

Cambridge IGCSE™

COMBINED SCIENCE**0653/41**

Paper 4 Theory (Extended)

October/November 2024**MARK SCHEME**Maximum Mark: 80

Published

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

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This document consists of **12** printed pages.

PUBLISHED**Generic Marking Principles**

These general marking principles must be applied by all examiners when marking candidate answers. They should be applied alongside the specific content of the mark scheme or generic level descriptions for a question. Each question paper and mark scheme will also comply with these marking principles.

GENERIC MARKING PRINCIPLE 1:

Marks must be awarded in line with:

- the specific content of the mark scheme or the generic level descriptors for the question
- the specific skills defined in the mark scheme or in the generic level descriptors for the question
- the standard of response required by a candidate as exemplified by the standardisation scripts.

GENERIC MARKING PRINCIPLE 2:

Marks awarded are always **whole marks** (not half marks, or other fractions).

GENERIC MARKING PRINCIPLE 3:

Marks must be awarded **positively**:

- marks are awarded for correct/valid answers, as defined in the mark scheme. However, credit is given for valid answers which go beyond the scope of the syllabus and mark scheme, referring to your Team Leader as appropriate
- marks are awarded when candidates clearly demonstrate what they know and can do
- marks are not deducted for errors
- marks are not deducted for omissions
- answers should only be judged on the quality of spelling, punctuation and grammar when these features are specifically assessed by the question as indicated by the mark scheme. The meaning, however, should be unambiguous.

GENERIC MARKING PRINCIPLE 4:

Rules must be applied consistently, e.g. in situations where candidates have not followed instructions or in the application of generic level descriptors.

GENERIC MARKING PRINCIPLE 5:

Marks should be awarded using the full range of marks defined in the mark scheme for the question (however; the use of the full mark range may be limited according to the quality of the candidate responses seen).

GENERIC MARKING PRINCIPLE 6:

Marks awarded are based solely on the requirements as defined in the mark scheme. Marks should not be awarded with grade thresholds or grade descriptors in mind.

Science-Specific Marking Principles

- 1 Examiners should consider the context and scientific use of any keywords when awarding marks. Although keywords may be present, marks should not be awarded if the keywords are used incorrectly.
- 2 The examiner should not choose between contradictory statements given in the same question part, and credit should not be awarded for any correct statement that is contradicted within the same question part. Wrong science that is irrelevant to the question should be ignored.
- 3 Although spellings do not have to be correct, spellings of syllabus terms must allow for clear and unambiguous separation from other syllabus terms with which they may be confused (e.g. ethane / ethene, glucagon / glycogen, refraction / reflection).
- 4 The error carried forward (ecf) principle should be applied, where appropriate. If an incorrect answer is subsequently used in a scientifically correct way, the candidate should be awarded these subsequent marking points. Further guidance will be included in the mark scheme where necessary and any exceptions to this general principle will be noted.
- 5 'List rule' guidance

For questions that require ***n*** responses (e.g. State **two** reasons ...):
 - The response should be read as continuous prose, even when numbered answer spaces are provided.
 - Any response marked *ignore* in the mark scheme should not count towards ***n***.
 - Incorrect responses should not be awarded credit but will still count towards ***n***.
 - Read the entire response to check for any responses that contradict those that would otherwise be credited. Credit should **not** be awarded for any responses that are contradicted within the rest of the response. Where two responses contradict one another, this should be treated as a single incorrect response.
 - Non-contradictory responses after the first ***n*** responses may be ignored even if they include incorrect science.

6 Calculation specific guidance

Correct answers to calculations should be given full credit even if there is no working or incorrect working, **unless** the question states 'show your working'.

For questions in which the number of significant figures required is not stated, credit should be awarded for correct answers when rounded by the examiner to the number of significant figures given in the mark scheme. This may not apply to measured values.

For answers given in standard form (e.g. $a \times 10^n$) in which the convention of restricting the value of the coefficient (a) to a value between 1 and 10 is not followed, credit may still be awarded if the answer can be converted to the answer given in the mark scheme.

