

Cambridge International AS & A Level

CHEMISTRY**9701/41**

Paper 4 A Level Structured Questions

May/June 2025**MARK SCHEME**Maximum Mark: 100

Published

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This document consists of **16** printed pages.

PUBLISHED**Generic Marking Principles**

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- the specific content of the mark scheme or the generic level descriptors for the question
- the specific skills defined in the mark scheme or in the generic level descriptors for the question
- the standard of response required by a candidate as exemplified by the standardisation scripts.

GENERIC MARKING PRINCIPLE 2:

Marks awarded are always **whole marks** (not half marks, or other fractions).

GENERIC MARKING PRINCIPLE 3:

Marks must be awarded **positively**:

- marks are awarded for correct/valid answers, as defined in the mark scheme. However, credit is given for valid answers which go beyond the scope of the syllabus and mark scheme, referring to your Team Leader as appropriate
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- answers should only be judged on the quality of spelling, punctuation and grammar when these features are specifically assessed by the question as indicated by the mark scheme. The meaning, however, should be unambiguous.

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- 4 The error carried forward (ecf) principle should be applied, where appropriate. If an incorrect answer is subsequently used in a scientifically correct way, the candidate should be awarded these subsequent marking points. Further guidance will be included in the mark scheme where necessary and any exceptions to this general principle will be noted.
- 5 'List rule' guidance

For questions that require ***n*** responses (e.g. State **two** reasons ...):
 - The response should be read as continuous prose, even when numbered answer spaces are provided.
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6 Calculation specific guidance

Correct answers to calculations should be given full credit even if there is no working or incorrect working, **unless** the question states 'show your working'.

For questions in which the number of significant figures required is not stated, credit should be awarded for correct answers when rounded by the examiner to the number of significant figures given in the mark scheme. This may not apply to measured values.

For answers given in standard form (e.g. $a \times 10^n$) in which the convention of restricting the value of the coefficient (a) to a value between 1 and 10 is not followed, credit may still be awarded if the answer can be converted to the answer given in the mark scheme.

Unless a separate mark is given for a unit, a missing or incorrect unit will normally mean that the final calculation mark is not awarded. Exceptions to this general principle will be noted in the mark scheme.

7 Guidance for chemical equations

Multiples / fractions of coefficients used in chemical equations are acceptable unless stated otherwise in the mark scheme.

State symbols given in an equation should be ignored unless asked for in the question or stated otherwise in the mark scheme.

Annotations guidance for centres

Examiners use a system of annotations as a shorthand for communicating their marking decisions to one another. Examiners are trained during the standardisation process on how and when to use annotations. The purpose of annotations is to inform the standardisation and monitoring processes and guide the supervising examiners when they are checking the work of examiners within their team. The meaning of annotations and how they are used is specific to each component and is understood by all examiners who mark the component.

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The annotations listed below were available to examiners marking this component in this series.

Annotations

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	Correct point or mark awarded
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	Part of the correct answer has been seen. Full credit has not been awarded.
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Annotation	Meaning
RE	Rounding error
REP	Repetition
SEEN	Blank page or part of script seen
SF	Error in number of significant figures
TE	Transcription error

Question	Answer	Marks
1(a)	$\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$ [1]	1
1(b)	M1: calcium carbonate AND calcium ion / Ca^{2+} smaller / ionic radius bigger down group M2: carbonate ion / anion / CO_3^{2-} is more distorted/polarised by Ca^{2+}	2
1(c)(i)	calcium hydroxide / $\text{Ca}(\text{OH})_2$	1
1(c)(ii)	M1: A / calcium hydroxide has more exothermic ΔH_{latt} AND ΔH_{hyd} than B / barium hydroxide M2: difference in ΔH_{latt} is greater for A / calcium hydroxide than for B / barium hydroxide M3: ΔH_{sol} is less exothermic for A / calcium hydroxide than for B / barium hydroxide	3

