Cambridge Assessment International Education

Cambridge International AS & A Level

CHEMISTRY **9701/31**

Paper 3 Advanced Practical Skills 1 **October/November 2020**

CONFIDENTIAL INSTRUCTIONS

This document gives details of how to prepare for and administer the practical exam.

The information in this document and the identity of any materials supplied by Cambridge International are confidential and must NOT reach candidates either directly or indirectly.

The supervisor must complete the report at the end of this document and return it with the scripts.

INSTRUCTIONS

- If you have any queries regarding these confidential instructions, contact Cambridge International stating the centre number, the syllabus and component number and the nature of the query.
  email info@cambridgeinternational.org
  phone +44 1223 553554

This document has **8** pages. Blank pages are indicated.

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General information about practical exams

Centres must follow the guidance on science practical exams given in the *Cambridge Handbook*.

Safety

Supervisors must follow national and local regulations relating to safety and first aid.

Only those procedures described in the question paper should be attempted.

Supervisors must inform candidates that materials and apparatus used in the exam should be treated with caution. Suitable eye protection should be used where necessary.

The following hazard codes are used in these confidential instructions, where relevant:

C	corrosive	**MH**	moderate hazard
HH	health hazard	**T**	acutely toxic
F	flammable	**O**	oxidising
N	hazardous to the aquatic environment		

Hazard data sheets relating to substances used in this exam should be available from your chemical supplier.

Before the exam

- The packets containing the question papers must **not** be opened before the exam.
- It is assumed that standard school laboratory facilities, as indicated in the *Guide to Planning Practical Science*, will be available.
- Spare materials and apparatus for the tasks set must be available for candidates, if required.

During the exam

- It must be made clear to candidates at the start of the exam that they may request spare materials and apparatus for the tasks set.
- Where specified, the supervisor **must** perform the experiments and record the results as instructed. This must be done **out of sight** of the candidates, using the same materials and apparatus as the candidates.
- Any assistance provided to candidates must be recorded in the supervisor's report.
- If any materials or apparatus need to be replaced, for example, in the event of breakage or loss, this must be recorded in the supervisor's report.

After the exam

- The supervisor must complete a report for each practical session held and each laboratory used.
- Each packet of scripts returned to Cambridge International must contain the following items:

  - the scripts of the candidates specified on the bar code label provided
  - the supervisor's results relevant to these candidates
  - the supervisor's reports relevant to these candidates
  - seating plans for each practical session, referring to each candidate by candidate number
  - the attendance register.

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Specific information for this practical exam

During the exam, the supervisor (NOT the invigilator) must do all the experiments and record the results on a spare copy of the question paper, clearly labelled 'supervisor's results'.

If chemicals are prepared in more than one batch, clearly labelled supervisor's results must be provided for each batch. The candidates using each batch must be listed on the supervisor's report.

Apparatus

The apparatus listed must be provided to each candidate.

1 × pipette filler
1 × 250 cm^3 graduated (volumetric) flask
1 × 50 cm^3 burette
1 × 25 cm^3 pipette
1 × 50 cm^3 measuring cylinder
1 × 25 cm^3 measuring cylinder
1 × burette stand and clamp
1 × funnel for filling burette
1 × 150 cm^3 or 250 cm^3 conical flask
1 × white tile
1 × 100 cm^3 beaker
1 × 250 cm^3 beaker
1 × glass rod
6 × test-tube*
1 × boiling tube*
1 × test-tube rack
2 × teat/dropping pipette
1 × stop-clock or sight of a clock with seconds display
wash bottle containing distilled water
pen for labelling glassware
paper towels
access to a balance weighing to **at least** 0.1g
red and blue litmus papers
aluminium foil for testing nitrate/nitrite
wooden splints
the apparatus normally used in the centre for use with limewater in testing for carbon dioxide

* Candidates are expected to rinse and reuse test-tubes and boiling tubes where possible. Additional tubes should be available.

Per five candidates
A bucket labelled **Quenching Bath** must be provided.
The bucket must contain 1 dm^3 of 5% sodium carbonate solution (made up by dissolving 50 g of Na_2CO_3 or 135 g of $Na_2CO_3 \cdot 10H_2O$ in 1 dm^3 of water) and Universal Indicator.
The supervisor must monitor the colour of the Universal Indicator in each quenching bath to check that the solution has not become acidic. If the solution becomes acidic, the supervisor must add more 5% sodium carbonate solution to the quenching bath.