Unless a separate mark is given for a unit, a missing or incorrect unit will normally mean that the final calculation mark is not awarded. Exceptions to this general principle will be noted in the mark scheme.

7 Guidance for chemical equations

Multiples / fractions of coefficients used in chemical equations are acceptable unless stated otherwise in the mark scheme.

State symbols given in an equation should be ignored unless asked for in the question or stated otherwise in the mark scheme.

Mark scheme abbreviations

;	separates marking points
/	separates alternative responses for the same marking point
ecf	error carried forward
AVP	any valid point
ORA	or reverse argument
AW	alternative wording
<u>underline</u>	actual word given must be used by candidate (grammatical variants accepted)
terms in bold	shows the idea that must be seen but AW acceptable
()	the word / phrase in brackets is not required but sets the context

Question	Answer	Marks
1(a)(i)	B ; D ;	2
1(a)(ii)	<i>any two from:</i> exchange / transfer, of (dissolved) nutrients ; exchange / transfer, of (dissolved) gases ; transfer / removal, of excretory products ; barrier to toxins ;	2
1(b)	flagellum ; (presence of) enzymes ;	2
1(c)	<i>any two from:</i> humans have sperm (cells) AND plants have pollen (grains) ; humans have eggs/ova AND plants have ovules ; takes place in, oviducts in humans AND ovaries in plants ;	2
1(d)	insect / animal / bird ; spikes to help them stick (to pollinators) ;	2

Question	Answer	Marks																		
2(a)(i)	idea that: both have one <u>electron</u> in their outer shell so are in Group 1 ; K is lower in group because it has one more electron shell than Na ;	2																		
2(a)(ii)	23 ;	1																		
2(b)	<table><tr><td></td><td>proton number</td><td>nucleon number</td><td>number of protons</td><td>number of neutrons</td><td>number of electrons</td></tr><tr><td>Li atom</td><td>3</td><td>7</td><td>3</td><td>4</td><td>3</td></tr><tr><td>Cu²⁺ ion</td><td>29</td><td>64</td><td>29</td><td>35</td><td>27</td></tr></table> number of protons column correct ; number of neutrons column correct ; number of electrons for Li atom correct ; number of electrons for Cu ²⁺ ion correct ;		proton number	nucleon number	number of protons	number of neutrons	number of electrons	Li atom	3	7	3	4	3	Cu ²⁺ ion	29	64	29	35	27	4
	proton number	nucleon number	number of protons	number of neutrons	number of electrons															
Li atom	3	7	3	4	3															
Cu ²⁺ ion	29	64	29	35	27															
2(c)(i)	high density / malleable / good conductors of heat / good conductors of electricity ;	1																		
2(c)(ii)	form coloured compounds ; act as catalysts ;	2																		

Question	Answer	Marks
3(a)	evidence of, $W = mg$; $24.1 \div 2.45 = 9.84$; N / kg ;	3
3(b)	evidence of, $GPE = mgh / W \times d / 24.1 \times 1.48$; $35.7 / 36$ (J) ;	2
3(c)	(calculation of area $A = 0.64 \times 0.25 / 0.16$ (m ²) ; evidence of, $p = F \div A / 24.1 \div 0.16$; $150 / 151$ (Pa) ;	3

Question	Answer	Marks
4(a)	<i>any three from:</i> increasing humidity causes decrease in transpiration rate ; concentration / percentage of water outside leaf increases ; reduces (water) concentration gradient ; reduces (rate of), diffusion / evaporation ;	3
4(b)	$6 \text{ CO}_2 + 6 \text{ H}_2\text{O} \longrightarrow \text{C}_6\text{H}_{12}\text{O}_6 + 6 \text{ O}_2$ correct formula for glucose ; equation correctly balanced dependent on glucose formula ;	2
4(c)	unequal / AW ; <u>elongation</u> ; tip / AW ;	3

