Surname	Centre Number	Candidate Number
Other Names		2

GCE A LEVEL

A410U20-1

S19-A410U20-1

TUESDAY, 11 JUNE 2019 – AFTERNOON

CHEMISTRY – A level component 2
Organic Chemistry and Analysis

2 hours 30 minutes

	For Examiner's use only		
	Question	**Maximum Mark**	**Mark Awarded**
Section A	**1. to 7.**	**15**	
Section B	**8.**	**21**	
	9.	**17**	
	10.	**18**	
	11.	**15**	
	12.	**17**	
	13.	**17**	
	Total	**120**	

A410U201
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ADDITIONAL MATERIALS

In addition to this examination paper, you will need a:
- calculator;
- **Data Booklet** supplied by WJEC.

INSTRUCTIONS TO CANDIDATES

Use black ink or black ball-point pen.

Write your name, centre number and candidate number in the spaces at the top of this page.

Section A Answer **all** questions in the spaces provided.
Section B Answer **all** questions in the spaces provided.

Candidates are advised to allocate their time appropriately between **Section A (15 marks)** and **Section B (105 marks)**.

INFORMATION FOR CANDIDATES

The number of marks is given in brackets at the end of each question or part-question.

The maximum mark for this paper is 120.

Your answers must be relevant and must make full use of the information given to be awarded full marks for a question.

The assessment of the quality of extended response (QER) will take place in **Q.9***(a)* and **Q.12***(a)*.

If you run out of space, use the additional page(s) at the back of the booklet, taking care to number the question(s) correctly.

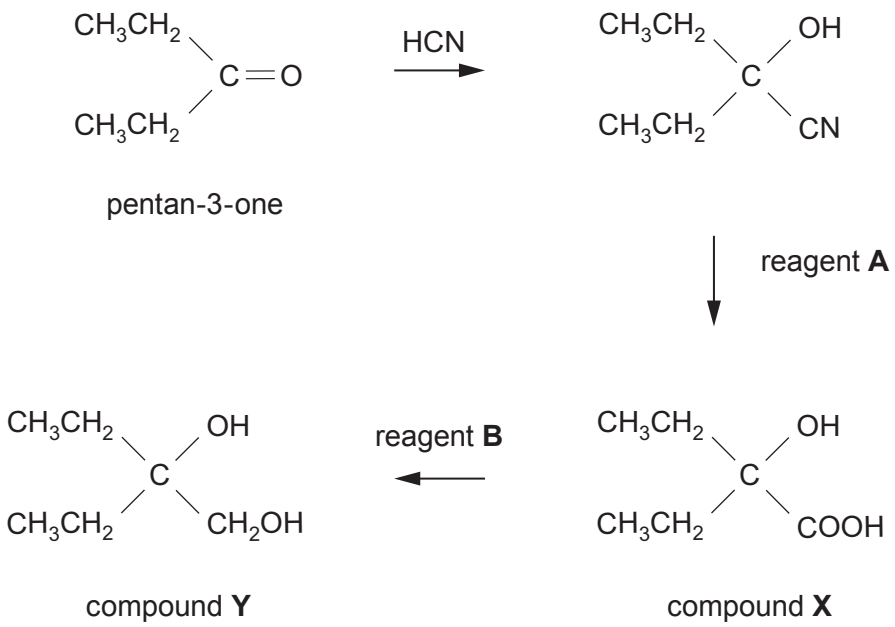
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SECTION A

Answer **all** *questions in the spaces provided.*

1. Compound **Y** can be made by the following method.



(a) State the type of reaction mechanism taking place when hydrogen cyanide reacts with pentan-3-one. [1]

.....

(b) State the name of reagent **A**. [1]

.....

(c) Give the empirical formula of compound **X**. [1]

.....

(d) Identify reagent **B**. [1]

.....

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2. *N*-(phenylazo)phenylamine, $C_6H_5—N{=}N—NH—C_6H_5$, is prepared by reacting benzenediazonium chloride and phenylamine.

(a) State the reagents and conditions needed to prepare a solution of benzenediazonium chloride from phenylamine. [2]

.....

.....

.....

(b) In white light *N*-(phenylazo)phenylamine is seen as a yellow solid.

Explain why this compound is yellow in white light. [1]

.....

.....

3. An aqueous solution contains sodium benzoate, $C_6H_5COO^-Na^+$, and benzamide, $C_6H_5CONH_2$.

A student was asked to suggest a simple chemical method to produce a solution containing only sodium benzoate and water. He said that he would heat the mixture with aqueous sodium hydroxide.

Explain why this suggestion is correct. [1]

.....

.....

