



CANDIDATE
NAME

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CENTRE
NUMBER

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9701/21

October/November 2021

1 hour 15 minutes

You must answer on the question paper.

You will need: Data booklet

- Answer **all** questions.
- Use a black or dark blue pen. You may use an HB pencil for any diagrams or graphs.
- Write your name, centre number and candidate number in the boxes at the top of the page.
- Write your answer to each question in the space provided.
- Do **not** use an erasable pen or correction fluid.
- Do **not** write on any bar codes.
- You may use a calculator.
- You should show all your working, use appropriate units and use an appropriate number of significant figures.

- The total mark for this paper is 60.
- The number of marks for each question or part question is shown in brackets [].

This document has **16** pages. Any blank pages are indicated.

Answer **all** the questions in the spaces provided.

1 Sulfides are compounds that contain sulfur but not oxygen.

(a) Carbon disulfide, CS_2 , is a volatile liquid at room temperature and pressure.

(i) State the meaning of *volatile*.

..... [1]

(ii) Draw a 'dot-and-cross' diagram of the CS_2 molecule.

[2]

(iii) Suggest the bond angle in a molecule of CS_2 .

..... [1]

(iv) CS_2 is a liquid under room conditions, while CO_2 is a gas.

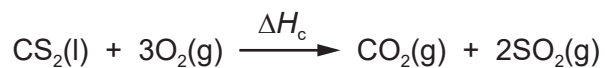
Explain what causes the difference in the physical properties between CS_2 and CO_2 .

.....

.....

..... [2]

(b) The enthalpy change of combustion of $\text{CS}_2(\text{l})$ is represented by the following equation.



(i) Define *enthalpy change of combustion*.

.....

 [2]

(ii) The table shows the enthalpy changes of formation of $\text{CS}_2(\text{l})$, $\text{CO}_2(\text{g})$ and $\text{SO}_2(\text{g})$.

compound	enthalpy change of formation, $\Delta H_f / \text{kJ mol}^{-1}$
$\text{CS}_2(\text{l})$	+89.7
$\text{CO}_2(\text{g})$	-394
$\text{SO}_2(\text{g})$	-297

Use the data in the table to calculate the enthalpy change of combustion, ΔH_c , of $\text{CS}_2(\text{l})$, in kJ mol^{-1} .

Show your working.

ΔH_c of $\text{CS}_2(\text{l})$ = ... kJ mol^{-1}
 [2]

(c) Hydrogen sulfide gas, $\text{H}_2\text{S}(\text{g})$, is slightly soluble in water. It acts as a weak acid in aqueous solution.

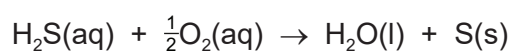
(i) State the meaning of *weak acid*.

.....
..... [1]

(ii) Give the formula of the conjugate base of H_2S .

..... [1]

(iii) $\text{H}_2\text{S}(\text{aq})$ reacts slowly with oxygen dissolved in water. The reaction is represented by the following equation.

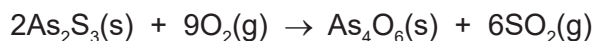


Explain, with reference to oxidation numbers, why this reaction is a redox reaction.

.....
.....
.....
..... [2]

- (d) The compound As_2S_3 is a common mineral.

When As_2S_3 is heated strongly in air, it forms a mixture of products, as shown.



- (i) A sample containing 0.198 g As_2S_3 is placed in 0.100 dm³ of pure oxygen, an excess, in a reaction chamber connected to a gas syringe at room temperature.

The reactants are heated until no further change is observed. The products are then allowed to cool to room temperature.

Calculate the volume, in dm³, of gas present at the end of the experiment.

The molar volume of gas is 24.0 dm³ mol⁻¹ under these conditions. Assume that the pressure is constant throughout the experiment.

Show your working.

volume of gas remaining = ... dm³
[4]

- (ii) State the environmental consequences of releasing $\text{SO}_2(\text{g})$ into the atmosphere.

..... [1]

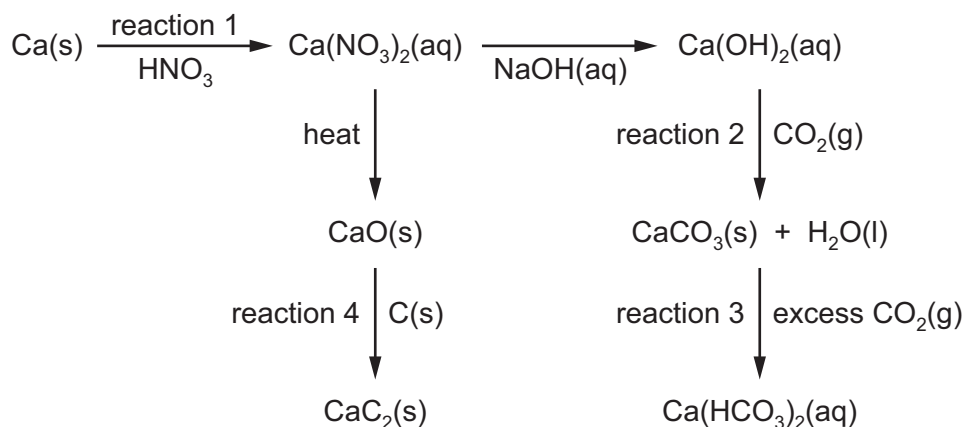
- (iii) $\text{SO}_2(\text{g})$ can be removed from the air by reacting it with $\text{NaOH}(\text{aq})$.

Construct an equation for the reaction of $\text{SO}_2(\text{g})$ with $\text{NaOH}(\text{aq})$. Include state symbols.

..... [2]

[Total: 21]

2 The reaction scheme shows some reactions of calcium.



(a) (i) Reaction 1 produces $\text{Ca(NO}_3)_2$ and one other product.

Identify the other product.

..... [1]

(ii) Construct an equation for the thermal decomposition of $\text{Ca(NO}_3)_2\text{(s)}$.

..... [1]

(iii) State the trend in the thermal stability of the Group 2 nitrates down the group.

..... [1]

(iv) In reaction 3, excess CO_2 is bubbled through water containing CaCO_3 . A solution of $\text{Ca(HCO}_3)_2\text{(aq)}$ forms.

Construct an equation for reaction 3.

..... [1]

(b) Describe how Ca(OH)_2 is used in agriculture.

.....
 [1]

(c) In reaction 4, calcium carbide, CaC_2 , is formed from CaO .

CaC_2 contains the C_2^{2-} anion. Each carbon in C_2^{2-} is sp hybridised.

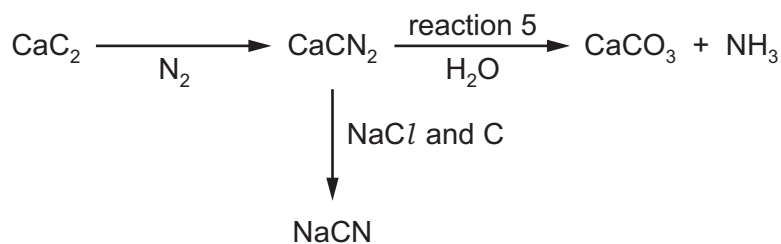
(i) Describe how sp hybridised orbitals are formed.