Question	Answer	Marks
5(a)	s ;	1
5(b)(i)	2HCl ; H_2O ;	2
5(b)(ii)	from 1 – 3 to 4 – 7 ; the idea that acid is being neutralised / used up ;	2
5(c)(i)	bromine ;	1
5(c)(ii)	bromine is less reactive than chlorine ORA ;	1
5(c)(iii)	bonding pair shown ; 6 non-bonding electrons in each atom ;	2

Question	Answer	Marks																
6(a)	the <u>temperature</u> at which a substance changes state from solid to liquid / AW ;	1																
6(b)(i)	highest pitch = highest frequency / 4186 Hz ; evidence of $v = f\lambda$ / $340 \div 4186$; 0.081(2) (m) ;	3																
6(b)(ii)	oscillations / vibrations (of the air / particles / molecules) ; (motion of air / particles / molecules is) parallel to the direction of, energy transfer / travel / the wave ;	2																
6(c)	refraction ; caused by change in speed / medium / density (of medium) ; sound wave moves faster in water / liquid / denser medium / ORA ;	3																
6(d)	<table><tr><th>property</th><th>solids</th><th>liquids</th><th>gases</th></tr><tr><td>fixed volume</td><td>✓</td><td>(✓)</td><td>×</td></tr><tr><td>fixed shape</td><td>✓</td><td>×</td><td>(X)</td></tr><tr><td>ability to flow</td><td>×</td><td>(✓)</td><td>✓</td></tr></table> solids column correct ; liquids AND gases columns correct ;	property	solids	liquids	gases	fixed volume	✓	(✓)	×	fixed shape	✓	×	(X)	ability to flow	×	(✓)	✓	2
property	solids	liquids	gases															
fixed volume	✓	(✓)	×															
fixed shape	✓	×	(X)															
ability to flow	×	(✓)	✓															

Question	Answer	Marks
7(a)(i)	eagle ; daisy ;	2
7(a)(ii)	<i>any two from:</i> respiration ; movement ; excretion ; egestion ; decomposition ;	2
7(b)(i)	<i>substrate</i> starch ; <i>enzyme</i> protease ; <i>product</i> fatty acids AND glycerol ;	3
7(b)(ii)	to kill bacteria ; to provide, acid pH / low pH, for enzymes ;	2

Question	Answer	Marks														
8(a)	<u>enhanced</u> ;	1														
8(b)	<table border="1"> <tr> <td rowspan="2">reaction that makes carbon dioxide</td><td colspan="2">compound</td></tr> <tr> <td>formula</td><td>name</td></tr> <tr> <td>complete combustion of a gaseous compound</td><td>CH₄</td><td>methane</td></tr> <tr> <td>a solid with a dilute acid</td><td>CaCO₃</td><td>calcium carbonate</td></tr> <tr> <td>thermal decomposition</td><td>CaCO₃</td><td>calcium carbonate</td></tr> </table> one mark for each correct row ;;;	reaction that makes carbon dioxide	compound		formula	name	complete combustion of a gaseous compound	CH ₄	methane	a solid with a dilute acid	CaCO ₃	calcium carbonate	thermal decomposition	CaCO ₃	calcium carbonate	3
reaction that makes carbon dioxide	compound															
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8(c)(i)	(similar) size / chain length ;	1														
8(c)(ii)	<i>any two from:</i> high temperature ; high pressure ; catalyst ;	2														

Question	Answer	Marks
9(a)	battery ;	1
9(b)	correct symbol for voltmeter ; connected in parallel across the fixed resistor only ;	2
9(c)(i)	evidence of, $R = V \div I$ / $6.0 \div 0.25$; = 24 (Ω);	2
9(c)(ii)	determination of total resistance of three-lamp combination, e.g. by subtraction of fixed resistor from total resistance of circuit / $24 - 15 / 9$; evidence / use of formula to calculate resistance of two lamps in parallel, $\frac{1}{R_T} = \frac{1}{R_1} + \frac{1}{R_2} / R_T = \frac{R_1 R_2}{R_1 + R_2} / \frac{R \times R}{R + R} / \frac{R^2}{2R} / \frac{6 \times 6}{6 + 6} / \frac{36}{12} ;$ calculation of resistance of three-lamp combination, $R + \frac{R \times R}{R + R} / R + \frac{R^2}{2R} / R + \frac{R}{2} / 1.5R / 6 + \frac{6 \times 6}{6 + 6} / R + \frac{R}{2} = 9 / 18 = 3R / 1.5R = 9$ AND $R = 6 (\Omega) ;$	3

