	AO2.6 ×1 AO2.1 ×2	DO NOT ALLOW $\text{Cl}_2 + \text{NaOH} \rightarrow \text{NaClO} + \text{HCl}$ ALLOW ECF from HCl in equation ALLOW 1 out of 2 redox marks if NaClO AND NaCl omitted, i.e. Cl is oxidised from 0 to +1 AND Cl is reduced from 0 to –1 ALLOW 1 out of 2 redox marks if oxidation number changes are BOTH correct ... BUT reduction/oxidation is incorrectly assigned, i.e. Cl is reduced from 0 (in Cl ₂) to +1 in NaClO Cl is oxidised from 0 (in Cl ₂) to –1 in NaCl/HCl ----- General: ALLOW number before sign in ox no, i.e. 1+ for +1 1– for –1 IGNORE ionic charges, e.g. Cl ¹⁺ IGNORE '1' (signs required) IGNORE references to electron loss/gain (even if wrong)

Question		Answer	Marks	AO element	Guidance
	(b)	Identification of halide Add (aqueous) silver nitrate OR AgNO₃ OR Ag⁺/silver ions ✓ Observations – mark independently Any 2 precipitate colours from Chloride/Cl⁻ gives white precipitate Bromide/Br⁻ gives cream precipitate Iodide/I⁻ gives yellow precipitate ✓ Precipitate/solid seen at least once Equation for at least one halide e.g. Ag⁺ + Cl⁻ → AgCl ALLOW Ag⁺ + X⁻ → AgX ✓ <i>IGNORE state symbols (ppt already assessed)</i> Identification of B and C B: NaBr OR sodium bromide ✓ C: CaCl₂ OR calcium chloride ✓	5	AO3.3 ×3 AO3.2 ×2	ANNOTATE ANSWER WITH TICKS AND CROSSES IGNORE addition of HNO₃ but HCl CONs AgNO₃ IGNORE references to solubility in NH₃ (dil or conc), even if incorrect ALLOW chlorine for chloride, etc ALLOW equation with Br⁻ OR I⁻ e.g. Ag⁺ + Br⁻ → AgBr ALLOW full/partial equations, e.g. AgNO₃ + Cl⁻ → AgCl + NO₃⁻ ALLOW explanation for identification: i.e. B (Group 1): Subtract molar/atomic mass of halide/Br from number in range 100–115/molar mass of B ✓ C (Group 2): Subtract 2 × molar/atomic mass of halide/Cl from number in range 100–115/molar mass of C ✓ ----- ALLOW displacement by addition of halogen ✓ 2 correct colours in water or organic solvent ✓ Equation, e.g. Cl₂ + 2Br⁻ → Br₂ + 2Cl⁻ ✓

Question			Answer	Marks	AO element	Guidance
25	(a)	(i)	FIRST, CHECK THE ANSWER ON ANSWER LINE IF $\Delta_r H = -116$ (kJ mol⁻¹) award 4 marks IF $\Delta_r H = +116$ (kJ mol⁻¹) award 3 marks ----- Energy released in J OR kJ $= 75.0 \times 4.18 \times 18.5 = 5799.75$ (J) OR 5.79975 (kJ) ✓ Correctly calculates $n(\text{Ba(OH)}_2)$ OR $n(\text{HNO}_3)$ $n(\text{Ba(OH)}_2) = 2 \times \frac{25.0}{1000} = 0.05(00)$ (mol) OR $n(\text{HNO}_3) = 2 \times \frac{50.0}{1000} = 0.1(00)$ (mol) ✓ ΔH per mole Ba(OH)_2 in J OR kJ Answer <i>MUST</i> divide energy by $n(\text{Ba(OH)}_2)$ OR $2 \times n(\text{HNO}_3)$ $\pm \frac{5799.75}{0.05}$ OR $\pm 2 \times \frac{5799.75}{0.1} = \pm 115995$ (J) OR $\pm \frac{5.79975}{0.05}$ OR $\pm 2 \times \frac{5.79975}{0.1} = \pm 115.995$ (kJ) ✓ ΔH in kJ mol⁻¹ to 3 SF AND – sign $\Delta_r H = -116$ (kJ mol⁻¹) ✓	4		ANNOTATE ANSWER WITH TICKS AND CROSSES ----- AO2.4 ALLOW 5799.8 OR 5800 J OR 5.7998 OR 5.8 kJ DO NOT ALLOW < 3 SF EXCEPT 5.8 (trailing zeroes) IGNORE any sign AO2.4 IGNORE units <i>i.e.</i> ALLOW correctly calculated number in J OR kJ OR no units ALLOW 3SF or more OR use of 5800 J OR 5.8 kJ AO2.8 Sign NOT needed 3 SF needed ----- AO2.8 Common errors 3 marks $\frac{5799.75}{0.1} \rightarrow -58.0$ no $2 \times$ using 0.1 $\frac{5799.75}{0.15} \rightarrow -38.7$ $\div$ by 0.05 + 0.10 $2 \times \frac{5799.75}{0.15} \rightarrow -77.3$