.....
..... [1]

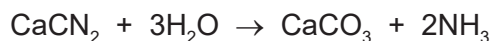
(ii) Sketch a diagram to show how two sp hybrid orbitals can form a sigma (σ) bond.

[2]

(d) The flowchart shows some reactions of CaC_2 .



(i) Reaction 5 can be used to prepare NH_3 .



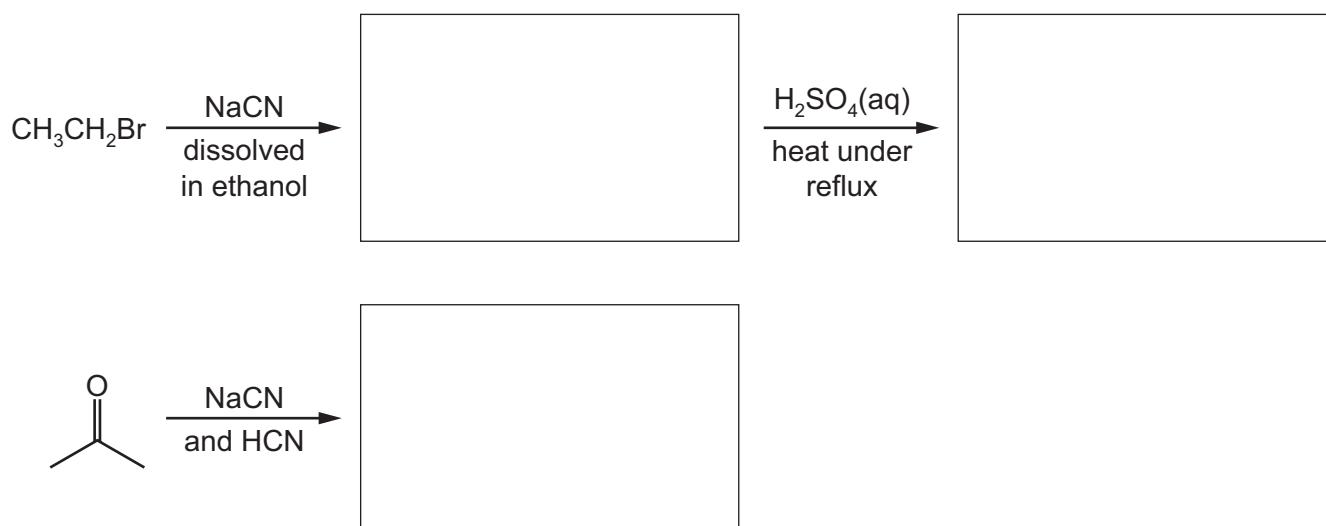
Calculate the minimum mass, in tonnes, of calcium cyanamide, CaCN_2 , that is required to produce 1.50×10^6 tonnes of NH_3 .

Show your working.

$$1 \text{ tonne} = 1.00 \times 10^6 \text{ g}$$

minimum mass of CaCN_2 = ... tonnes
[2]

(ii) Draw the structure of the organic products formed in the following reactions.



[3]

[Total: 13]

3 Phosphorus is a reactive Period 3 element.

(a) Phosphorus has several allotropes. Details of two allotropes are given.

allotrope of phosphorus	formula	melting point/°C
white	P ₄	44
red	P	590

(i) White phosphorus and red phosphorus both have covalent bonding.

Suggest the types of structure shown by white phosphorus (P₄) and red phosphorus (P).

Explain why red phosphorus (P) has a higher melting point than white phosphorus (P₄).

structure of P₄

structure of P

explanation

.....

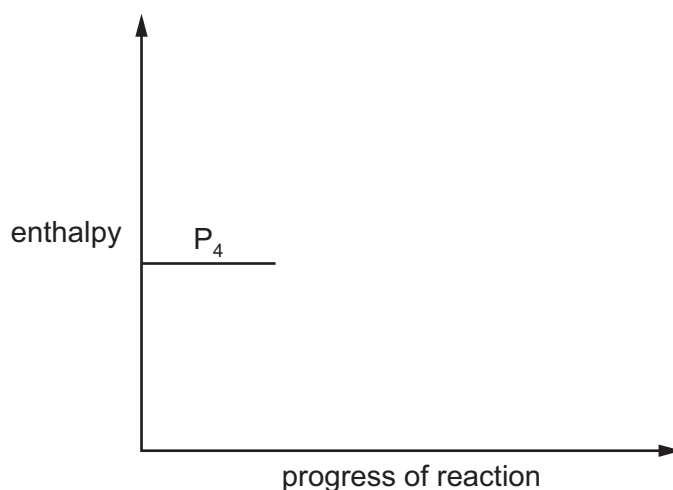
.....

[3]

(ii) Red phosphorus (P) forms when white phosphorus (P₄) is exposed to sunlight.

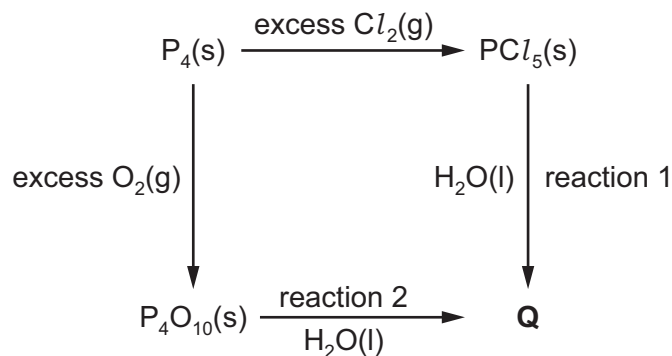


Use this information to draw a reaction pathway diagram to show the formation of red phosphorus (P) from white phosphorus (P₄).



[1]

(b) Some reactions of $P_4(s)$ are shown in the reaction scheme.



(i) State the oxidation number of phosphorus in P_4O_{10} .

..... [1]

(ii) Deduce the identity of **Q** and hence construct chemical equations for reactions 1 and 2.