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()	the word / phrase in brackets is not required but sets the context

Question	Answer	Marks
1(a)(i)	anther labelled A on flower X ;	1
1(a)(ii)	insect / bird / animal ; anthers inside flower / stigma inside flower ;	2
1(b)	palisade ; chloroplasts ; phloem ;	3
1(c)(i)	organic AND energy ;	1
1(c)(ii)	primary ;	1
1(d)	increased aerobic respiration by decomposers in the water ; increased decomposition of dead producers in the water ;	2

Question	Answer	Marks															
2(a)(i)	hydrogen; rubidium hydroxide ;	2															
2(a)(ii)	idea that rubidium is more reactive / faster / more vigorous ORA ;	1															
2(a)(iii)	red ;	1															
2(b)	11 ; 12 ; 10 ;	3															
2(c)	<table> <thead> <tr> <th></th><th>true</th><th>false</th></tr> </thead> <tbody> <tr> <td>Sodium is a liquid at room temperature.</td><td>☐</td><td>☒</td></tr> <tr> <td>Rubidium is more dense than potassium.</td><td>☒</td><td>☐</td></tr> <tr> <td>The density of sodium is less than 0.50 g/cm³.</td><td>☐</td><td>☒</td></tr> <tr> <td>The melting point of rubidium is 70 °C.</td><td>☐</td><td>☒</td></tr> </tbody> </table> any two correct ; all correct ;		true	false	Sodium is a liquid at room temperature.	☐	☒	Rubidium is more dense than potassium.	☒	☐	The density of sodium is less than 0.50 g/cm ³ .	☐	☒	The melting point of rubidium is 70 °C.	☐	☒	2
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The density of sodium is less than 0.50 g/cm ³ .	☐	☒															
The melting point of rubidium is 70 °C.	☐	☒															
2(d)	positive and negative <u>ions</u> ; alternating / regular (arrangement) ;	2															

Question	Answer	Marks
3(a)	chemical (potential) ;	1
3(b)(i)	acceleration = gradient of graph / $20 \div 10$ / change in speed $\div$ time ; $2 \text{ (m/s}^2\text{)} ;$	2
3(b)(ii)	(any relevant area) $100 / 800 / 200 ;$ (all three relevant areas) $100 + 800 + 200 / 1100 ;$ (unit conversion) $1.1 \text{ km} ;$	3
3(b)(iii)	recognition that maximum speed is $20 \text{ m/s} ;$ evidence of, $\text{KE} = \frac{1}{2} m v^2 / \frac{1}{2} \times 1200 \times 20^2 ;$ $240\,000 \text{ (J)} ;$	3

Question	Answer	Marks
4(a)	blood vessel adaptation function W X Y thick elastic walls permeable walls large lumen for diffusion of materials to withstand high blood pressure for less resistance to blood flow two lines correct ; four lines correct ; six lines correct ;	3
4(b)(i)	idea of blockage of, a blood vessel / blood flow ; <u>coronary artery</u> ;	2
4(b)(ii)	329 000 000 / 0.659 million / 0.002 / ratio of N:P $\times$ 100 ; 0.200 (%) ;	2
4(b)(iii)	<i>any two from:</i> chronic obstructive pulmonary disease / COPD ; lung cancer ; AVP ;	2

Question	Answer	Marks
5(a)	carbon ; platinum ;	2
5(b)(i)	hydrogen ;	1
5(b)(ii)	Cl_2 ; $2e^-$;	2
5(c)	sodium ; chlorine ;	2