Where balance provision is limited, some candidates should be instructed to start the exam with different questions. See the current syllabus for balance : candidate ratio.

Materials

The materials listed in the table must be provided to each candidate.

Warning: small amounts of SO_2 **[C][T]**, which can cause respiratory distress in some people, may be produced. **The laboratory must be well ventilated.**

label	per candidate	identity	notes
FA 1	6.1–6.3 g	hydrated sodium thiosulfate	Provide 6.20 ± 0.1 g of $Na_2S_2O_3 \cdot 5H_2O$ in a stoppered container.
FA 3	150 cm^3	0.0500 mol dm^{-3} iodine	Dissolve 40.0 g of KI and 12.7 g of I_2 **[MH][N]** in each dm^3 of solution.
FA 4	60 cm^3	2.00 mol dm^{-3} hydrochloric acid	See preparation instructions in the current syllabus.
FA 5	80 cm^3	0.100 mol dm^{-3} sodium thiosulfate	Dissolve 24.8 g of $Na_2S_2O_3 \cdot 5H_2O$ in each dm^3 of solution.
FA 6 [C][MH]	0.4–0.6 g	sodium sulfite	Provide 0.5 ± 0.1 g of Na_2SO_3 **[C][MH]** in a stoppered bottle. A freshly opened supply should be used.
FA 8 [C][N]	10 cm^3	0.25 mol dm^{-3} copper(II) sulfate	Dissolve 15.6 g of $CuSO_4 \cdot 5H_2O$ **[C][N][MH]** in each dm^3 of solution.
aqueous hydrogen peroxide [MH]	10 cm^3	20 'vol' hydrogen peroxide	Dilute 200 cm^3 of 100 'vol' (8.3 mol dm^{-3}) H_2O_2 **[C]** to 1 dm^3.
aqueous potassium iodide	10 cm^3	1.00 mol dm^{-3} potassium iodide	Dissolve 166 g of KI in each dm^3 of solution.
distilled water	400 cm^3	distilled water	
starch indicator	10 cm^3	2% aqueous starch	See preparation instructions in the current syllabus.

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label	per candidate	identity	notes
dilute hydrochloric acid	10 cm^3	2.0 mol dm^{-3} HCl	See preparation instructions in the current syllabus. If necessary, each of these reagents can be provided as a communal supply for groups of up to 6 candidates. Invigilators must be alert to the risk of contamination and the opportunity for malpractice when using a communal supply.
dilute nitric acid **[C]**	10 cm^3	2.0 mol dm^{-3} HNO_3	
dilute sulfuric acid **[MH]**	10 cm^3	1.0 mol dm^{-3} H_2SO_4	
aqueous ammonia **[C][MH][N]**	10 cm^3	2.0 mol dm^{-3} NH_3	
aqueous sodium hydroxide **[C]**	10 cm^3	2.0 mol dm^{-3} NaOH	
aqueous barium chloride or aqueous barium nitrate	10 cm^3	0.1 mol dm^{-3} $BaCl_2$ or 0.1 mol dm^{-3} $Ba(NO_3)_2$	
limewater **[MH]**	10 cm^3	saturated aqueous calcium hydroxide, $Ca(OH)_2$	
aqueous silver nitrate	10 cm^3	0.05 mol dm^{-3} $AgNO_3$	
aqueous acidified potassium manganate(VII) **[MH]**	10 cm^3	0.01 mol dm^{-3} $KMnO_4$ in 0.5 mol dm^{-3} H_2SO_4	

- An excess of at least 10% of each material must be prepared to cover accidental loss.
- All solutions must be thoroughly mixed.
- If you are unable to source any of these chemicals, you must contact Cambridge International as far as possible in advance of the exam for advice.
- **Materials must be labelled only as specified in the 'label' column. The identities of chemicals labelled with letter codes, e.g. FA 1, may be different from their descriptions in the question paper. Candidates must use the descriptions given in the question paper.**

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Supervisor's report

Syllabus and component number | | | | | / | | |

Centre number | | | | | |

Centre name .....

Time of the practical session .....

Laboratory name/number .....

Give details of any difficulties experienced by the centre or by candidates (include the relevant candidate names and candidate numbers).