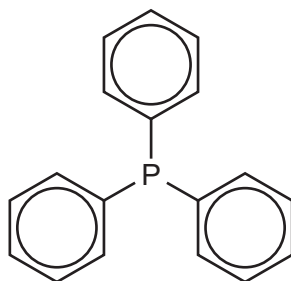
reaction 1 $PCl_5 + \dots H_2O \rightarrow \dots$

reaction 2 $P_4O_{10} + \dots H_2O \rightarrow \dots$

[2]

(c) Triphenylphosphine is used in a type of reaction known as a *Wittig reaction*.

triphenylphosphine

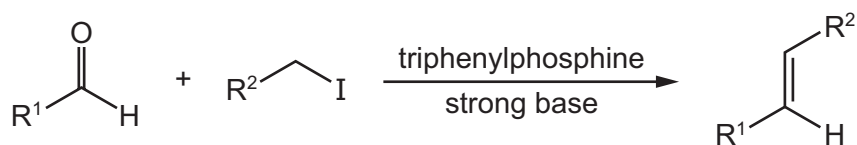


where = $-C_6H_5$

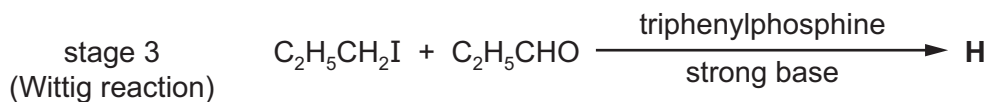
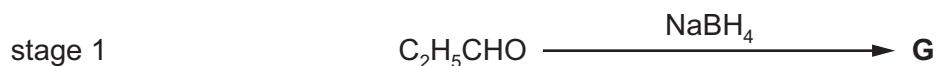
(i) Give the empirical formula of triphenylphosphine.

..... [1]

In a Wittig reaction, an aldehyde reacts with a halogenoalkane to form an alkene. The conversion is shown in the following unbalanced equation.



Compound **H** can be made from propanal, $\text{C}_2\text{H}_5\text{CHO}$. Stage 3 in the reaction scheme is a Wittig reaction.



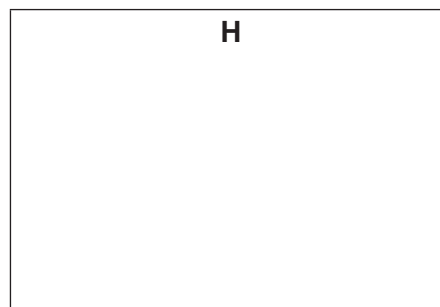
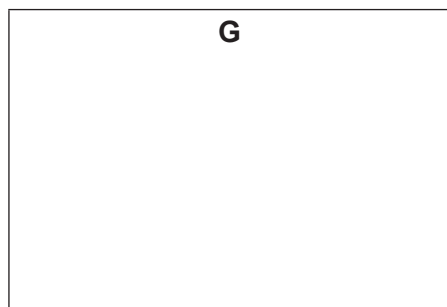
(ii) State the types of reaction that occur in stages 1 and 2.

stage 1

stage 2

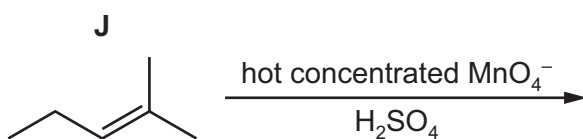
[2]

(iii) Draw the structures of **G** and **H** in the boxes provided.



[2]

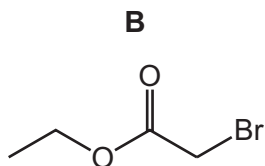
(d) Identify the organic products formed when compound **J**, shown below, is heated with hot concentrated acidified manganate(VII) ions.



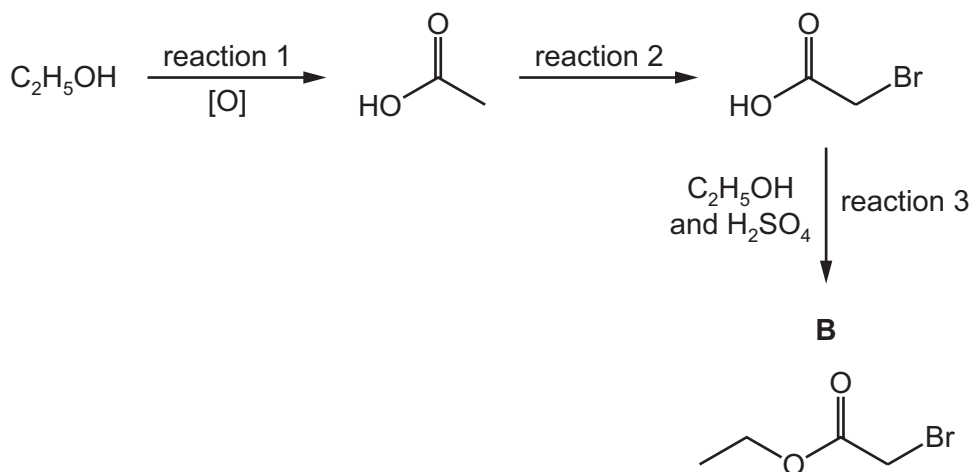
[2]

[Total: 14]

- 4 Compound **B** is a liquid with a fruity smell.



The reaction scheme shows how **B** can be made from ethanol, $\text{C}_2\text{H}_5\text{OH}$.



- (a) (i) Reaction 1 is an oxidation reaction.

Give the reagent(s) and conditions required for reaction 1.

reagent(s)

conditions

[2]

- (ii) Construct an equation to represent reaction 1.

Use [O] to represent an oxygen atom from the oxidising agent in this reaction.

..... [1]

- (iii) Suggest the type of reaction that occurs in reaction 2.

..... [1]

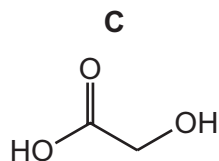
- (iv) H_2SO_4 acts as a homogeneous catalyst in reaction 3.

Explain why H_2SO_4 is described as *homogeneous*.

.....

..... [1]

- (b) Reaction 2 needs to take place in the absence of water to prevent formation of compound **C**.



If **C** is present in the reaction mixture of reaction 3, a different compound, compound **D**, will also form. Compound **D** has two identical functional groups.

The infrared spectrum of **D** shows strong absorptions at 1100 cm^{-1} and 1720 cm^{-1} , but no absorption due to O–H bonds.

Use the *Data Booklet* to identify the functional group present in **D**.

Explain your answer as fully as you can.

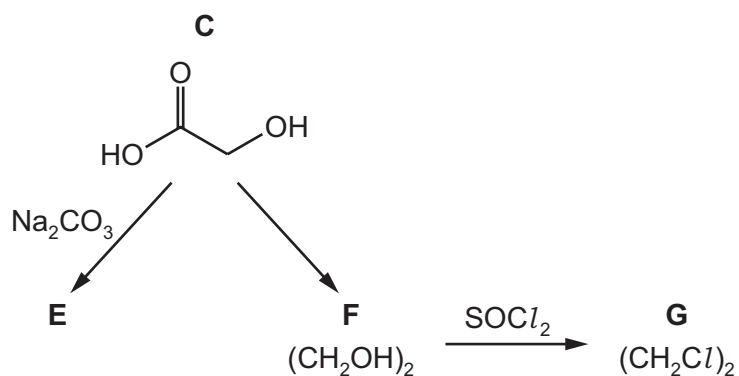
.....

.....

.....

..... [3]

(c) Some other reactions of **C** are shown.



(i) Draw the structure of **E**.

[1]

(ii) Suggest why NaBH_4 is not a suitable reagent to make **F**, $(\text{CH}_2\text{OH})_2$, from **C**. Explain your answer.