Question	Answer	Marks
6(a)(i)	(total resistance =) $18 + 15 / 33 (\Omega)$; $V = IR$ (in any form) / 0.20×33 ; 6.6 (V) ;	3
6(a)(ii)	evidence of, $E = VIt / \frac{330}{6.6 \times 0.20}$ / 1.32 (W) ; 250 (s) ;	2
6(b)(i)	$(1.05 - 0.41 =) 0.64 (A)$;	1
6(b)(ii)	evidence of, $R_T = \frac{R_F \times R_G}{R_F + R_G}$ OR $\frac{1}{R_T} = \frac{1}{R_F} + \frac{1}{R_G}$ OR $\frac{51 \times 33}{51 + 33}$; 20 (Ω) ;	2

Question	Answer	Marks
7(a)	C ; is <u>complementary</u> (in shape, to the active site of the enzyme) ;	2
7(b)(i)	biuret (solution) ;	1
7(b)(ii)	protein intake needs to increase ; for the growth / development of the fetus ;	2
7(b)(iii)	gastric juice / stomach, is a lower pH / acidic pH OR protease works (best) in, a lower pH / (more) acidic pH/conditions ; protease is denatured (in small intestine) ; active site changes shape / enzyme substrate complex cannot form / substrate/protein no longer fits (active site of) protease;	3

Question	Answer	Marks
8(a)	calcium chloride ; water ;	2
8(b)	calcium carbonate is heated (to a high temperature) ; (and) breaks down / breaks bonds / separates atoms ;	2
8(c)	$\text{Fe}_2\text{O}_3 + 3\text{CO} \rightarrow 2\text{Fe} + (3\text{CO}_2)$ correct formulae ; balanced ;	2
8(d)(i)	methane / AVP ;	1
8(d)(ii)	increase in global temperature / climate change / named effect of climate change, e.g. ice caps melting, flooding, etc. ;	1

Question	Answer	Marks
9(a)(i)	continuation of the three straight rays to converge at X ;	1
9(a)(ii)	focal point / (principal) focus ;	1
9(b)	evidence of, $\rho = m \div V$ / $0.23 \div 920$; $0.00025 / 2.5 \times 10^{-4} \text{ (m}^3\text{)}$;	2
9(c)	(intermolecular) forces are strong(er) in solid / weak(er) in liquid ; molecules / they (only) vibrate, in solid ; molecules / they move around / move freely / flow / slide over each other, in liquid ;	3
9(d)	<i>any three from:</i> heated wax, expands / increases in volume / particles move further apart / ORA ; hot wax has lower density / cool wax has higher density ; hot wax rises ; cool wax sinks ;	3

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6 Calculation specific guidance

Correct answers to calculations should be given full credit even if there is no working or incorrect working, **unless** the question states 'show your working'.

For questions in which the number of significant figures required is not stated, credit should be awarded for correct answers when rounded by the examiner to the number of significant figures given in the mark scheme. This may not apply to measured values.

For answers given in standard form (e.g. $a \times 10^n$) in which the convention of restricting the value of the coefficient (a) to a value between 1 and 10 is not followed, credit may still be awarded if the answer can be converted to the answer given in the mark scheme.

Unless a separate mark is given for a unit, a missing or incorrect unit will normally mean that the final calculation mark is not awarded. Exceptions to this general principle will be noted in the mark scheme.

7 Guidance for chemical equations

Multiples / fractions of coefficients used in chemical equations are acceptable unless stated otherwise in the mark scheme.

State symbols given in an equation should be ignored unless asked for in the question or stated otherwise in the mark scheme.