You must include:

- any difficulties experienced by the centre in the preparation of materials
- any difficulties experienced by candidates, e.g. due to faulty materials or apparatus
- any specific assistance given to candidates.

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If chemicals have been prepared in more than one batch, list the candidates using each batch.

Declaration

1 Each packet that I am returning to Cambridge International contains the following items:

- ☐ the scripts of the candidates specified on the bar code label provided
- ☐ the supervisor's results relevant to these candidates
- ☐ the supervisor's reports relevant to these candidates
- ☐ seating plans for each practical session, referring to each candidate by candidate number
- ☐ the attendance register.

2 Where the practical exam has taken place in more than one practical session, I have clearly labelled the supervisor's results, supervisor's reports and seating plans with the time and laboratory name/number for each practical session.

3 I have included details of difficulties relating to each practical session experienced by the centre or by candidates.

4 I have reported any other adverse circumstances affecting candidates, e.g. illness, bereavement or temporary injury, directly to Cambridge International on a *special consideration form*.

Signed ..... (supervisor)

Name (in block capitals) .....

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Cambridge Assessment International Education

Cambridge International AS & A Level

CHEMISTRY **9701/33**

Paper 3 Advanced Practical Skills 1 **October/November 2020**

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If chemicals are prepared in more than one batch, clearly labelled supervisor's results must be provided for each batch. The candidates using each batch must be listed on the supervisor's report.

Apparatus

The apparatus listed must be provided to each candidate.

2 × 50 cm^3 burette (one should be labelled **FA 1**)
1 × 25 cm^3 measuring cylinder
1 × 50 cm^3 measuring cylinder
2 × burette stand and clamp
2 × funnel (for filling burette)
1 × 100 cm^3 beaker
1 × white tile
1 × stop-clock (to measure to an accuracy of 1 s)
1 × Bunsen burner
1 × heat-proof mat
1 × glass rod
1 × spatula
6 × test-tube*
1 × hard-glass test-tube
1 × boiling tube
1 × test-tube rack
1 × test-tube holder
2 × teat/dropping pipette
1 × wash bottle containing distilled water
1 × pen (for labelling glassware)
paper towels
red and blue litmus papers
aluminium foil for testing nitrate/nitrite
wooden splints
the apparatus normally used in the centre for use with limewater in testing for carbon dioxide

*Candidates are expected to rinse and reuse tubes where possible.
Additional tubes should be available.

Materials

The materials listed in the table must be provided to each candidate.

Warning: small amounts of SO_2 **[C][T]**, which can cause respiratory distress in some people, may be produced. **The laboratory must be well ventilated.**

label	per candidate	identity	notes
FA 1	100 cm^3	0.0500 mol dm^{-3} potassium iodide	Dissolve 8.30 g KI in each dm^3 of solution.
FA 2	75 cm^3	0.0500 mol dm^{-3} acidified iron(III) chloride	Dissolve 8.12 g of $FeCl_3$ **[MH][C]** or 13.52 g of $FeCl_3 \cdot 6H_2O$ **[MH][C]** in 500 cm^3 of 2.0 mol dm^{-3} HCl and make up to 1.0 dm^3 of solution with distilled water.
FA 3	150 cm^3	0.00500 mol dm^{-3} sodium thiosulfate	Dissolve 0.79 g of $Na_2S_2O_3$ or 1.24 g of $Na_2S_2O_3 \cdot 5H_2O$ in each dm^3 of solution.
FA 4	75 cm^3	starch indicator	See preparation instructions in the current syllabus.
FA 5 [MH]	2 g	hydrated manganese(II) chloride	Provide 2.0–2.2 g $MnCl_2 \cdot 4H_2O$ **[MH]** in a stoppered container.
FA 6 [MH]	1 g	manganese(IV) oxide	Provide 0.9–1.1 g MnO_2 **[MH]** in a stoppered container.
FA 7	5 cm^3	0.2 mol dm^{-3} potassium iodide	Dissolve 33.2 g KI in each dm^3 of solution.
aqueous hydrogen peroxide	5 cm^3	10-volume hydrogen peroxide	Dilute 100 cm^3 of 100-volume H_2O_2 **[C]** to 1 dm^3 of solution.
distilled water	200 cm^3	distilled water	