.....
 [1]

(iii) Construct an equation for the reaction of $(\text{CH}_2\text{OH})_2$ with SOCl_2 to form **G**, $(\text{CH}_2\text{Cl})_2$.

..... [1]

(d) Explain why **C** is very soluble in water.

.....

 [1]

[Total: 12]

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This document has **12** pages.

Answer **all** the questions in the spaces provided.

1 Hydrogen iodide, HI, is a colourless gas at room temperature.

(a) (i) Explain why HI has a higher boiling point than HCl and HBr.

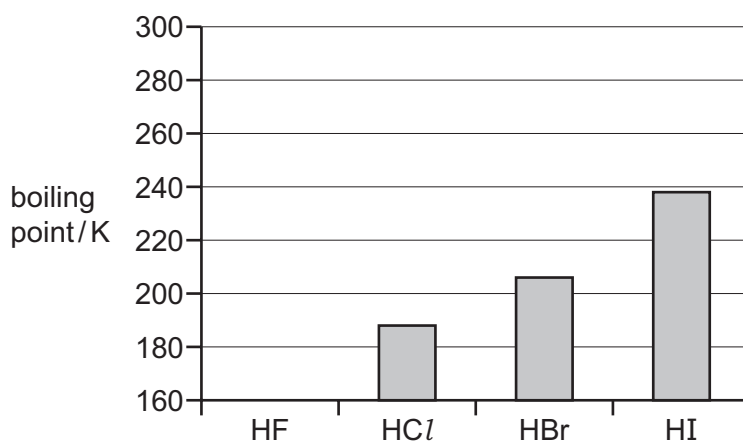
.....

.....

.....

..... [2]

(ii) The bar chart shows the boiling points of HCl, HBr and HI. The boiling point of HF is not shown.



Hydrogen bonds form between HF molecules.

Draw a bar on the bar chart to predict the boiling point of HF.

Explain your answer.

.....

..... [2]

(b) The standard enthalpy change of formation, $\Delta H_f^\ominus$, of HI(g) is $+26.5 \text{ kJ mol}^{-1}$.

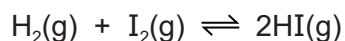
Define the term *standard enthalpy change of formation*.

.....

.....

..... [2]

- (c) HI(g) can be formed by reacting H₂(g) with I₂(g). The reaction is reversible, and an equilibrium forms quickly at high temperatures.



- (i) Construct an expression for the equilibrium constant, K_p , for the reaction of H₂(g) and I₂(g) to form HI(g).

$K_p =$

[1]

- (ii) The equilibrium partial pressures of the gases at 200 °C are as follows.

$$p_{\text{H}_2(\text{g})} = 895 \text{ Pa}$$

$$p_{\text{I}_2(\text{g})} = 895 \text{ Pa}$$

$$p_{\text{HI}(\text{g})} = 4800 \text{ Pa}$$

Calculate K_p for this reaction.

$K_p = \dots\dots\dots$ [1]

- (iii) State how the value of K_p would change, if at all, if the reaction were carried out at 100 °C rather than 200 °C.

Explain your answer.

.....

.....

.....

..... [2]

(d) HI reacts with oxygen to form iodine and water.

(i) Construct an equation for the reaction of HI with oxygen.

..... [1]

(ii) Explain, with reference to oxidation numbers, why this reaction is a redox reaction.

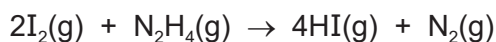
.....

.....

.....

..... [2]

(e) HI(g) can also be formed by the reaction of I₂(g) with hydrazine, N₂H₄(g).



State the change in pressure that would occur when 2 mol I₂(g) fully reacts with 1 mol N₂H₄(g) in a sealed container at constant temperature. Explain your answer.

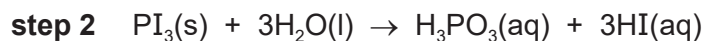
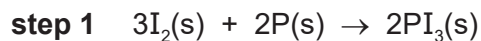
.....

.....

.....

..... [2]

(f) In the laboratory, HI(aq) can be formed in a two-step process.



(i) Draw a 'dot-and-cross' diagram of a PI_3 molecule.

[2]

(ii) Name the type of reaction in **step 2**.

..... [1]

(iii) $\text{H}_3\text{PO}_3(\text{aq})$ and $\text{HI}(\text{aq})$ are both strong Brønsted–Lowry acids.

Give the meaning of the term *strong Brønsted–Lowry acid*.

.....

 [2]

(iv) Give the formula of the conjugate base of H_3PO_3 .

..... [1]

(g) HI(g) reacts with propene, $\text{CH}_3\text{CH}=\text{CH}_2(\text{g})$ to form a mixture of 1-iodopropane and 2-iodopropane.

(i) Identify which of 1-iodopropane and 2-iodopropane is the major product of this reaction.

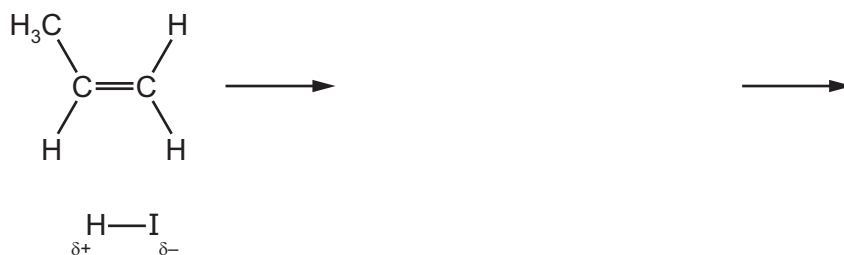
Explain your answer.

.....

 [2]

(ii) Complete the diagram to show the mechanism of the reaction between HI and $\text{CH}_3\text{CH}=\text{CH}_2$ that forms the major product identified in (g)(i).

Include curly arrows, lone pairs of electrons and charges as necessary.



[3]

[Total: 26]

- 2 (a) Table 1 gives physical data for some of the Period 3 elements.

Table 1

atomic number, Z	11	12	13	14	15	16	17
bonding present in element	M						C
first ionisation energy / kJ mol ⁻¹	494	736	577	786	1060	1000	1260
maximum oxidation number							+7
anionic radius / nm	–	–	–	0.271	0.212	0.184	0.181

- (i) Complete the row in the table labelled 'bonding present in element'.

Use C = covalent, I = ionic, M = metallic, as appropriate.

[1]

- (ii) Explain the difference between the first ionisation energies of the elements with atomic numbers 11 and 17.

.....

 [2]

- (iii) Explain the difference between the first ionisation energies of the elements with atomic numbers 15 and 16.

.....

 [2]

- (iv) Complete the row in the table labelled 'maximum oxidation number'.