Question	Answer	Marks
2(a)(i)	$\text{mol}^{-2} \text{dm}^6 \text{s}^{-1}$	1
2(a)(ii)	M1: $2.57 \times 10^{-6} = 26.4 \times [\text{concentration}]^3$ M2: $[\text{concentration}] = 4.60 \times 10^{-3}$	2
2(a)(iii)	2.57×10^{-3}	1
2(b)(i)	overall 1st order / pseudo 1 st order OR 1st order with respect to Cl_2 AND concentration of NO doesn't change	1
2(b)(ii)	$\ln 2 / 105.6$ or $0.693 / 105.6$	1
2(b)(iii)	$0.0262 / 2.62 \times 10^{-2}$	1

Question	Answer	Marks
2(c)	M1: homogeneous catalyst OR catalyst because it is regenerated M2: $2\text{NO} + \text{O}_2 \rightarrow 2\text{NO}_2$ M3: $\text{NO}_2 + \text{SO}_2 \rightarrow \text{NO} + \text{SO}_3$	3

Question	Answer	Marks
3(a)(i)	M1: $K_{\text{sp}} = [\text{Cr}^{3+}][\text{OH}^-]^3$ M2: $\text{mol}^4 \text{dm}^{-12}$	2
3(a)(ii)	M1: $[\text{OH}^-] = 3[\text{Cr}^{3+}]$ M2: $2.47 \times 10^{-9} \text{mol dm}^{-3}$ M3: $2.54 \times 10^{-7} \text{g dm}^{-3}$	3
3(a)(iii)	common ion effect	1
3(b)(i)	M1: $[\text{H}^+] = 1.23 \times 10^{-3}$ M2: $\text{pH} = 2.91$	2
3(b)(ii)	13.0	1
3(b)(iii)	M1: $[\text{CH}_3\text{CH}_2\text{CH}_2\text{COOH}] = [\text{CH}_3\text{CH}_2\text{CH}_2\text{COO}^-]$ M2: $\text{pH} = 4.82$	2
3(c)	M1: 0.148 g is left in water M2: $K_{\text{pc}} = 3.01$ OR 0.333	2

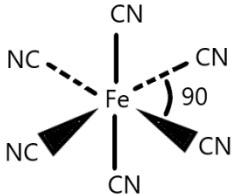
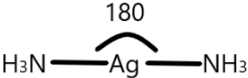
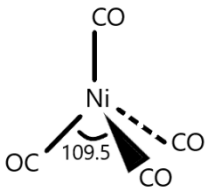
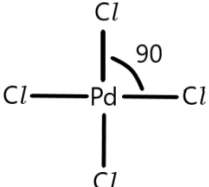
Question	Answer	Marks
4(a)(i)	energy change when one mole of gaseous ions dissolves in water	1
4(a)(ii)	M1: K^+ has larger radius AND K^+ has smaller charge M2: Ca^{2+} has larger attraction for H_2O	2
4(a)(iii)	energy change when one mole of ionic solid forms from gaseous ions	1
4(a)(iv)	M1: $-1650 - (2 \times 506) + 2602$ M2: -60	2
4(b)	M1: $\Delta G = \Delta H - T\Delta S$ AND $T = 298$ M2: -174.5 [1]	2

Question	Answer	Marks
5(a)(i)	(Ar) $3d^{10}$ AND (Ar) $3d^9$	1
5(a)(ii)	similar energy of 3d and 4s subshells	1
5(b)(i)	pale blue precipitate	1
5(b)(ii)	$[Cu(H_2O)_6]^{2+} + 2OH^- \rightarrow Cu(OH)_2(H_2O)_4 + 2H_2O$	1
5(c)(i)	deep blue solution	1
5(c)(ii)	$[Cu(H_2O)_6]^{2+} + 4NH_3 \rightarrow [Cu(NH_3)_4(H_2O)_2]^{2+} + 4H_2O$	1
5(c)(iii)	ligand exchange	1
5(d)(i)	M1: $E = E^\ominus + (0.059 / z) \log(\text{ox/red})$ M2: $E = 1.52 + (0.059 / 5) \times \log(0.002)$	2