Question			Answer	Marks	AO element	Guidance
						2 marks for answers above with wrong sign or not to 3 SF Other multiples by using <i>m</i> as 50 or 25: Mark using same principal Use of 50 → -77.3 3 marks Use of 25 → -38.7 3 marks
		(ii)	Reason for incorrect conclusion neutralisation forms 1 mol H₂O OR $\Delta_r H$ forms 2 mol H₂O ✓ Value for $\Delta_{\text{neut}} H = \pm \frac{\text{answer to 25a(i)}}{2}$ (kJ mol ⁻¹) ✓ 2 SF or more	2	AO3.2 ×1	H₂O essential IGNORE sign, even if wrong ALLOW 2 SF, e.g. 58

Question	Answer	Marks	AO element	Guidance
26 (a)	Curly arrow from HO^- to C atom of C–Cl bond ✓ Dipole shown on C–Cl bond, $\text{C}^{\delta+}$ and $\text{Cl}^{\delta-}$ AND curly arrow from C–Cl bond to Cl atom ✓  IGNORE presence of Na^+ but OH^- needed i.e. Na^+OH^- can be allowed if criteria met DO NOT ALLOW H_2O instead of OH^- Correct organic product AND Cl^- ✓  IGNORE presence of Na^+ but Cl^- needed i.e. Na^+Cl^- can be allowed BUT NaCl does NOT show Cl^-	3	AO1.2 AO1.2 AO2.5 ×1	ANNOTATE ANSWER WITH TICKS AND CROSSES NOTE: curly arrows can be straight, snake-like, etc. but NOT double headed or half headed arrows 1st curly arrow must - go to C of C–Cl AND - start from, OR be traced back to any point across width of lone pair on O of OH^-  OR start from – charge on O of OH^- ion  (Lone pair NOT needed if curly arrow shown from O^-) 2nd curly arrow must start from, OR be traced back to, any part of C–Cl bond and go to Cl  ALLOW ECF NaCl^- ONLY from NaOH^-

Question	Answer	Marks	AO element	Guidance
				ALLOW S_N1 mechanism First mark Dipole shown on C–Cl bond, C^{δ+} and Cl^{δ-}, AND curly arrow from C–Cl bond to Cl atom ✓  Second mark Correct carbocation AND curly arrow from HO⁻ to carbocation  Curly arrow must come from lone pair on O of HO⁻ OR OH⁻ OR from minus on O of HO⁻ ion (no need to show lone pair if curly came from negative charge) ✓ Third mark Correct organic product AND Cl⁻ ✓ -----

Question	Answer	Marks	AO element	Guidance
(b)	FIRST check the molar mass on answer line MUST be derived from $pV = nRT$, Award 4 marks for calculation for: answer = 136.9 OR 137 ----- <i>Rearranging ideal gas equation to make n subject</i> $n = \frac{pV}{RT} \checkmark$ <i>Substituting all values including conversion to m^3 and K</i> $n = \frac{(1.01 \times 10^5) \times (74.0 \times 10^{-6})}{8.314 \times 373} \checkmark$ $n = 2.410095443 \times 10^{-3} \rightarrow \begin{matrix} 2.41 \times 10^{-3} \text{ (mol)} \checkmark \\ \text{unrounded} & \text{rounded to 3 SF} \end{matrix}$ <i>Calculation of molar mass, M</i> $M = \frac{m}{n} = \frac{0.330}{2.410095443 \times 10^{-3}} = 136.9.. \text{ (g mol}^{-1}\text{)}$ $\rightarrow \frac{0.330}{2.41 \times 10^{-3}} = 136.9 \text{ (g mol}^{-1}\text{)} \checkmark$ ALLOW calculated M in range 136.9 – 137 <i>Molecular formula of D</i> $C_4H_9Br \checkmark$ ----- IF candidate has failed to derive suitable value of n, ALLOW value of M from 0.330 AND 24000 with haloalkane closest to calculated value for last 2 marks See Guidance column.	5	AO2.4 ×4 AO3.2	ANNOTATE ANSWER WITH TICKS AND CROSSES ----- If there is an alternative answer, check to see if there is any ECF credit possible using working below 1st mark may be implicit by direct substitution of correct values below into rearranged equation. ALLOW use of 8.31 for $R \rightarrow 2.411 \times 10^{-3}$ ONLY award this mark if n has been derived from correct rearranged ideal gas equation ALLOW 3 SF up to calculator value, correctly rounded 2.41 × 10⁻³ OR 0.002411255537 → first 3 marks → 136.868581616 → C_4H_9Br NOTE: ALLOW 137 (i.e. to 3 SF) ALLOW any unambiguous structure ALLOW ECF provided that formula given is a haloalkane and matches M calculated from 0.330 g AND $pV = nRT$ ----- $M = \frac{0.330}{74.0/24000} \text{ OR } \frac{0.330}{3.0833.. \times 10^{-3}}$ = 107 to 3 SF ✓ From 107, ONLY ALLOW = C_2H_5Br (108.9) ✓

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