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label	per candidate	identity	notes
dilute hydrochloric acid	10 cm^3	2.0 mol dm^{-3} HCl	See preparation instructions in the current syllabus. If necessary, each of these reagents can be provided as a communal supply for groups of up to 6 candidates. Invigilators must be alert to the risk of contamination and the opportunity for malpractice when using a communal supply.
dilute nitric acid **[C]**	10 cm^3	2.0 mol dm^{-3} HNO_3	
dilute sulfuric acid **[MH]**	10 cm^3	1.0 mol dm^{-3} H_2SO_4	
aqueous ammonia **[C][MH][N]**	10 cm^3	2.0 mol dm^{-3} NH_3	
aqueous sodium hydroxide **[C]**	10 cm^3	2.0 mol dm^{-3} NaOH	
aqueous barium chloride or aqueous barium nitrate	10 cm^3	0.1 mol dm^{-3} $BaCl_2$ or 0.1 mol dm^{-3} $Ba(NO_3)_2$	
limewater **[MH]**	10 cm^3	saturated aqueous calcium hydroxide, $Ca(OH)_2$	
aqueous silver nitrate	10 cm^3	0.05 mol dm^{-3} $AgNO_3$	
aqueous acidified potassium manganate(VII) **[MH]**	10 cm^3	0.01 mol dm^{-3} $KMnO_4$ in 0.5 mol dm^{-3} H_2SO_4	

- An excess of at least 10% of each material must be prepared to cover accidental loss.
- All solutions must be thoroughly mixed.
- If you are unable to source any of these chemicals, you must contact Cambridge International as far as possible in advance of the exam for advice.
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Supervisor's report

Syllabus and component number | | | | | / | | |

Centre number | | | | | |

Centre name .....

Time of the practical session .....

Laboratory name/number .....

Give details of any difficulties experienced by the centre or by candidates (include the relevant candidate names and candidate numbers).

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Signed ..... (supervisor)

Name (in block capitals) .....

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Cambridge Assessment International Education

Cambridge International AS & A Level

CHEMISTRY **9701/34**

Paper 3 Advanced Practical Skills 2 **October/November 2020**

CONFIDENTIAL INSTRUCTIONS

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If chemicals are prepared in more than one batch, clearly labelled supervisor's results must be provided for each batch. The candidates using each batch must be listed on the supervisor's report.

Apparatus

The apparatus listed must be provided to each candidate.

1 × pipette filler
1 × 10 cm^3 pipette
2 × 50 cm^3 burette
2 × burette stand and clamp
1 × funnel (for filling burette)
1 × thermometer (–10 °C to +110 °C at 1 °C)
1 × foamed plastic (or cardboard) cup
1 × 250 cm^3 beaker
1 × glass rod
1 × stop-clock (to measure to an accuracy of 1 s)
2 × teat/dropping pipette
1 × spatula
1 × crucible with lid (approximate capacity 15 cm^3)
1 × crucible tongs
1 × pipe-clay triangle
1 × tripod
1 × Bunsen burner
1 × heat-proof mat
8 × test-tube*
2 × boiling tube*
1 × test-tube rack
1 × test-tube holder
1 × wash bottle of distilled water
1 × pen suitable for labelling glassware
paper towels
access to a balance weighing to **at least** 0.1 g
red and blue litmus papers
aluminium foil for testing nitrate/nitrite
wooden splints
apparatus normally used in the centre in testing for carbon dioxide with limewater

*Candidates are expected to rinse and reuse test-tubes and boiling tubes where possible.
Additional tubes should be available.

Where balance provision is limited, some candidates should be instructed to start the exam with different questions. See the current syllabus for balance : candidate ratio.

Materials

The materials listed in the table must be provided to each candidate.

Warning: small amounts of SO_2 **[C][T]**, which can cause respiratory distress in some people, may be produced. **The laboratory must be well ventilated.**