Mark scheme abbreviations

;	separates marking points
/	separates alternative responses for the same marking point
ecf	error carried forward
AVP	alternative valid point
ORA	or reverse argument
AW	alternative wording
<u>underline</u>	actual word given must be used by candidate (grammatical variants accepted)
()	the word / phrase in brackets is not required but sets the context

Question	Answer	Marks
1(a)(i)	grasshopper / rabbit / caterpillar ;	1
1(a)(ii)	hawk AND owl ;	1
1(a)(iii)	snake AND hawk AND (blue) jay ;	1
1(b)	energy is lost at each trophic level ; named example of energy loss, e.g. heat / respiration / excretion / protein synthesis / metabolism / not all organism is eaten ; not enough energy left (to support trophic level five) ;	3

Question	Answer	Marks
2(a)(i)	$\text{Zn(s)} + \dots 2 \dots \text{HCl(aq)} \rightarrow \text{ZnCl}_2 \text{(aq)} + \dots \text{H}_2 \dots \text{(g)}$ 2 added for balancing ; H_2 as product ;	2
2(a)(ii)	<i>any two from:</i> solid zinc reduces in size / AW ; fizzing / bubbles (of gas) ; AVP ;	2
2(a)(iii)	by filtration ;	1
2(a)(iv)	heat ; evaporate / remove water ;	2
2(b)	sodium has a +1 charge AND zinc has a +2 charge ; chloride ions have a -1 charge AND idea that charges must balance ;	2

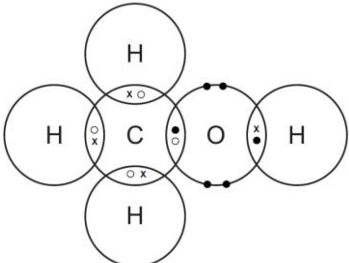
Question	Answer	Marks																		
3(a)	vibrating ; 20 ; gases ;	3																		
3(b)	(because of its) change in speed / different speeds ;	1																		
3(c)	<table><tr><th>volume and shape</th><th>fluidity</th><th>molecular motion</th><th>molecular separation</th><th>intermolecular forces</th></tr><tr><td>fixed volume, fixed shape</td><td>can flow</td><td>molecules move only by vibrating about fixed positions</td><td>molecules <u>are</u> close together</td><td>no forces between molecules</td></tr><tr><td>fixed volume, no fixed shape</td><td rowspan="2">cannot flow</td><td>molecules move around <u>while</u> still touching each other</td><td rowspan="2">molecules <u>are</u> far apart</td><td>moderate forces between molecules</td></tr><tr><td>no fixed volume, no fixed shape</td><td>molecules move <u>quickly</u> in all directions</td><td>strong forces between molecules</td></tr></table> two columns correct ; four columns correct ;	volume and shape	fluidity	molecular motion	molecular separation	intermolecular forces	fixed volume, fixed shape	can flow	molecules move only by vibrating about fixed positions	molecules <u>are</u> close together	no forces between molecules	fixed volume, no fixed shape	cannot flow	molecules move around <u>while</u> still touching each other	molecules <u>are</u> far apart	moderate forces between molecules	no fixed volume, no fixed shape	molecules move <u>quickly</u> in all directions	strong forces between molecules	2
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3(d)	evidence of, $v = f\lambda / 3.0 \times 10^8 \div 1.2 \times 10^7$; 25 (m) ;	2																		

Question	Answer	Marks
4(a)(i)	(most) enzyme denatured; <u>active site</u> changes shape / substrate no longer fits <u>active site</u> ; rate of photosynthesis decreases (so fewer bubbles) ;	3
4(a)(ii)	gas exchange takes place through stomata / carbon dioxide enters through stomata / oxygen exits through stomata ; carbon dioxide is needed for photosynthesis / oxygen is produced in photosynthesis ;	2
4(a)(iii)	iodine (solution) ; colour changes (from brown) to blue-black ;	2
4(b)	amylase ; (simpler) sugars ; salivary ; pancreas ;	4

Question	Answer	Marks
5(a)	compound ; electricity ;	2
5(b)	$2e^-$; H_2 ;	2
5(c)	(hydroxide ions) lose electrons / ions are discharged; water (also) formed ;	2
5(d)	acid remains / acid is not used up / only water is electrolysed ;	1