4. A colourless solution contains either benzoic acid or 2-hydroxybenzaldehyde, $C_6H_4(OH)CHO$.

State **two** chemical tests, apart from using indicator paper, which can be used to identify which compound is present. [2]

	Reagent(s) used	Result with benzoic acid	Result with 2-hydroxybenzaldehyde
Test 1			
Test 2			

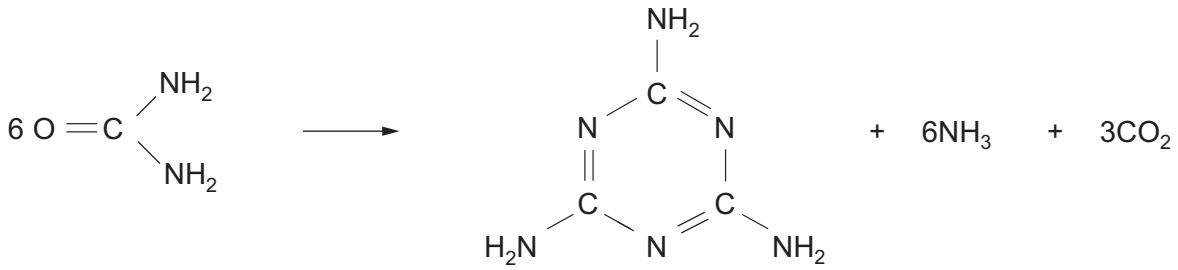
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5. Draw the **displayed** formula of the ionic compound calcium ethanoate, $(CH_3COO)_2Ca$. [1]

6. Draw a dot and cross diagram of the methyl radical. [1]

7. Melamine is made by heating urea.



(a) Calculate the atom economy of this reaction. [2]

Atom economy = ... %

(b) There have been a number of instances of the illegal addition of melamine to baby milk. This addition appears to increase the percentage of protein in the milk, by increasing its nitrogen content.

Calculate the percentage of nitrogen by mass in melamine. [1]

Percentage = ... %

15

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SECTION B

Answer **all** *questions in the spaces provided.*

8. *(a)* Alkanes react with chlorine to give chloroalkanes.

(i) Give the equation for the reaction of methylpropane, $(CH_3)_3CH$, with chlorine, giving 2-chloro-2-methylpropane as one of the products.

Your answer should show the structures of the reactants and products. [1]

(ii) The chlorination of methylpropane also produces a number of polychlorinated alkanes.

State how these polychlorinated products are formed. [1]

.....

.....

(iii) The formula of methylpropane shows that 9 of the 10 hydrogen atoms are equivalent and therefore there is a greater chance of 1-chloro-2-methylpropane being the principal organic product in the first stage of the chlorination reaction.

However, the main product of this chlorination is 2-chloro-2-methylpropane.
Suggest a reason for this compound being the main product. [1]

.....

.....

.....

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(iv) Another product of this reaction is a hydrocarbon of molecular formula C_8H_{18}, whose high resolution ^1H NMR spectrum shows only a singlet at 0.86 ppm.

Suggest a structure for this hydrocarbon, explaining why the spectrum shows a singlet. [2]

.....

.....

(v) Methylpropane is replacing CFCs as a domestic refrigerant. While this material is seen as more environmentally friendly, it is very flammable.

Give the equation for the complete combustion of methylpropane. [1]

.....

(b) Halogenated alkanes cause environmental problems in the ozone layer by their interaction with UV radiation.

Complete the equation below, which shows the effect of UV radiation on chlorofluoromethane. Explain why these products are formed. [2]

$CH_2ClF \longrightarrow$

.....

.....

.....

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(c) Methane clathrate is the name given to methane trapped inside ice crystals. The permafrost contains extensive quantities of this material and there are worries that global warming may lead to the uncontrolled emission of vast quantities of methane.

Use the information below to show that the mole ratio of methane to water (ice) in methane clathrate is about 1 : 5.8.

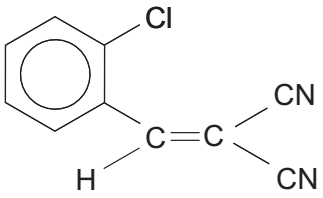
You should assume that methane clathrate is a material consisting of only methane and ice. [4]

- 1 dm^3 of methane clathrate gives 168 dm^3 of methane measured at 273 K and 1 atm pressure
- 1 dm^3 of methane clathrate has a mass of 900 g

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(d) CS gas (actually a solid) is a substance used to quell social unrest.



(i) The disposal of surplus stocks of CS gas is a problem which is causing concern. One way is by incineration.

CS gas —(high temperature)→ chlorinated hydrocarbons, hydrogen cyanide gas, polycyclic aromatic hydrocarbons

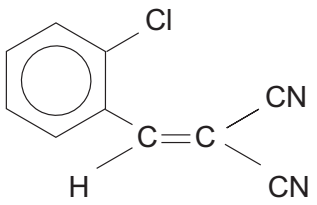
Suggest **one** reason why this method is seen as an unsatisfactory method of disposal. [1]

.....

.....