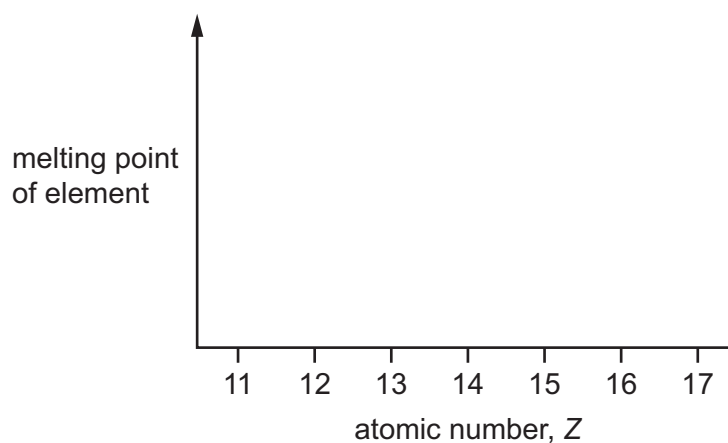
[1]

- (v) Explain the variation in anionic radius for the elements with atomic numbers 14 to 17.

.....

 [2]

- (b) Use the axes to sketch a graph that shows the trend in melting points of the elements with atomic numbers 11 to 17.



[2]

- (c) Dmitri Mendeleev published the first Periodic Table in 1869.

Mendeleev used his knowledge of chemical periodicity to propose the properties of gallium, $_{31}\text{Ga}$, a Group 13 element.

Table 2 gives some chemical and physical data of elements in Group 13.

Table 2

element	density / g cm^{-3}	boiling point /K	cationic radius /nm
$_{5}\text{B}$	2.34	3930	0.020
$_{13}\text{Al}$		2470	0.050
$_{31}\text{Ga}$	5.91	2400	
$_{49}\text{In}$	7.30		0.081
$_{81}\text{Tl}$	11.8	1460	0.095

Complete the table by predicting values for the missing data.

[3]

(d) Indium and aluminium are elements in Group 13 of the Periodic Table.

Indium has very similar chemical properties to aluminium.

- Indium reacts vigorously with hydrochloric acid to form a colourless gas and a salt in solution.
- Indium oxide, In_2O_3 , is amphoteric.
- Gaseous indium bromide has the formula In_2Br_6 . This molecule contains coordinate bonds.

(i) Identify the formula of the salt formed when indium reacts with hydrochloric acid.

..... [1]

(ii) Construct an equation for the reaction of In_2O_3 with excess aqueous NaOH .

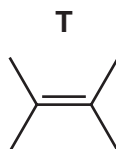
..... [1]

(iii) Draw a diagram that clearly shows the types of bond present in $\text{In}_2\text{Br}_6(\text{g})$.

[2]

[Total: 17]

- 3 Compound **T** is an isomer of C_6H_{12} .



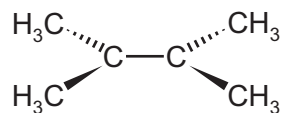
- (a) Name **T**.

..... [1]

- (b) Draw the skeletal formula of a structural isomer of **T** that shows *cis-trans* (geometrical) isomerism.

[1]

- (c) Each carbon atom in **T** forms a sigma (σ) bond to at least one other carbon atom, as shown.

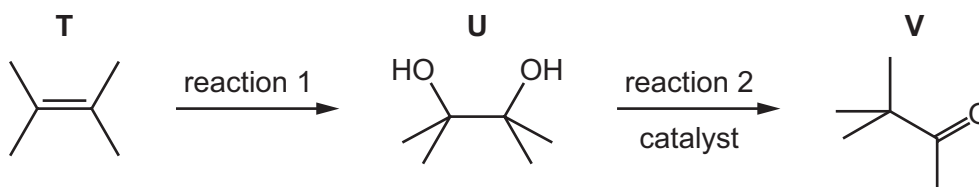


- (i) On the diagram, draw the orbitals that represent the pi (π) bond that is also present in **T**. [1]

- (ii) State the hybridisation of the two carbon atoms between which the pi (π) bond forms.

..... [1]

- (d) A reaction scheme starting with **T** is shown. Reaction 2 occurs in the presence of a catalyst; knowledge of the mechanism for this reaction is not required.



- (i) Give the reagent(s) and conditions for reaction 1.

..... [1]

- (ii) State and explain how 2,4-dinitrophenylhydrazine (2,4-DNPH) can be used to detect the presence of **V** as a product of reaction 2.

.....

 [2]

- (iii) The progress of reaction 2 can be monitored by infrared spectroscopy.

The absorption caused by O–H bonds is always present because water is used as a solvent.

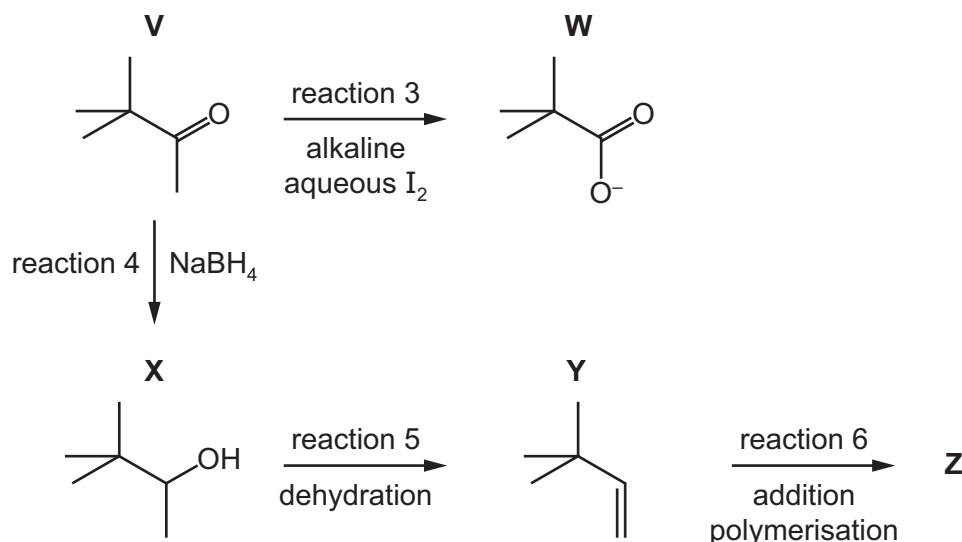
Identify two absorptions, and the bonds responsible for these absorptions, whose appearance will change significantly during the reaction.

1

 2
 [2]

- (e) **V** is used in a wide range of organic reactions.

Some reactions of **V** are shown.



- (i) **V** and **W** are colourless and soluble in water.

State what you would observe in reaction 3.

..... [1]

- (ii) Reaction 3 is a redox reaction.

Identify which of the **reactants** is reduced in this reaction.

..... [1]

- (iii) Construct an equation for reaction 4.

Use [H] in the equation to represent an atom of hydrogen from NaBH₄.

C₆H₁₂O + ... [1]

- (iv) **X** is a mixture of two optical isomers.

Draw the two optical isomers in the boxes provided.