Question	Answer	Marks
6(a)(i)	0.40 (m / s) ;	1
6(a)(ii)	(calculation of any relevant area) 160 / 1280; (all three relevant areas) 160 + 1280 + 160 / 1600 (m) ; 1.6 (km) ;	3
6(b)(i)	evidence of, $W = Fd / 462\,000 \div 1100$; 420 (m) ;	2
6(b)(ii)	idea that some energy is also transferred to the surroundings ; as, thermal / sound / AVP ;	2
6(b)(iii)	evidence of, $P = W \div t / 792\,000 \div 330$; 2400 ; W / watt(s) ;	3

Question	Answer	Marks
7(a)(i)	A pulmonary artery ; B vena cava ;	2
7(a)(ii)	left atrium labelled ;	1
7(a)(iii)	reference to contraction ; (muscles of) ventricles (contract) ;	2
7(b)(i)	<i>any two from:</i> diet ; stress ; smoking ; genetic predisposition ; age ; AVP ;	2
7(b)(ii)	<i>any two from:</i> both / male / female / number of people increasing AND then decreasing ; males higher (than females) ORA ; AVP ; <i>plus one for:</i> correct use of supporting data, e.g. both graphs reach peak in 2004 / 10 000 females and 22 000 males in 1978 ;	3

Question	Answer	Marks												
8(a)	 four pairs bonding electrons around carbon correct ; one bonding pair between oxygen and hydrogen (O-H bond) correct ; four non-bonding electrons on oxygen shell AND no extra electrons (on hydrogen) ;	3												
8(b)(i)	ethene, has a double bond / shares 4 electrons OR ethane, has only single bonds / has single bond between carbons / shares 2 electrons ;	1												
8(b)(ii)	<table border="0"> <thead> <tr> <th></th><th>true</th><th>false</th></tr> </thead> <tbody> <tr> <td>One molecule of ethene produces the same number of molecules of carbon dioxide as one molecule of ethane.</td><td>☒</td><td>☐</td></tr> <tr> <td>One molecule of ethene produces the same number of molecules of water as one molecule of ethane.</td><td>☐</td><td>☒</td></tr> <tr> <td>The combustion reactions of ethene and ethane are both endothermic.</td><td>☐</td><td>☒</td></tr> </tbody> </table> two correct ; three correct ;		true	false	One molecule of ethene produces the same number of molecules of carbon dioxide as one molecule of ethane.	☒	☐	One molecule of ethene produces the same number of molecules of water as one molecule of ethane.	☐	☒	The combustion reactions of ethene and ethane are both endothermic.	☐	☒	2
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The combustion reactions of ethene and ethane are both endothermic.	☐	☒												
8(c)	<i>alkene</i> C ₅ H ₁₀ ; <i>alkane</i> C ₄ H ₁₀ ;	2												
8(d)	<i>alkene</i> no colour / colourless ; <i>alkane</i> orange ;	2												

Question	Answer	Marks
9(a)	fuse ;	1
9(b)(i)	(current in headlight branch = $11.0 - 1.0 = 10.0$ (A) ;	1
9(b)(ii)	(current in one headlight = $\frac{1}{2} \times 10.0 = 5.0$ (A) ; evidence of, $R = V \div I$ / $12 \div 5.0$; 2.4 (Ω) ; <i>alternative method:</i> evidence of, $R = V \div I$ or $12 \div 10.0$ or 1.2(Ω) ; (resistance of parallel arrangement) 1.2×2 OR double the calculated resistance ; 2.4 (Ω) ;	3
9(c)	resistance of wire proportional to length / $3\times$ longer gives $3\times$ resistance / 3×3 or 9 ; resistance of wire inversely proportional to cross-sectional area / half cross-sectional area gives $2\times$ resistance / 2×3 or 6 ; (resistance of rear lamp is $6\times$ resistance of headlamp $\Rightarrow$) 18 (Ω) ;	3