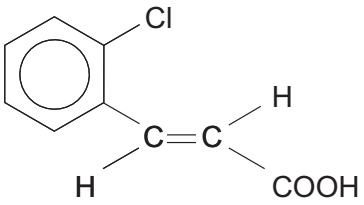
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(ii) Another method of disposal is by the conversion of CS gas to (2-chlorophenyl)ethene, ammonium sulfate, carbon dioxide and other products.



M_r 189

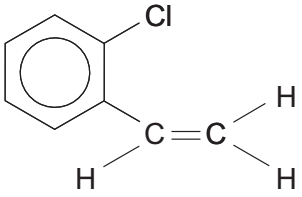
stage 1 →



+ CO_2 + $(NH_4)_2SO_4$

compound **P**

stage 2 ↓ heat with sodalime



+ Na_2CO_3 + H_2O

M_r 139

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I. Stage 1 is an example of a **hydrolysis** reaction.

State what is meant by this term. [1]

.....

.....

II. If stage 1 were carried out in the laboratory a mixture would be obtained, containing ammonium sulfate in solution, together with a precipitate of the organic acid.

State how you would obtain pure crystals of the acid from this mixture. [1]

.....

.....

III. The melting temperature of compound **P** is 209 °C.

State how this figure would change if a sample of compound **P** were impure. [1]

.....

.....

IV. State the type of reaction occurring in stage 2. [1]

.....

V. The yield of compound **P** in stage 1 was 90.0 % and when compound **P** was converted to (2-chlorophenyl)ethene in stage 2 the yield was 50.0 %.

Calculate the **overall** percentage yield and use your answer to calculate the mass of (2-chlorophenyl)ethene obtained if 75.0 kg of CS gas were destroyed in this way.

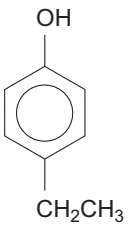
Give your answer to an **appropriate** number of significant figures. [4]

Mass = ... kg

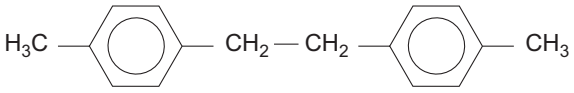
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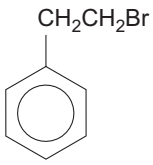
9. *(a)* An unknown compound is one of the six compounds whose structures are shown below.



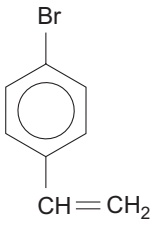
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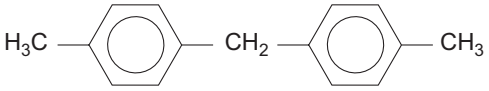
B



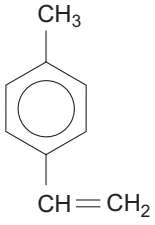
C



D



E



F

Show how you could use infrared spectroscopy to identify **four** of these compounds. In your answer you should state the bonds identified and their absorption values.

For the remaining **two** compounds discuss how high resolution ^1H NMR spectroscopy could be used to find the structure of the unknown compound. Your answer should include comments about the splitting pattern seen in the spectra. [6 QER]

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(b) Compound **L** is an aromatic compound of formula $C_8H_8Br_2$ (M_r 264).

A sample of compound **L** of mass 3.22 g was refluxed with excess aqueous sodium hydroxide and the resulting mixture acidified with nitric acid. On adding aqueous silver nitrate to this mixture 1.22×10^{-2} mol of silver bromide was obtained.

Use this information to suggest a structure for compound **L**. Explain your answer. [4]

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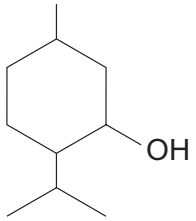
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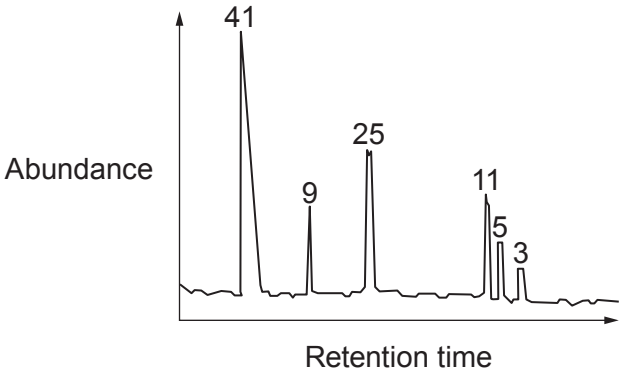
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(c) Menthol, $C_{10}H_{20}O$, is a natural product that has important medicinal uses.



(i) Menthol can be isolated from peppermint oil. Analysis of this oil shows that its major component is menthol. It contains smaller quantities of other compounds. A simplified gas chromatogram of peppermint oil (showing relative peak areas) is shown below.



I. Calculate the percentage by volume of menthol in the oil. [1]

Percentage = ... %

II. The peak of relative area 5 is given by limonene. Limonene is a non-cyclic hydrocarbon whose mass spectrum shows a molecular ion (m/z) at 136.