[2]

- (v) Both optical isomers of **X** can be dehydrated to form a single product, **Y**.
Give the reagent(s) and conditions required for reaction 5.

..... [1]

- (vi) **Y** can form an addition polymer **Z**.

Draw one repeat unit of **Z**.

[1]

- (vii) Reaction 6 does not proceed quickly at room temperature.

Suggest why this is the case.

.....

..... [1]

[Total: 17]

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9701/23

October/November 2021

1 hour 15 minutes

You must answer on the question paper.

You will need: Data booklet

- Answer **all** questions.
- Use a black or dark blue pen. You may use an HB pencil for any diagrams or graphs.
- Write your name, centre number and candidate number in the boxes at the top of the page.
- Write your answer to each question in the space provided.
- Do **not** use an erasable pen or correction fluid.
- Do **not** write on any bar codes.
- You may use a calculator.
- You should show all your working, use appropriate units and use an appropriate number of significant figures.

- The total mark for this paper is 60.
- The number of marks for each question or part question is shown in brackets [].

This document has **16** pages. Any blank pages are indicated.

Answer **all** the questions in the spaces provided.

1 Sulfides are compounds that contain sulfur but not oxygen.

(a) Carbon disulfide, CS_2 , is a volatile liquid at room temperature and pressure.

(i) State the meaning of *volatile*.

..... [1]

(ii) Draw a 'dot-and-cross' diagram of the CS_2 molecule.

[2]

(iii) Suggest the bond angle in a molecule of CS_2 .

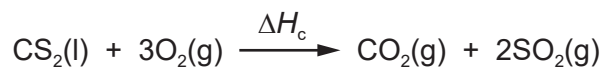
..... [1]

(iv) CS_2 is a liquid under room conditions, while CO_2 is a gas.

Explain what causes the difference in the physical properties between CS_2 and CO_2 .

.....
.....
..... [2]

(b) The enthalpy change of combustion of $\text{CS}_2(\text{l})$ is represented by the following equation.



(i) Define *enthalpy change of combustion*.

.....

 [2]

(ii) The table shows the enthalpy changes of formation of $\text{CS}_2(\text{l})$, $\text{CO}_2(\text{g})$ and $\text{SO}_2(\text{g})$.

compound	enthalpy change of formation, $\Delta H_{\text{f}}/\text{kJ mol}^{-1}$
$\text{CS}_2(\text{l})$	+89.7
$\text{CO}_2(\text{g})$	−394
$\text{SO}_2(\text{g})$	−297

Use the data in the table to calculate the enthalpy change of combustion, ΔH_{c} , of $\text{CS}_2(\text{l})$, in kJ mol^{-1} .

Show your working.

ΔH_{c} of $\text{CS}_2(\text{l})$ = ... kJ mol^{-1}
 [2]

(c) Hydrogen sulfide gas, $\text{H}_2\text{S}(\text{g})$, is slightly soluble in water. It acts as a weak acid in aqueous solution.

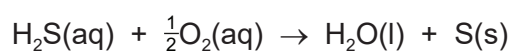
(i) State the meaning of *weak acid*.

.....
..... [1]

(ii) Give the formula of the conjugate base of H_2S .

..... [1]

(iii) $\text{H}_2\text{S}(\text{aq})$ reacts slowly with oxygen dissolved in water. The reaction is represented by the following equation.

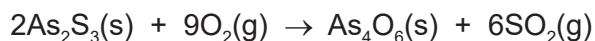


Explain, with reference to oxidation numbers, why this reaction is a redox reaction.

.....
.....
.....
..... [2]

- (d) The compound As_2S_3 is a common mineral.

When As_2S_3 is heated strongly in air, it forms a mixture of products, as shown.



- (i) A sample containing 0.198 g As_2S_3 is placed in 0.100 dm³ of pure oxygen, an excess, in a reaction chamber connected to a gas syringe at room temperature.

The reactants are heated until no further change is observed. The products are then allowed to cool to room temperature.

Calculate the volume, in dm³, of gas present at the end of the experiment.

The molar volume of gas is 24.0 dm³ mol⁻¹ under these conditions. Assume that the pressure is constant throughout the experiment.

Show your working.

volume of gas remaining = ... dm³
[4]

- (ii) State the environmental consequences of releasing $\text{SO}_2(\text{g})$ into the atmosphere.

..... [1]

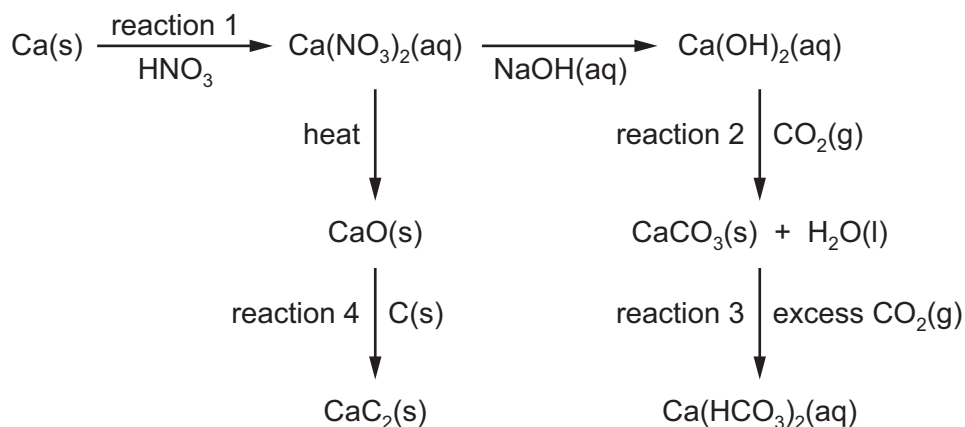
- (iii) $\text{SO}_2(\text{g})$ can be removed from the air by reacting it with $\text{NaOH}(\text{aq})$.

Construct an equation for the reaction of $\text{SO}_2(\text{g})$ with $\text{NaOH}(\text{aq})$. Include state symbols.

..... [2]

[Total: 21]

2 The reaction scheme shows some reactions of calcium.



(a) (i) Reaction 1 produces $\text{Ca(NO}_3)_2$ and one other product.

Identify the other product.

..... [1]

(ii) Construct an equation for the thermal decomposition of $\text{Ca(NO}_3)_2\text{(s)}$.

..... [1]

(iii) State the trend in the thermal stability of the Group 2 nitrates down the group.

..... [1]

(iv) In reaction 3, excess CO_2 is bubbled through water containing CaCO_3 . A solution of $\text{Ca(HCO}_3)_2\text{(aq)}$ forms.

Construct an equation for reaction 3.

..... [1]

(b) Describe how Ca(OH)_2 is used in agriculture.

.....
 [1]

(c) In reaction 4, calcium carbide, CaC_2 , is formed from CaO .

CaC_2 contains the C_2^{2-} anion. Each carbon in C_2^{2-} is sp hybridised.