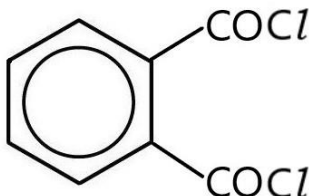
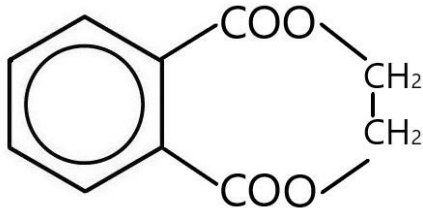
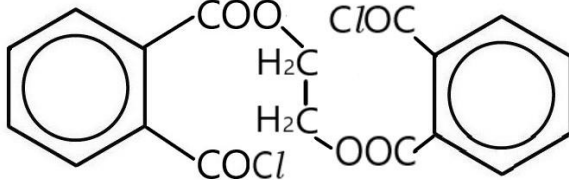
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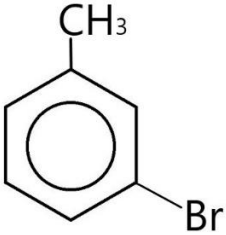
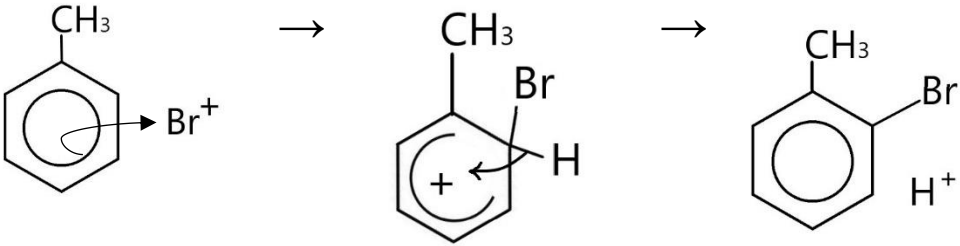
Question	Answer	Marks
5(d)(ii)	1.15	1
5(d)(iii)	$5\text{Cu} + 2\text{MnO}_4^- + 16\text{H}^+ \rightarrow 5\text{Cu}^{2+} + 2\text{Mn}^{2+} + 8\text{H}_2\text{O}$	1
5(d)(iv)	$\text{Cu}^{2+} / \text{Cu}$, $\text{Cu}^{2+} / \text{Cu}$, $\text{MnO}_4^- / \text{Mn}^{2+} / \text{platinum}$	1
5(e)	M1: moles of copper = 0.0120 mol M2: coulombs required = 2322 C M3: current required = 0.129 A	3

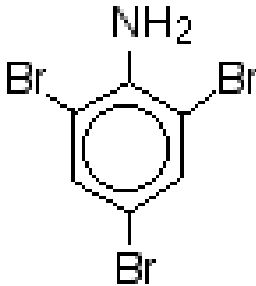
Question	Answer	Marks
6(a)(i)	$[\text{Fe}(\text{CN})_6]^{3-}$ AND $[\text{Fe}(\text{CN})_6]^{4-}$	1
6(a)(ii)	M1: different ΔE M2: absorption of different frequency of visible light OR wavelength of visible light	2
6(a)(iii)	dative covalent / coordinate	1

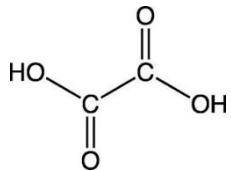
Question	Answer	Marks
6(b)	 	4
	 	
	Each correct structure = [1]	