label	per candidate	identity	notes
FB 1 [C]	2.5–2.6 g	hydrated aluminium sulfate	Provide 2.5–2.6 g of $Al_2(SO_4)_3 \cdot 16H_2O$ **[C]** in a stoppered container.
FB 2 [C]	120 cm^3	2.50 mol dm^{-3} sodium hydroxide	Dissolve 100.0 g of NaOH **[C]** in each dm^3 of solution.
FB 3 [C]	100 cm^3	1.95 mol dm^{-3} sulfuric acid	Cautiously pour 107 cm^3 of concentrated (98%) sulfuric acid **[C]** into 500 cm^3 of distilled water with continuous stirring. Make the solution up to 1 dm^3 with distilled water.
FB 4 [N]	20 cm^3	0.2 mol dm^{-3} zinc nitrate	Dissolve 59.5 g of $Zn(NO_3)_2 \cdot 6H_2O$ **[O][MH][N]** in each dm^3 of solution. Provide 20 cm^3 in a stoppered container.
FB 5 [C]	20 cm^3	0.2 mol dm^{-3} sodium sulfite	Dissolve 25.2 g of Na_2SO_3 **[C][MH]** in each dm^3 of solution. This solution should be freshly made. Provide 20 cm^3 in a stoppered container.
aqueous sodium carbonate	10 cm^3	0.5 mol dm^{-3} sodium carbonate	Dissolve 53.0 g of Na_2CO_3 **[MH]** or 143.0 g of $Na_2CO_3 \cdot 10H_2O$ **[MH]** in each dm^3 of water. Provide 10 cm^3 in a suitable container.
distilled water	200 cm^3	distilled water	

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label	per candidate	identity	notes
dilute hydrochloric acid	10 cm^3	2.0 mol dm^{-3} HCl	See preparation instructions in the current syllabus. If necessary, each of these reagents can be provided as a communal supply for groups of up to 6 candidates. Invigilators must be alert to the risk of contamination and the opportunity for malpractice when using a communal supply.
dilute nitric acid **[C]**	10 cm^3	2.0 mol dm^{-3} HNO_3	
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aqueous sodium hydroxide **[C]**	10 cm^3	2.0 mol dm^{-3} NaOH	
aqueous barium chloride or aqueous barium nitrate	10 cm^3	0.1 mol dm^{-3} $BaCl_2$ or 0.1 mol dm^{-3} $Ba(NO_3)_2$	
limewater **[MH]**	10 cm^3	saturated aqueous calcium hydroxide, $Ca(OH)_2$	
aqueous silver nitrate	10 cm^3	0.05 mol dm^{-3} $AgNO_3$	
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Centre number					

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4 I have reported any other adverse circumstances affecting candidates, e.g. illness, bereavement or temporary injury, directly to Cambridge International on a *special consideration form*.

Signed ..... (supervisor)

Name (in block capitals) .....

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Cambridge Assessment International Education

Cambridge International AS & A Level

CHEMISTRY **9701/35**

Paper 3 Advanced Practical Skills 1 **October/November 2020**

CONFIDENTIAL INSTRUCTIONS

This document gives details of how to prepare for and administer the practical exam.

The information in this document and the identity of any materials supplied by Cambridge International are confidential and must NOT reach candidates either directly or indirectly.

The supervisor must complete the report at the end of this document and return it with the scripts.

INSTRUCTIONS

- If you have any queries regarding these confidential instructions, contact Cambridge International stating the centre number, the syllabus and component number and the nature of the query.
  email info@cambridgeinternational.org
  phone +44 1223 553554

This document has **8** pages. Blank pages are indicated.

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General information about practical exams

Centres must follow the guidance on science practical exams given in the *Cambridge Handbook*.

Safety

Supervisors must follow national and local regulations relating to safety and first aid.

Only those procedures described in the question paper should be attempted.

Supervisors must inform candidates that materials and apparatus used in the exam should be treated with caution. Suitable eye protection should be used where necessary.

The following hazard codes are used in these confidential instructions, where relevant:

C	corrosive	**MH**	moderate hazard
HH	health hazard	**T**	acutely toxic
F	flammable	**O**	oxidising
N	hazardous to the aquatic environment		

Hazard data sheets relating to substances used in this exam should be available from your chemical supplier.

Before the exam

- The packets containing the question papers must **not** be opened before the exam.
- It is assumed that standard school laboratory facilities, as indicated in the *Guide to Planning Practical Science*, will be available.
- Spare materials and apparatus for the tasks set must be available for candidates, if required.

During the exam

- It must be made clear to candidates at the start of the exam that they may request spare materials and apparatus for the tasks set.
- Where specified, the supervisor **must** perform the experiments and record the results as instructed. This must be done **out of sight** of the candidates, using the same materials and apparatus as the candidates.
- Any assistance provided to candidates must be recorded in the supervisor's report.
- If any materials or apparatus need to be replaced, for example, in the event of breakage or loss, this must be recorded in the supervisor's report.