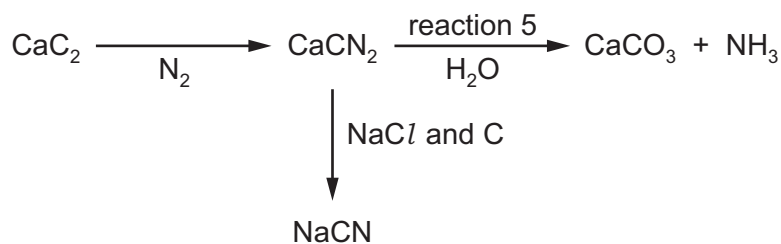
(i) Describe how sp hybridised orbitals are formed.

.....
..... [1]

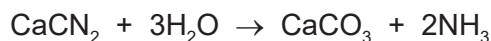
(ii) Sketch a diagram to show how two sp hybrid orbitals can form a sigma (σ) bond.

[2]

(d) The flowchart shows some reactions of CaC_2 .



(i) Reaction 5 can be used to prepare NH_3 .



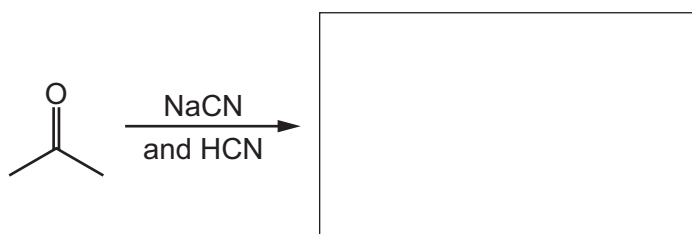
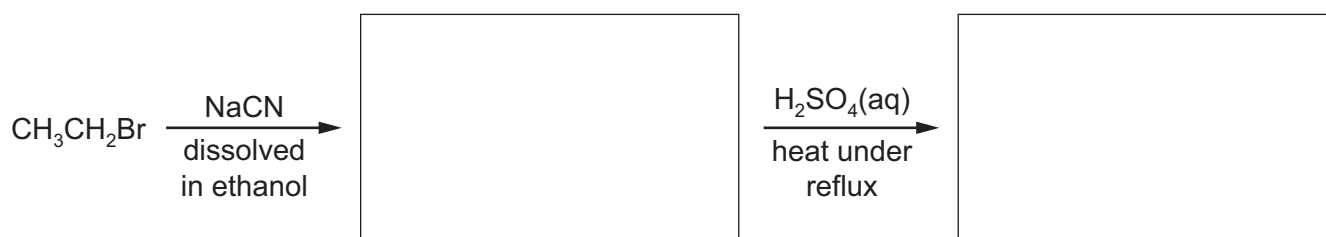
Calculate the minimum mass, in tonnes, of calcium cyanamide, CaCN_2 , that is required to produce 1.50×10^6 tonnes of NH_3 .

Show your working.

$$1 \text{ tonne} = 1.00 \times 10^6 \text{ g}$$

minimum mass of CaCN_2 = ... tonnes
[2]

(ii) Draw the structure of the organic products formed in the following reactions.



[3]

[Total: 13]

3 Phosphorus is a reactive Period 3 element.

(a) Phosphorus has several allotropes. Details of two allotropes are given.

allotrope of phosphorus	formula	melting point/°C
white	P ₄	44
red	P	590

(i) White phosphorus and red phosphorus both have covalent bonding.

Suggest the types of structure shown by white phosphorus (P₄) and red phosphorus (P).

Explain why red phosphorus (P) has a higher melting point than white phosphorus (P₄).

structure of P₄

structure of P

explanation

.....

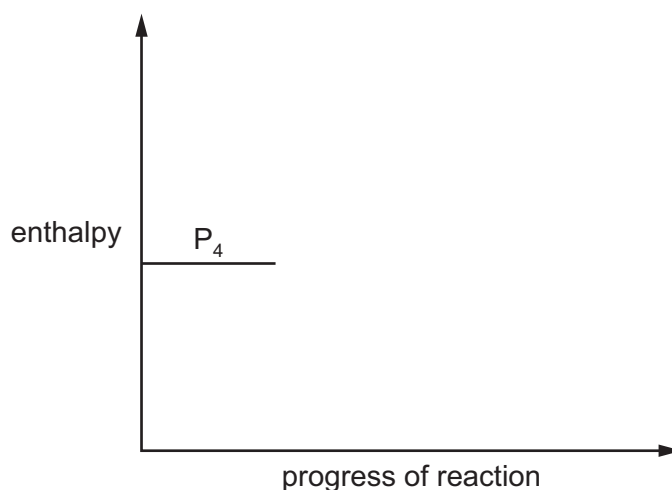
.....

[3]

(ii) Red phosphorus (P) forms when white phosphorus (P₄) is exposed to sunlight.

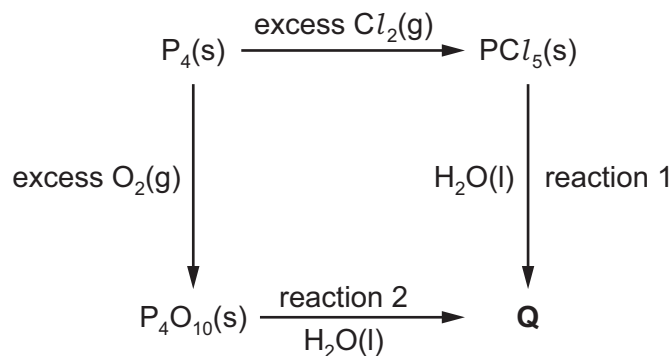


Use this information to draw a reaction pathway diagram to show the formation of red phosphorus (P) from white phosphorus (P₄).



[1]

(b) Some reactions of $P_4(s)$ are shown in the reaction scheme.



(i) State the oxidation number of phosphorus in P_4O_{10} .

..... [1]

(ii) Deduce the identity of **Q** and hence construct chemical equations for reactions 1 and 2.

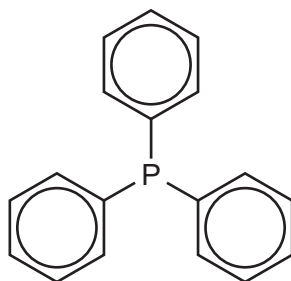
reaction 1 $PCl_5 + \dots H_2O \rightarrow \dots$

reaction 2 $P_4O_{10} + \dots H_2O \rightarrow \dots$

[2]

(c) Triphenylphosphine is used in a type of reaction known as a *Wittig reaction*.

triphenylphosphine

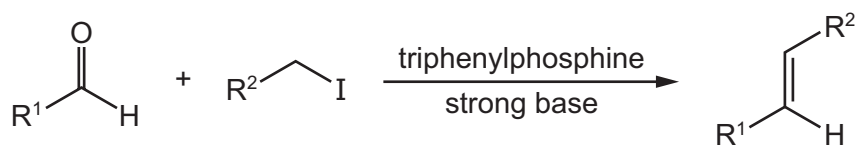


where = $-C_6H_5$

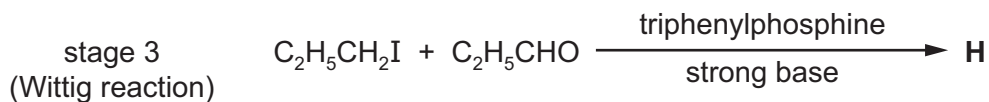
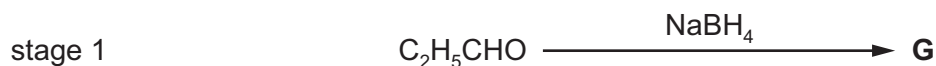
(i) Give the empirical formula of triphenylphosphine.