Question	Answer	Marks																
7(a)	1,3-diethylbenzene	1																
7(b)(i)	CDCl_3	1																
7(b)(ii)	no difference AND no H that can be exchanged with D	1																
7(b)(iii)	<table><tr><th>δ / ppm</th><th>number of protons</th><th>group responsible</th><th>name of splitting pattern</th></tr><tr><td>1.3</td><td>6</td><td>CH_3</td><td>triplet</td></tr><tr><td>2.7</td><td>4</td><td>CH_2</td><td>quartet / quadruplet</td></tr><tr><td>7.1</td><td>4</td><td> </td><td> </td></tr></table> Each correct column = [1]	δ / ppm	number of protons	group responsible	name of splitting pattern	1.3	6	CH_3	triplet	2.7	4	CH_2	quartet / quadruplet	7.1	4			3
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7(b)(iv)	5 AND 6	1																
7(c)(i)	$\text{C}_{10}\text{H}_{14} + 12[\text{O}] \rightarrow \text{C}_8\text{H}_6\text{O}_4 + 2\text{CO}_2 + 4\text{H}_2\text{O}$	1																

Question	Answer		Marks
7(c)(ii)	compound G		3
	compound J		
	compound K		
	Each correct structure = [1]		

Question	Answer	Marks
8(a)(i)	$\text{Br}_2 + \text{AlBr}_3 \rightarrow \text{Br}^+ + \text{AlBr}_4^-$	1
8(a)(ii)	 - reference to methyl group having an effect - methyl group is 2,4,6 directing Correct structure and one correct point [1] correct structure and two correct points [2]	2
8(a)(iii)	 M1: first curly arrow from inside hexagon towards Br^+ M2: correct intermediate M3: second curly arrow from bond into hexagon AND organic product, H^+	3
8(b)(i)	chlorobutane AND white precipitate	1
8(b)(ii)	M1: $\text{C}_4\text{H}_9\text{Cl} + \text{H}_2\text{O} \rightarrow \text{C}_4\text{H}_9\text{OH} + \text{H}^+ + \text{Cl}^-$ M2: $\text{Cl}^- + \text{Ag}^+ \rightarrow \text{AgCl}$ [1]	2

Question	Answer	Marks
8(b)(iii)	M1: delocalisation of lone pair of Cl into benzene ring / π system M2: strengthens $C-Cl$ bond / causes $C-Cl$ bond to have partial double bond nature	2
9(a)(i)	- nitrobenzene / $C_6H_5NO_2$ Sn and HCl - concentrated HCl and heat / boil / reflux Two correct point [1] all three correct points [2]	2
9(a)(ii)	M1: propylamide / propenamide / $C_2H_5CONH_2$ OR propanenitrile / C_2H_5CN M2: $LiAlH_4$	2
9(a)(iii)	M1: bromine (aq) AND phenylamine / $C_6H_5NH_2$  M2: correct structure	2
9(b)	M1: phenylamine ammonia propylamine M2: a base has a lone pair that can accept a proton / H^+ M3: the lone pair on the nitrogen atom of phenylamine is delocalised into the benzene ring / π -system M4: propylamine has an electron donating alkyl group	4

Question	Answer	Marks												
10(a)		1												
10(b)	M1: reagent 1: Fehling's OR Tollens' reagent M2: acidified MnO_4^- <table border="1"> <thead> <tr> <th></th><th>observation when treated with reagent 1</th><th>observation when treated with reagent 2</th></tr> </thead> <tbody> <tr> <td>propanoic acid</td><td>no change</td><td>no change</td></tr> <tr> <td>methanoic acid</td><td>red ppt OR silver mirror</td><td>colourless</td></tr> <tr> <td>ethanedioic acid</td><td>no change</td><td>colourless</td></tr> </tbody> </table> M3 / M4: One completely correct column [1] both completely correct columns [2]		observation when treated with reagent 1	observation when treated with reagent 2	propanoic acid	no change	no change	methanoic acid	red ppt OR silver mirror	colourless	ethanedioic acid	no change	colourless	4
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10(c)	M1: chlorine atoms are electron withdrawing / electronegative M2: weakening O–H bond OR stabilising conjugate base	2												