Use this information to suggest a molecular formula for limonene.
Explain why limonene must be an unsaturated compound. [2]

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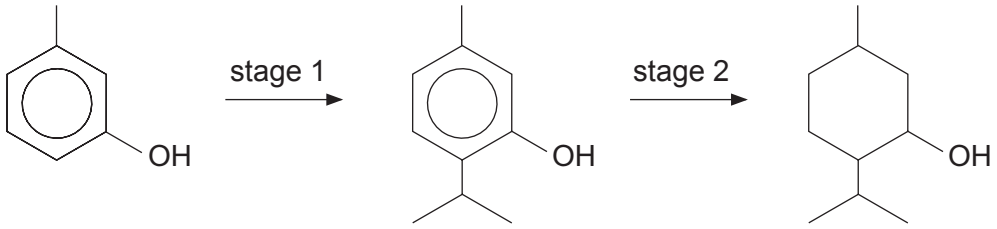
.....

.....

Molecular formula .....

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(ii) One method for producing menthol is shown below.



menthol

I. Give the systematic name for the starting material. [1]

.....

II. Stage 1 is the Friedel-Crafts alkylation of the starting material.

Give the names of the reagent and the catalyst used in this stage. [1]

Reagent .....

Catalyst .....

III. State the type of reaction occurring in Stage 2. [1]

.....

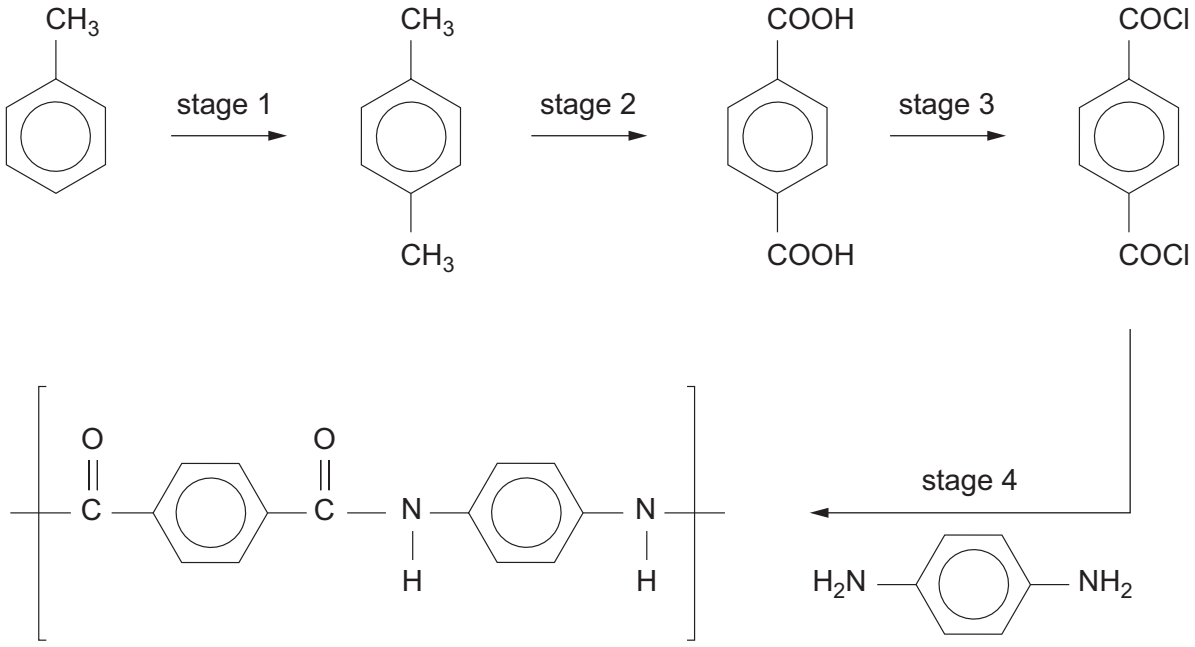
(iii) Menthol is a secondary alcohol.

Suggest the **skeletal** formula of menthane, $C_{10}H_{18}O$, which is formed by the oxidation of menthol. [1]

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10. *(a)* A student was asked to propose a synthesis for the polyamide Kevlar®, starting from methylbenzene. He suggested the following sequence of reactions.



Kevlar®

(i) State the type of reaction mechanism occurring in stage 1. [1]

.....

(ii) Suggest why the yield of 1,4-dimethylbenzene in stage 1 will be considerably less than 100 %. Problems with the method should not be considered as part of the answer. [1]

.....

.....

(iii) State the reagent(s) needed to convert 1,4-dimethylbenzene to benzene-1,4-dicarboxylic acid in Stage 2. [1]

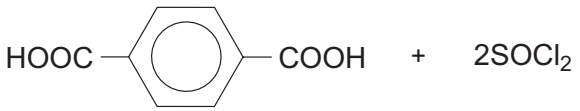
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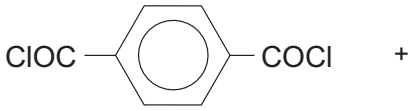
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(iv) In stage 3 the dicarboxylic acid is converted to the acid dichloride.
The student found that the preferred reagent for this reaction was sulfur dichloride oxide, $SOCl_2$.