..... [1]

In a Wittig reaction, an aldehyde reacts with a halogenoalkane to form an alkene. The conversion is shown in the following unbalanced equation.



Compound **H** can be made from propanal, $\text{C}_2\text{H}_5\text{CHO}$. Stage 3 in the reaction scheme is a Wittig reaction.



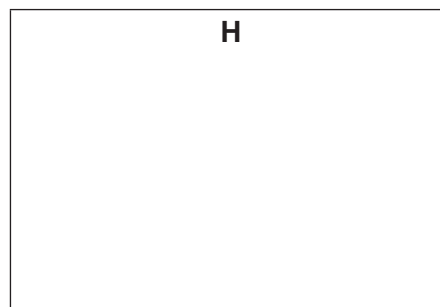
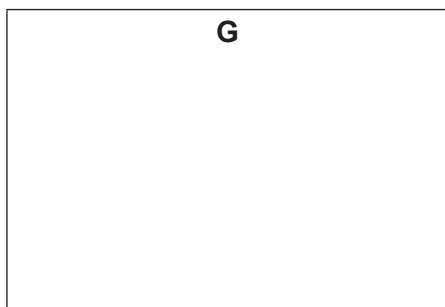
(ii) State the types of reaction that occur in stages 1 and 2.

stage 1

stage 2

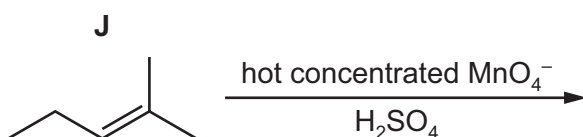
[2]

(iii) Draw the structures of **G** and **H** in the boxes provided.



[2]

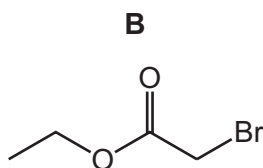
(d) Identify the organic products formed when compound **J**, shown below, is heated with hot concentrated acidified manganate(VII) ions.



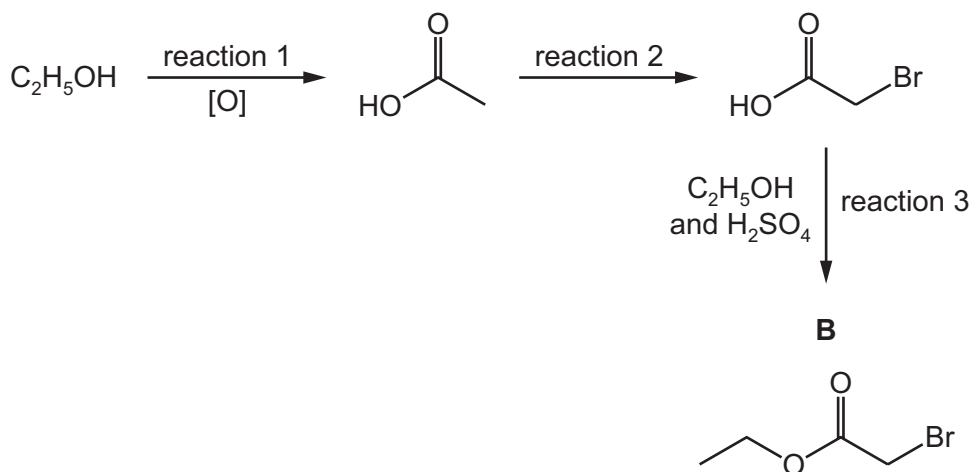
[2]

[Total: 14]

- 4 Compound **B** is a liquid with a fruity smell.



The reaction scheme shows how **B** can be made from ethanol, $\text{C}_2\text{H}_5\text{OH}$.



- (a) (i) Reaction 1 is an oxidation reaction.

Give the reagent(s) and conditions required for reaction 1.

reagent(s)

conditions

[2]

- (ii) Construct an equation to represent reaction 1.

Use [O] to represent an oxygen atom from the oxidising agent in this reaction.

..... [1]

- (iii) Suggest the type of reaction that occurs in reaction 2.

..... [1]

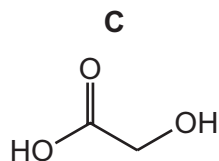
- (iv) H_2SO_4 acts as a homogeneous catalyst in reaction 3.

Explain why H_2SO_4 is described as *homogeneous*.

.....

..... [1]

- (b) Reaction 2 needs to take place in the absence of water to prevent formation of compound **C**.



If **C** is present in the reaction mixture of reaction 3, a different compound, compound **D**, will also form. Compound **D** has two identical functional groups.

The infrared spectrum of **D** shows strong absorptions at 1100 cm^{-1} and 1720 cm^{-1} , but no absorption due to O–H bonds.

Use the *Data Booklet* to identify the functional group present in **D**.

Explain your answer as fully as you can.

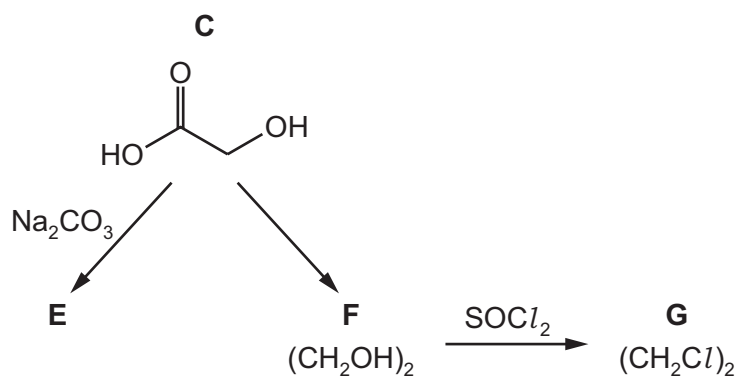
.....

.....

.....

..... [3]

(c) Some other reactions of **C** are shown.



(i) Draw the structure of **E**.

[1]

(ii) Suggest why NaBH_4 is not a suitable reagent to make **F**, $(\text{CH}_2\text{OH})_2$, from **C**. Explain your answer.

.....
 [1]

(iii) Construct an equation for the reaction of $(\text{CH}_2\text{OH})_2$ with SOCl_2 to form **G**, $(\text{CH}_2\text{Cl})_2$.

..... [1]

(d) Explain why **C** is very soluble in water.

.....

 [1]

[Total: 12]

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