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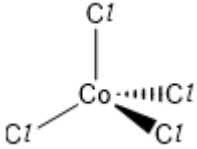
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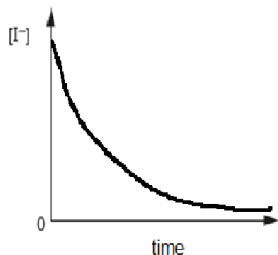
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1(a)(ii)	M1: temperature increases down the group M2: ionic radius OR size of cation / $\text{M}^{(2)+}$ AND increases M3: less polarisation / less distortion AND of anion / nitrate ion OR less weakening of NO bond / harder to break NO bond	3
1(b)	M1: moles of $\text{SrO} = 0.333 \div 103.6 = 3.214 \times 10^{-3}$ AND moles of $\text{Sr}(\text{OH})_2 = 3.214 \times 10^{-3}$ OR 9 / 2800 M2: moles of OH^- in $250 \text{ cm}^3 = 3.214 \times 10^{-3} \times 2 = 6.429 \times 10^{-3}$ AND moles of OH^- in $1 \text{ dm}^3 = 6.429 \times 10^{-3} \times 4 = 0.0257$ M3: $[\text{H}^+] = 1.00 \times 10^{-14} \div 0.0257 = 3.89 \times 10^{-13}$ OR $\text{pOH} = -\log[\text{OH}^-] = 1.59$ M4: $\text{pH} = -\log(3.89 \times 10^{-13}) = 12.41$ must be to 2dp OR $\text{pH} = 14 - 1.59 = 12.41$ must be 2 dp	4
2(a)(i)	A = $[\text{Co}(\text{H}_2\text{O})_6]^{2+}$ pink B = $[\text{Co}(\text{H}_2\text{O})_4(\text{OH})_2] / \text{Co}(\text{OH})_2$ blue C = $[\text{Co}(\text{NH}_3)_6]^{2+}$ straw / yellow-brown B = precipitation / deprotonation / acid-base C = ligand exchange Any two [1] any four [2] any six [3] all eight [4]	4
2(a)(ii)	concentrated HCl	1
2(b)(i)	- five (3d) orbitals (lines, boxes) in the isolated Co^{2+} ion of same energy - splitting three higher and two lower d orbitals - energy of all d orbitals in the complex clearly higher than all d orbitals in isolated ion any two [1] all three [2]	2

Question	Answer	Marks
2(b)(ii)		1
2(c)	M1: H₂O₂ acting as oxidising agent with Mn(OH)₂ / Mn²⁺ M2: Mn²⁺ + H₂O₂ → MnO₂ + 2H⁺ AND E^o_{cell} = +0.56 OR Mn(OH)₂ + H₂O₂ → MnO₂ + 2H₂O AND E^o_{cell} = +0.92 M3: H₂O₂ acting as reducing agent with MnO₂ M4: MnO₂ + 2H⁺ + H₂O₂ → Mn²⁺ + 2H₂O + O₂ AND E^o_{cell} = +0.54	4
2(d)(i)	colourless to pale pink	1
2(d)(ii)	- moles MnO₄[−] = 0.0500 × 18.60 / 1000 = 9.30 × 10^{−4} in 25.0 cm³ - moles Fe²⁺ = 5 × 9.30 × 10^{−4} = 4.65 × 10^{−3} in 25.0 cm³ - moles Fe²⁺ = 6 × 4.65 × 10^{−3} = 2.79 × 10^{−2} in 150.0 cm³ M1 / M2: any two of the above bullets [1] all three [2] OR - moles MnO₄[−] = 0.0500 × 18.60 / 1000 = 9.30 × 10^{−4} in 25.0 cm³ - moles MnO₄[−] = 6 × 9.30 × 10^{−4} = 5.58 × 10^{−3} in 150.0 cm³ - moles Fe²⁺ = 5 × 5.58 × 10^{−3} = 2.79 × 10^{−2} in 150.0 cm³ M1 / M2: any two of the above bullets [1] all three [2] M3: mass Fe = 55.8 × 2.79 × 10^{−2} = 1.55682 (g) in two tablets M4: mass Fe = 0.5 × 1.55682 = 0.77841 (g) in one tablet AND mass Fe = 7.7841 × 10^{−2} × 1000 = 778.4	4