After the exam

- The supervisor must complete a report for each practical session held and each laboratory used.
- Each packet of scripts returned to Cambridge International must contain the following items:

  - the scripts of the candidates specified on the bar code label provided
  - the supervisor's results relevant to these candidates
  - the supervisor's reports relevant to these candidates
  - seating plans for each practical session, referring to each candidate by candidate number
  - the attendance register.

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Specific information for this practical exam

During the exam, the supervisor (NOT the invigilator) must do all the experiments and record the results on a spare copy of the question paper, clearly labelled 'supervisor's results'.

If chemicals are prepared in more than one batch, clearly labelled supervisor's results must be provided for each batch. The candidates using each batch must be listed on the supervisor's report.

Apparatus

The apparatus listed must be provided to each candidate.

1 × pipette filler
1 × 250 cm^3 graduated (volumetric) flask
1 × 50 cm^3 burette
1 × 25 cm^3 pipette
1 × 50 cm^3 measuring cylinder
1 × 25 cm^3 measuring cylinder
1 × burette stand and clamp
1 × funnel for filling burette
1 × 150 cm^3 or 250 cm^3 conical flask
1 × white tile
1 × 100 cm^3 beaker
1 × 250 cm^3 beaker
1 × glass rod
6 × test-tube*
1 × boiling tube*
1 × test-tube rack
2 × teat/dropping pipette
1 × stop-clock or sight of a clock with seconds display
wash bottle containing distilled water
pen for labelling glassware
paper towels
access to a balance weighing to **at least** 0.1g
red and blue litmus papers
aluminium foil for testing nitrate/nitrite
wooden splints
the apparatus normally used in the centre for use with limewater in testing for carbon dioxide

* Candidates are expected to rinse and reuse test-tubes and boiling tubes where possible. Additional tubes should be available.

Per five candidates
A bucket labelled **Quenching Bath** must be provided.
The bucket must contain 1 dm^3 of 5% sodium carbonate solution (made up by dissolving 50 g of Na_2CO_3 or 135 g of $Na_2CO_3 \cdot 10H_2O$ in 1 dm^3 of water) and Universal Indicator.
The supervisor must monitor the colour of the Universal Indicator in each quenching bath to check that the solution has not become acidic. If the solution becomes acidic, the supervisor must add more 5% sodium carbonate solution to the quenching bath.

Where balance provision is limited, some candidates should be instructed to start the exam with different questions. See the current syllabus for balance : candidate ratio.

Materials

The materials listed in the table must be provided to each candidate.

Warning: small amounts of SO_2 **[C][T]**, which can cause respiratory distress in some people, may be produced. **The laboratory must be well ventilated.**

label	per candidate	identity	notes
FA 1	6.1–6.3 g	hydrated sodium thiosulfate	Provide 6.20 ± 0.1 g of $Na_2S_2O_3 \cdot 5H_2O$ in a stoppered container.
FA 3	150 cm^3	0.0500 mol dm^{-3} iodine	Dissolve 40.0 g of KI and 12.7 g of I_2 **[MH][N]** in each dm^3 of solution.
FA 4	60 cm^3	2.00 mol dm^{-3} hydrochloric acid	See preparation instructions in the current syllabus.
FA 5	80 cm^3	0.100 mol dm^{-3} sodium thiosulfate	Dissolve 24.8 g of $Na_2S_2O_3 \cdot 5H_2O$ in each dm^3 of solution.
FA 6 [C][MH]	0.4–0.6 g	sodium sulfite	Provide 0.5 ± 0.1 g of Na_2SO_3 **[C][MH]** in a stoppered bottle. A freshly opened supply should be used.
FA 8 [C][N]	10 cm^3	0.25 mol dm^{-3} copper(II) sulfate	Dissolve 15.6 g of $CuSO_4 \cdot 5H_2O$ **[C][N][MH]** in each dm^3 of solution.
aqueous hydrogen peroxide [MH]	10 cm^3	20 'vol' hydrogen peroxide	Dilute 200 cm^3 of 100 'vol' (8.3 mol dm^{-3}) H_2O_2 **[C]** to 1 dm^3.
aqueous potassium iodide	10 cm^3	1.00 mol dm^{-3} potassium iodide	Dissolve 166 g of KI in each dm^3 of solution.
distilled water	400 cm^3	distilled water	
starch indicator	10 cm^3	2% aqueous starch	See preparation instructions in the current syllabus.

