

CAMBRIDGE INTERNATIONAL EXAMINATIONS

GCE Advanced Level

MARK SCHEME for the October/November 2012 series

9701 CHEMISTRY

9701/51

Paper 5 (Planning, Analysis and Evaluation),
maximum raw mark 30

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

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Page 2	Mark Scheme	Syllabus	Paper
	GCE A LEVEL – October/November 2012	9701	51

Question	Sections	Indicative material	Mark
1 (a)	PLAN Problem	PbO 1:1, Pb ₃ O ₄ 1:1.33, PbO ₂ 1:2 All three correct two marks. Two correct one mark.	[2]
(b)	PLAN Problem	Correctly labelled axes and three straight lines drawn converging at the origin. Correct order of the lines. If 'O' is on the y-axis, order on axes is PbO ₂ (steepest gradient), Pb ₃ O ₄ , PbO. Allow 'Pb' on y-axis, order reversed.	[2]
(c)	PLAN Problem	(i) lead (allow lead oxide or oxide) AND (ii) oxygen (allow O ₂ OR lead)	[1]
(d)	PLAN Method	Diagram shows a heated piece of apparatus containing some lead oxide with hydrogen passing over it with inlet and outlet shown. Diagram shows apparatus to generate hydrogen using Mg/Al/Zn/Fe AND any dilute acid (labelled) OR group 1 metal/alcohol OR Ca with water or dilute acid. Shows excess hydrogen being burned OR led away from apparatus/collected.	[1] [1] [1]
(e)	PLAN Method	Chooses mass (M) of lead oxide between 1 g and 25 g. Re-heats to constant mass. Calculates a volume of hydrogen sufficient to reduce the oxide. (mark is for the method, units are required.) Suggests calculating the moles of Pb and O/mole ratio of Pb to O.	[1] [1] [1] [1]
(f)	Plan Method	Hydrogen is explosive in air, so expel air from the apparatus before lighting flame to burn hydrogen OR lead/lead oxide is harmful/toxic, so wear a mask/use a fume cupboard to prevent inhalation of hydrogen/lead/lead oxide OR acids are corrosive/irritant, use chemically resistant gloves OR reduction tube is hot, allow to cool before handling/use heat resistant gloves/tongs.	[1]
(g)	PLAN Method	Columns are: mass/weight of the oxide; mass/weight of lead; mass/weight of oxygen; (mass units needed for these three) moles of lead; moles of oxygen; (no units). If five/four are fully correct, 2 marks, if only three/two are correct, 1 mark.	[2]
	Total		[15]

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2 (a)	ACE Data	Correct log column heading as $\log C/\log(a-x)/\log(1-B)$.	[1]																																														
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(g)	ACE Conclusion	Draws a straight line from the origin with a different gradient.	[1]
		Shows shorter elapsed times. (Steeper gradient)	[1]
	Total		[15]

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Question	Sections	Expected Answer	Mark
1(a)	PLAN Problem	(i) Pressure increases AND <u>frequency</u> of the collisions increases.	[1]
		(ii) Axes labelled AND graph shows a decrease of volume with increased pressure.	[1]
		(iii) Draws (approx) parallel line or curve above the original line. (At least one of the lines must be unambiguously labelled to identify it.)	[1]
(b)	PLAN Problem	(i) volume	[1]
		(ii) pressure	[1]
			[5]
2(a)	PLAN Method	Diagram shows a heated piece of apparatus containing some solid CuCO_3 alone AND apparatus is air-tight (not lids).	[1]
		Shows how the gas is collected by syringe OR over water/other liquid.	[1]
		(Apparatus is labelled and) the size or capacity of the vessel used to collect the gas produced is shown. (Volume of vessel must be greater or equal to 10 cm^3 , maximum 1000 cm^3 .)	[1]
(b)	PLAN Method	(i) 30 dm^3	[1]
		(ii) 24 dm^3	[1]
		(iii) Calculates the mass of copper carbonate which produces a volume of gas which will fit in the collecting vessel, unit essential. Calculation must be shown and give a mass that would fit in the collecting vessel if decomposition was as given by either equation (2.1 or 2.2, need not be stated). See appendix to mark scheme.	[1]
		(iv) (Reheats) copper carbonate to constant volume of gas.	[1]
		(v) Relates volume of gas collected to the two equations.	[1]
(c)	Plan Method	Harmful by inhalation/injection OR hot reaction vessel (not hot Bunsens).	[1]
		Dispose of CuCO_3 by reacting with ethanoic acid	[1]

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	Total		[10]
3 (a)	ACE Data	One mark for each correctly completed column given to 3 dp excluding the 'mass of water' column. Use table below. If 2 dp used allow 2 marks for 4 columns correct or 1 mark for 2 columns correct. If no columns are completely correct allow 1 mark if at least six values are correct.	[4]

percentage by mass of sulphuric acid	mass of sulphuric acid /g	mass of water /g	volume of sulphuric acid /cm ³	volume of water /cm ³	total volume of 100 g of solution /cm ³	calculated density of the solution /g cm ³	measured density of the solution /g cm ³
0	0.000	100.000	0.000	100.301	100.301	0.997	0.997
10	10.000	90.000	5.476	90.271	95.747	1.044	1.064
20	20.000	80.000	10.953	80.241	91.194	1.097	1.137
30	30.000	70.000	16.429	70.211	86.640	1.154	1.215
40	40.000	60.000	21.906	60.181*	82.087	1.218	1.299
50	50.000	50.000	27.382	50.150*	77.532	1.290*	1.391
60	60.000	40.000	32.859	40.120	72.979	1.370	1.494
70	70.000	30.000	38.335	30.090	68.425	1.461*	1.606
80	80.000	20.000	43.812	20.060	63.872	1.566	1.722
90	90.000	10.000	49.288	10.030	59.318*	1.686	1.809
100	100.000	0.000	54.765	0.000	54.765	1.826	1.826

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(b)	ACE Data	y-axis labelled as 'density /g cm ⁻³ ' and x-axis as '% by mass' of sulfuric acid AND all the plotted points cover at least half the grid in both directions.	[1]
		All 20 points present and correctly plotted.	[1]
		Two labelled continuous curves of best fit that do not deviate to accommodate a misplot or incorrect point. Do not allow points connected by straight lines.	[1]
		Both lines are smooth.	[1]
(c)	ACE	(i) Difference is 0.09, unit necessary AND higher density is the measured density.	[1]
	Evaluation	(ii) Both liquids have hydrogen bonding.	[1]
	ACE Conclusions	Explains difference as change/formation in hydrogen bonding between water and sulfuric acid in the mixture OR the ionisation of sulfuric acid in the mixture.	[1]
(d)	ACE Conclusions	Gives equation:	
		$\frac{40.000 + M}{60.000} = \frac{70.000}{30.000}$ decimal places not required	
		where M is mass of water required.	[1]
		100.000 g of water must be added.	[1]
(e)	ACE Evaluations	Allow inverse of equation or correct use of V.	
		Mass error either 0.01% OR 0.02%.	
		0.228% or 0.456%	[1]
	Total	If no % given a percentage calculation must be seen.	[1]
			[15]

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Appendix

Guide for 2(b)(iii) and 2(b)(v)

Volume of gas collected /cm³	Mass according to equation 2.1 /g	Mass according to equation 2.2 /g
10	0.0412	0.0515
20	0.0823	0.103
30	0.124	0.154
40	0.165	0.206
50	0.206	0.257
60	0.247	0.309
70	0.288	0.360
80	0.329	0.412
90	0.370	0.463
100	0.412	0.515
250	1.029	1.286

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Graphs for 3(b)