I. Complete the equation for this reaction. [1]





II. The conversion to the acid dichloride can also be carried out using PCl_3 or PCl_5.

Suggest why the preferred reagent for this reaction is $SOCl_2$. [1]

.....

.....

III. Suggest why ethanol should not be used as a solvent for this reaction. [1]

.....

(v) In Stage 4 Kevlar® is produced.

State why this polymer is described as a polyamide. [1]

.....

.....

(vi) Benzene-1,4-diamine, $C_6H_4(NH_2)_2$, used in stage 4, can be converted to benzene-1,4-diol, $C_6H_4(OH)_2$.

Suggest a reagent(s) that can be used for this reaction. [1]

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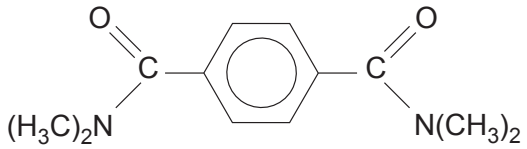
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(vii) The student suggested that benzene-1,4-diol can react with the acid dichloride formed in stage 3 to give a different polymer.

Suggest a formula for this polymer, showing a repeating section. [1]

(viii) While researching a method for making Kevlar®, the student came across the compound below.



I. Describe the **low** resolution ^{1}H NMR spectrum of this compound.
Relative peak areas are required in your answer but **not** the position of each signal. [2]

.....

.....

II. Describe the ^{13}C NMR spectrum of this compound identifying the carbon atoms involved. The position of the signals is **not** required. [2]

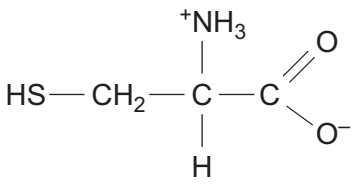
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(b) (i) At pH 5.1 the *α*-amino acid cysteine exists as the zwitterion.

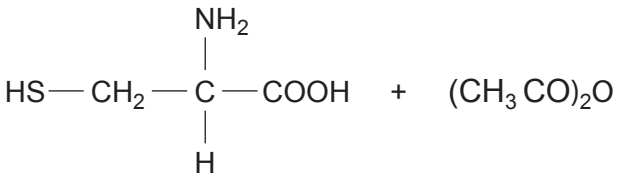


Write the structure of the organic species present when cysteine is added to an aqueous solution of pH 1.0. [1]

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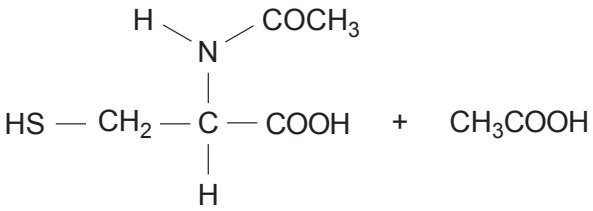
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(ii) *N*-ethanoylcysteine is used to treat an overdose of paracetamol.
It can be produced by the reaction of cysteine with ethanoic anhydride.



cysteine

M_r 121



N-ethanoylcysteine

M_r 163

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In an experiment 0.250 mol of cysteine reacts in a 1 : 1 ratio with ethanoic anhydride.

I. Ethanoic anhydride has a density of 1.08 g cm^{-3}.

Calculate the volume of ethanoic anhydride needed. [2]

Volume = ... cm^3

II. After the reaction a 90 % yield of *N*-ethanoylcysteine is formed.

Calculate the mass of *N*-ethanoylcysteine formed. [1]

Mass = ... g

(iii) Cysteine is an *α*-amino acid that contains a chiral centre.

Write the displayed formula of an *α*-amino acid that does **not** contain a chiral centre. [1]

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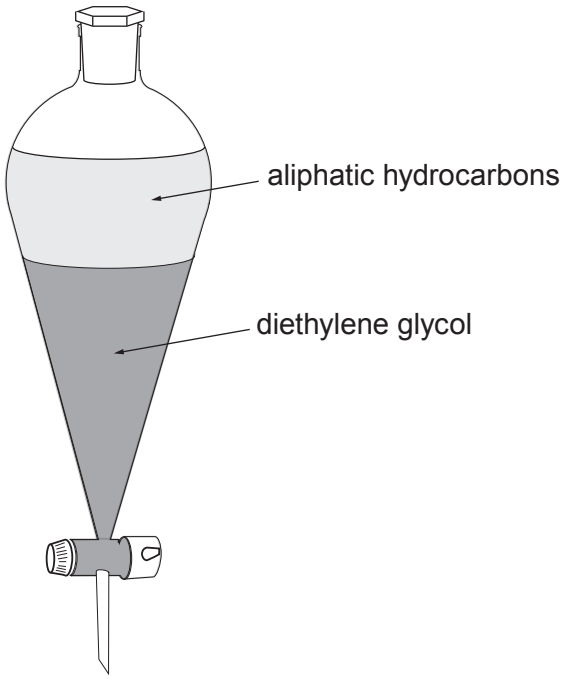
11. *(a)* (i) Naphtha, obtained from crude oil (petroleum) contains hexane.