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label	per candidate	identity	notes
dilute hydrochloric acid	10 cm^3	2.0 mol dm^{-3} HCl	See preparation instructions in the current syllabus. If necessary, each of these reagents can be provided as a communal supply for groups of up to 6 candidates. Invigilators must be alert to the risk of contamination and the opportunity for malpractice when using a communal supply.
dilute nitric acid **[C]**	10 cm^3	2.0 mol dm^{-3} HNO_3	
dilute sulfuric acid **[MH]**	10 cm^3	1.0 mol dm^{-3} H_2SO_4	
aqueous ammonia **[C][MH][N]**	10 cm^3	2.0 mol dm^{-3} NH_3	
aqueous sodium hydroxide **[C]**	10 cm^3	2.0 mol dm^{-3} NaOH	
aqueous barium chloride or aqueous barium nitrate	10 cm^3	0.1 mol dm^{-3} $BaCl_2$ or 0.1 mol dm^{-3} $Ba(NO_3)_2$	
limewater **[MH]**	10 cm^3	saturated aqueous calcium hydroxide, $Ca(OH)_2$	
aqueous silver nitrate	10 cm^3	0.05 mol dm^{-3} $AgNO_3$	
aqueous acidified potassium manganate(VII) **[MH]**	10 cm^3	0.01 mol dm^{-3} $KMnO_4$ in 0.5 mol dm^{-3} H_2SO_4	

- An excess of at least 10% of each material must be prepared to cover accidental loss.
- All solutions must be thoroughly mixed.
- If you are unable to source any of these chemicals, you must contact Cambridge International as far as possible in advance of the exam for advice.
- **Materials must be labelled only as specified in the 'label' column. The identities of chemicals labelled with letter codes, e.g. FA 1, may be different from their descriptions in the question paper. Candidates must use the descriptions given in the question paper.**

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Supervisor's report

Syllabus and component number | | | | | / | | |

Centre number | | | | | |

Centre name .....

Time of the practical session .....

Laboratory name/number .....

Give details of any difficulties experienced by the centre or by candidates (include the relevant candidate names and candidate numbers).

You must include:

- any difficulties experienced by the centre in the preparation of materials
- any difficulties experienced by candidates, e.g. due to faulty materials or apparatus
- any specific assistance given to candidates.

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If chemicals have been prepared in more than one batch, list the candidates using each batch.

Declaration

1 Each packet that I am returning to Cambridge International contains the following items:

- ☐ the scripts of the candidates specified on the bar code label provided
- ☐ the supervisor's results relevant to these candidates
- ☐ the supervisor's reports relevant to these candidates
- ☐ seating plans for each practical session, referring to each candidate by candidate number
- ☐ the attendance register.

2 Where the practical exam has taken place in more than one practical session, I have clearly labelled the supervisor's results, supervisor's reports and seating plans with the time and laboratory name/number for each practical session.

3 I have included details of difficulties relating to each practical session experienced by the centre or by candidates.

4 I have reported any other adverse circumstances affecting candidates, e.g. illness, bereavement or temporary injury, directly to Cambridge International on a *special consideration form*.

Signed ..... (supervisor)

Name (in block capitals) .....

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Cambridge Assessment International Education

Cambridge International AS & A Level

CHEMISTRY **9701/36**

Paper 3 Advanced Practical Skills 2 **October/November 2020**

CONFIDENTIAL INSTRUCTIONS

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General information about practical exams

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Safety

Supervisors must follow national and local regulations relating to safety and first aid.

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Supervisors must inform candidates that materials and apparatus used in the exam should be treated with caution. Suitable eye protection should be used where necessary.

The following hazard codes are used in these confidential instructions, where relevant:

C	corrosive	**MH**	moderate hazard
HH	health hazard	**T**	acutely toxic
F	flammable	**O**	oxidising
N	hazardous to the aquatic environment		

Hazard data sheets relating to substances used in this exam should be available from your chemical supplier.