Question	Answer	Marks
3(a)(i)	number of possible arrangements of particles/molecules AND energy in a system	1
3(b)(i)	T_1 melting OR solid to liquid T_2 boiling OR liquid to gas	1
3(b)(ii)	change in disorder from a liquid to a gas is much bigger than solid to liquid OR change in intermolecular distance between liquid & gas is much bigger than solid to liquid	1
3(c)	M1: $\Delta H^\circ_r = (3 \times -393.5) - ((-824.2) + (3 \times -110.5)) = -24.8 \text{ (kJ mol}^{-1}\text{)}$ M2: $\Delta G^\circ = \Delta H^\circ - T\Delta S^\circ$ AND $T = 723$ used M3: $(24800 - 36200) \div 723 = - [3 \times 213.8 + 2 \times 27.3] - (87.4) - (3x)$ $-15.8 = -608.6 + 3x$, so $3x = 592.8$, so $x = 197.6 / 198 \text{ (J K}^{-1} \text{ mol}^{-1}\text{)}$ min 3sf	3
3(d)	(as temperature increases) the reaction is less feasible - as ΔG° becomes less negative / (more) positive - due to $T\Delta S$ becoming more negative / $-T\Delta S$ (becoming more) positive M1 / M2: any two of the above bullets [1] all three [2]	2

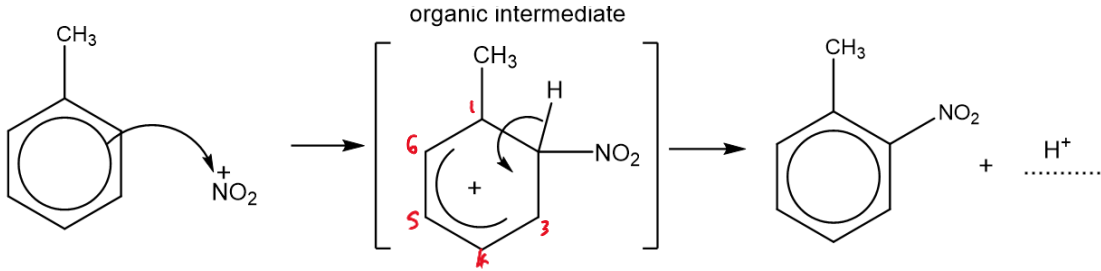
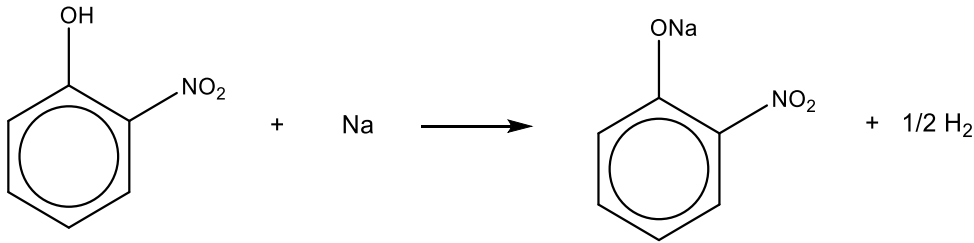
Question	Answer	Marks								
4(a)(i)	the power/exponent to which a concentration AND of a reactant is raised in the rate equation	1								
4(a)(ii)	<table><tr><td>the order of reaction with respect to [H⁺]</td><td>0</td></tr><tr><td>the order of reaction with respect to [I⁻]</td><td>1</td></tr><tr><td>the order of reaction with respect to [H₂O₂]</td><td>1</td></tr><tr><td>overall order of the reaction</td><td>2</td></tr></table> Any two [1] all three [2]	the order of reaction with respect to [H ⁺]	0	the order of reaction with respect to [I ⁻]	1	the order of reaction with respect to [H ₂ O ₂]	1	overall order of the reaction	2	2
the order of reaction with respect to [H ⁺]	0									
the order of reaction with respect to [I ⁻]	1									
the order of reaction with respect to [H ₂ O ₂]	1									
overall order of the reaction	2									