Write the equation for the conversion (reforming) of hexane into benzene and hydrogen. [1]

.....

(ii) The reforming of naphtha produces a number of other liquid products in addition to benzene.

I. One method of separating benzene from aliphatic products is to use solvent extraction. Benzene dissolves in diethylene glycol whereas aliphatic hydrocarbons are largely insoluble in this solvent. In the laboratory this can be demonstrated using a separating funnel.



Describe the use of a separating funnel containing the two layers shown above, to obtain a separate sample of the diethylene glycol layer. [1]

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Examiner only

II. The boiling temperatures of benzene and diethylene glycol are 80 °C and 245 °C respectively.

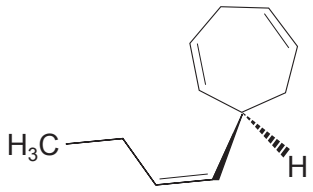
State the name of the technique used to separate these two liquids and state how you could make sure that the risk of fire is reduced to a minimum. [2]

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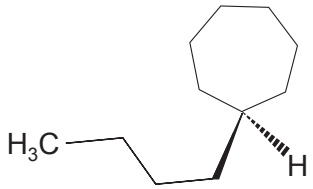
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(b) Ectocarpene is a hydrocarbon produced by certain algae.



(i) 4.31 dm^3 of hydrogen, measured at 312 K and at a pressure of 1.01×10^5 Pa, was needed to completely react with 8.29 g of ectocarpene to produce compound **G**.



compound **G**

Use this information to confirm that the relative molecular mass of ectocarpene is 148. [4]

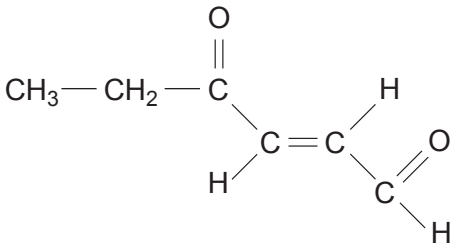
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(ii) Find the molecular formula of ectocarpene and deduce the structural formula of an aromatic compound that has the same molecular formula as ectocarpene. [2]

Molecular formula .....

Structural formula

(c) A particular insect defends itself by producing a mixture of compounds that contains compound **W**.



compound **W**

(i) Give the structure of the compound obtained when compound **W** reacts with sodium tetrahydridoborate(III), $NaBH_4$. [1]

(ii) Compound **W** is warmed with Tollens' reagent.

Describe what is seen (if anything) and explain your observation. [2]

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Examiner only

(d) Another insect produces a defensive mixture that contains 2-methylbutanoic acid.

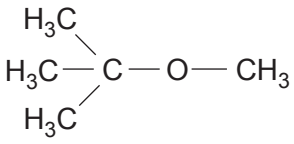
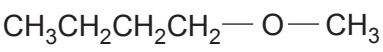
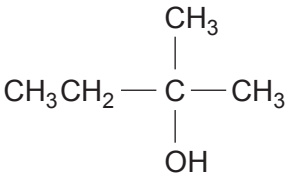
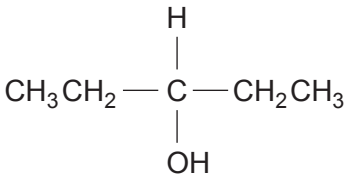
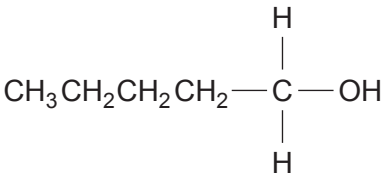
Write the displayed formula of 2-methylbutanoic acid. Give an isomer of this compound that has only three signals in its ^{13}C NMR spectrum, one of which occurs at 160-185 ppm. [2]

Structure of 2-methylbutanoic acid

Structure of the isomer

15

12. *(a)* The boiling temperatures of five isomers of formula $C_5H_{12}O$ are shown in the table below.

Compound	Formula	Boiling temperature / °C
A		56
B		71
C		102
D		116
E		137

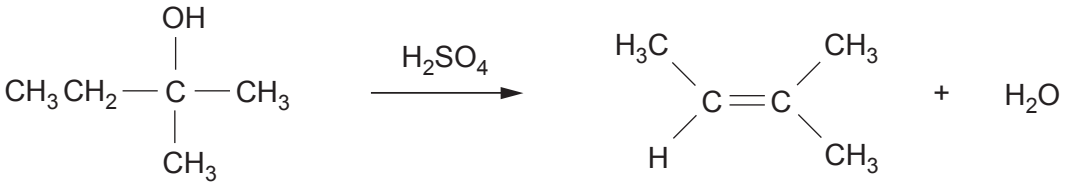
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Discuss the intermolecular bonding present in these compounds and relate this bonding to the differences observed in their boiling temperatures. [6 QER]

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(b) The equation below shows the reaction of 2-methylbutan-2-ol with sulfuric acid.