Before the exam

- The packets containing the question papers must **not** be opened before the exam.
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During the exam

- It must be made clear to candidates at the start of the exam that they may request spare materials and apparatus for the tasks set.
- Where specified, the supervisor **must** perform the experiments and record the results as instructed. This must be done **out of sight** of the candidates, using the same materials and apparatus as the candidates.
- Any assistance provided to candidates must be recorded in the supervisor's report.
- If any materials or apparatus need to be replaced, for example, in the event of breakage or loss, this must be recorded in the supervisor's report.

After the exam

- The supervisor must complete a report for each practical session held and each laboratory used.
- Each packet of scripts returned to Cambridge International must contain the following items:

  - the scripts of the candidates specified on the bar code label provided
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Specific information for this practical exam

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If chemicals are prepared in more than one batch, clearly labelled supervisor's results must be provided for each batch. The candidates using each batch must be listed on the supervisor's report.

Apparatus

The apparatus listed must be provided to each candidate.

1 × 50 cm^3 measuring cylinder
1 × 250 cm^3 measuring cylinder
1 × side-arm conical flask, labelled **X**, with bung connected to approximately 50 cm of plastic/rubber delivery tube or 150 cm^3 or 250 cm^3 conical flask, labelled **X**, with one-hole bung connected to approximately 50 cm of plastic/rubber delivery tube
1 × tub suitable for acting as trough (minimum capacity 1 dm^3)
2 × stand, clamp and boss
1 × 250 cm^3 beaker
1 × thermometer (–10 °C to +110 °C at 1 °C)
2 × foamed plastic (or cardboard) cup
1 × 25 cm^3 measuring cylinder
1 × Bunsen burner
1 × heat-proof mat
1 × glass rod
8 × test-tube*
1 × test-tube rack
2 × teat/dropping pipette
1 × wash bottle containing distilled water
1 × pen (for labelling glassware)
paper towels
access to a balance weighing to **at least** 0.1 g
red and blue litmus papers
wooden splints
apparatus normally used in the centre in testing for carbon dioxide with limewater
aluminium foil for testing nitrate/nitrite

*Candidates are expected to rinse and reuse test-tubes where possible.
Additional tubes should be available.

Where balance provision is limited, some candidates should be instructed to start the exam with different questions. See the current syllabus for balance : candidate ratio.

Materials

The materials listed in the table must be provided to each candidate.

label	per candidate	identity	notes
FB 1 [MH]	2.50 g	hydrated sodium carbonate	Provide 2.50 ± 0.05 g of $Na_2CO_3 \cdot 10H_2O$ **[MH]** in a stoppered container.
FB 2	80 cm^3	0.50 mol dm^{-3} hydrochloric acid	Dilute 250 cm^3 of 2.0 mol dm^{-3} HCl to 1.0 dm^3 of solution with distilled water.
FB 3 [MH]	2.2 g	anhydrous sodium carbonate	Provide 2.2 ± 0.1 g of Na_2CO_3 **[MH]** in a stoppered container.
FB 4 [MH]	5.5 g	hydrated sodium carbonate	Provide 5.5 ± 0.1 g of $Na_2CO_3 \cdot 10H_2O$ **[MH]** in a stoppered container.
FB 5	70 cm^3	2.0 mol dm^{-3} hydrochloric acid	See preparation instructions in the current syllabus.
FB 6	15 cm^3	0.5 mol dm^{-3} sodium carbonate	Dissolve 143.0 g of $Na_2CO_3 \cdot 10H_2O$ **[MH]** or 53.0 g Na_2CO_3 **[MH]** in each dm^3 of solution.
FB 7	15 cm^3	0.1 mol dm^{-3} chromium(III) chloride	Dissolve 26.7 g of $CrCl_3 \cdot 6H_2O$ **[MH]** in each dm^3 of solution.
FB 8	15 cm^3	0.2 mol dm^{-3} iron(II) sulfate	Dissolve 55.6 g of $FeSO_4 \cdot 7H_2O$ **[MH]** in each dm^3 of solution. This solution must be freshly prepared.

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label	per candidate	identity	notes
dilute hydrochloric acid	10 cm^3	2.0 mol dm^{-3} HCl	See preparation instructions in the current syllabus. If necessary, each of these reagents can be provided as a communal supply for groups of up to 6 candidates. Invigilators must be alert to the risk of contamination and the opportunity for malpractice when using a communal supply.
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Supervisor's report

Syllabus and component number					/		

Centre number					

Centre name .....

Time of the practical session .....

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Signed ..... (supervisor)

Name (in block capitals) .....

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