(i) State why the sulfuric acid acts as a catalyst in this reaction. [1]

.....

.....

(ii) In the first stage of the mechanism of this reaction, 2-methylbutan-2-ol acts as a base.

Suggest how it can react as a base. [1]

.....

.....

(iii) The main organic product of the reaction is 2-methylbut-2-ene but a small percentage of 2-methylbut-1-ene, $CH_3CH_2C(CH_3)=CH_2$ is also produced.

Suggest why this alkene is also a product. [1]

.....

.....

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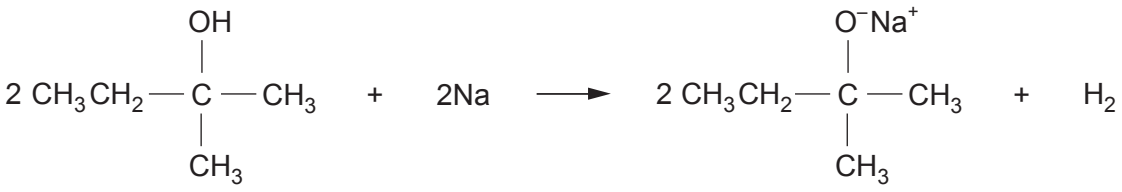
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(iv) The mixture at the end of the reaction contains the two alkenes and a small quantity of unreacted 2-methylbutan-2-ol.

A student found that 2-methylbutan-2-ol reacts with sodium to give hydrogen as one of the products.



A 10.0 g sample of the mixture was reacted with an excess of sodium. This produced 125 cm^3 of hydrogen measured at 298 K and 1 atm pressure.

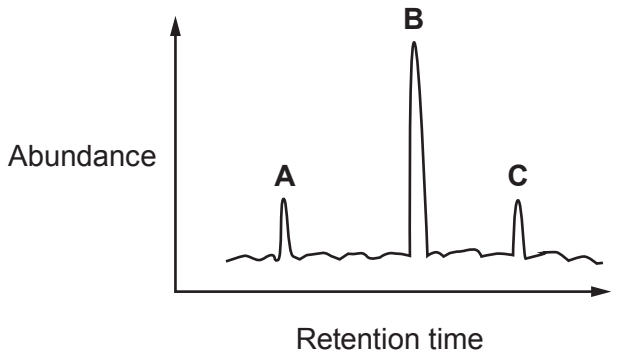
Calculate the mass of 2-methylbutan-2-ol in the mixture. [2]

Mass = ... g

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(v) A sample of the product mixture from this reaction produced the following gas chromatogram.



I. State, giving a reason, which of these peaks is given by 2-methylbut-2-ene. [1]

.....

.....

II. Suggest how you would confirm which one of the remaining two peaks is given by 2-methylbutan-2-ol. [1]

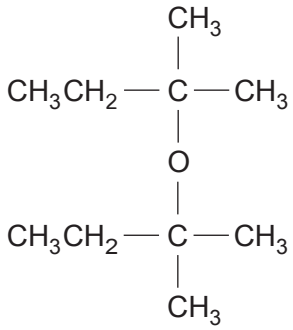
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(vi) In this reaction the alcohol is added slowly to an excess of sulfuric acid. However, if sulfuric acid is added slowly to the alcohol, then the product below is also formed.



Suggest how this compound can be formed in the reaction. [1]

.....

.....

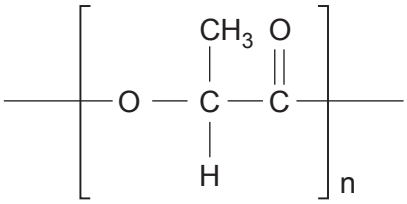
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(c) (i) Many polymers persist in the environment for many years and are causing severe pollution problems. There is an increasing use of materials that lead to fewer environmental problems.

Some cardboard drinking cups have an impervious PLA lining. PLA is made from plants and is biodegradable.

The formula of the repeating section of PLA is shown below.



Give the structure of the monomer of PLA. [1]

(ii) Suggest **two** desirable properties of PLA if it is to be used for lining the cardboard of a cup used for hot drinks. [2]

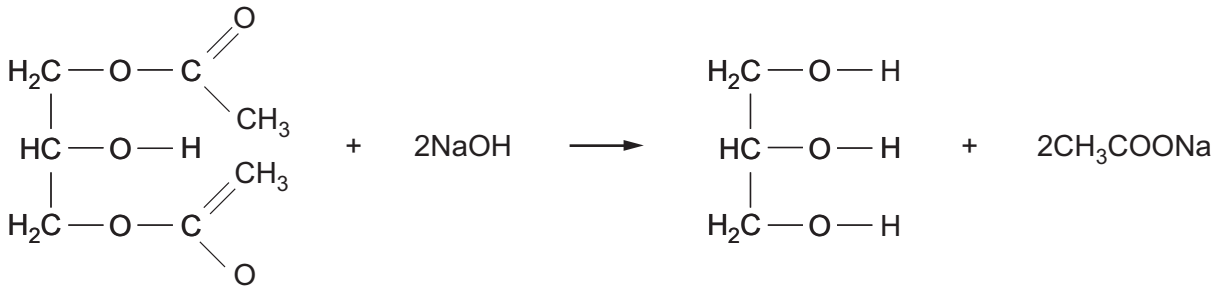
1. .....

2. .....

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13. *(a)* Diacetin is an ester of propane-1,2,3-triol that is used by some plants to attract bees. It reacts with aqueous sodium hydroxide to produce propane-1,2,3-triol and sodium ethanoate.



diacetin

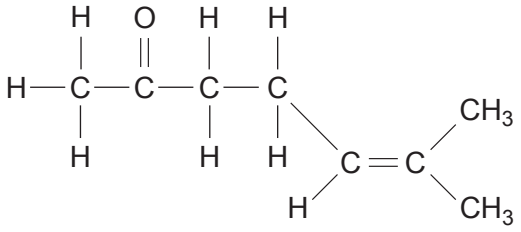
1.58 g of diacetin was heated with 50.0 cm^3 of aqueous sodium hydroxide of concentration 0.500 mol dm^{-3}. After complete reaction with the diacetin 7.00×10^{-3} mol of sodium hydroxide remained.

Use this information to show that the relative molecular mass of diacetin is 176. [4]

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(b) 2-Methylhept-2-en-6-one acts as an alarm pheromone in certain species of ants.



2-methylhept-2-en-6-one

(i) State why this compound does not show *E*-*Z* isomerism. [1]

.....

.....

(ii) Draw the structure of a ketone with the same molecular formula that **can** exist as *E*-*Z* isomers. [1]

(iii) 2-Methylhept-2-en-6-one gives a positive triiodomethane test.

State the reagent(s) used for this test and the observation. [2]

Reagent(s) .....

Observation .....

(iv) The reaction of 2-methylhept-2-en-6-one with aqueous sodium tetrahydridoborate(III) results in the formation of a new compound which has forms that rotate the plane of plane polarised light.

Give the structure of the compound formed and explain why it has forms that rotate the plane of plane polarised light. [2]

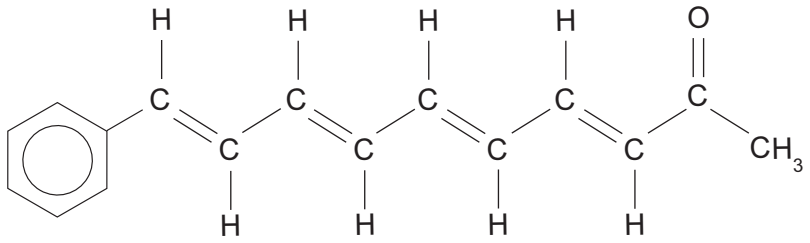
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(c) (i) A species of sea slug produces an alarm pheromone which is a mixture of three compounds. One of these compounds is Navenone B.



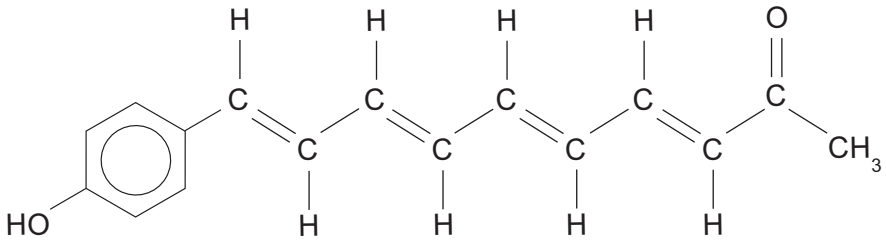
The visible spectrum of this compound shows an absorption maximum at 377 nm.

Calculate the energy of this absorption in kJ mol^{-1}. [3]

Energy = ... kJ mol^{-1}

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(ii) Navenone C is a similar compound.



This compound reacts with hydrogen bromide and with bromine.

I. State what is seen when bromine is added to a solution of Navenone C. [2]

.....

.....

II. State the type of reaction mechanism occurring when Navenone C reacts with hydrogen bromide. [1]

.....

.....

III. Complete the table below to show the mole ratios of the added reagent and Navenone C.

The table is already completed for the reaction with bromine. [1]

Reagent added	Mole ratio Added reagent : Navenone C
hydrogen bromide	
bromine	6 : 1

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