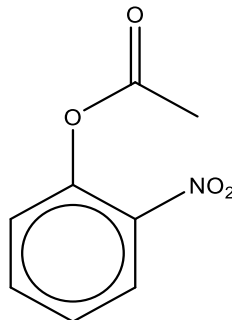
Question	Answer	Marks
4(a)(iii)	 negative gradient curve	1
4(b)(i)	rate = $4 \times 4 = 16$ times	1
4(b)(ii)	slowest step / slow step in the mechanism / multi-step reaction	1
4(b)(iii)	M1: step 1 $\text{O}_3 + \text{NO}_2 \rightarrow \text{NO}_3 + \text{O}_2$ M2: step 2 $\text{NO}_3 + \text{NO}_2 \rightarrow \text{N}_2\text{O}_5$ AND sum of step 1 and 2 is consistent with equation for reaction 2	2
4(c)(i)	Use of graph $k = 12 \times 10^{-4} / 0.04 = 0.03$	1
4(c)(ii)	$k = 0.693 / t_{1/2}$ $t_{1/2} = 0.693 / 0.030 = 23.1 \text{ (s) min 2sf}$	1

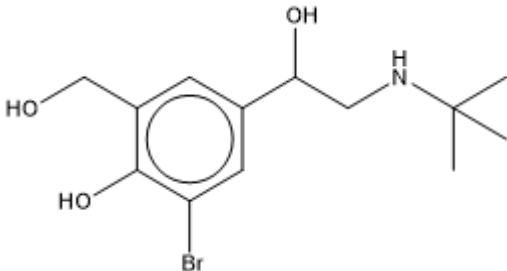
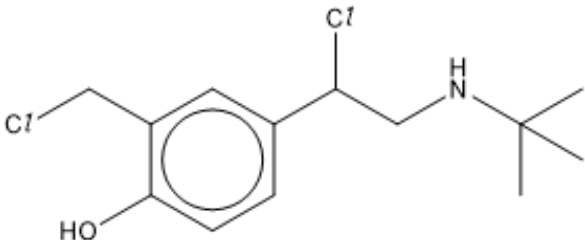
Question	Answer	Marks
5(a)(i)	acid I HClO conjugate base of acid I ClO^- acid II H_2O conjugate base of acid II OH^-	1
(a)(ii)	M1: $[\text{H}^+] = 10^{-4.51} = 3.09 \times 10^{-5}$ M2: $[\text{HClO}] = (3.09 \times 10^{-5})^2 \div (3.70 \times 10^{-8}) = 0.0258 \text{ min 2sf}$	2
(a)(iii)	$3 \text{HClO} \rightarrow \text{HClO}_3 + 2 \text{HCl}$	1

Question	Answer	Marks
(b)(i)	(a solution that) resists / minimises changes pH AND when small amounts of acid / H^+ and base / OH^- are added to it	1
5(b)(ii)	hydroxide ions / NaOH / OH^- AND it reacts with the acid to form the conjugate base / ethanoate ions OR ethanoate ions / sodium ethanoate AND as it is the conjugate base (of ethanoic acid)	1
5(d)(i)	M1: Use of $[Ca^{2+}] = [PO_4^{3-}] = 1.14 \times 10^{-7}$ in the expression OR $K_{sp} = (3.42 \times 10^{-7})^3 \times (2.28 \times 10^{-7})^2$ M2: $K_{sp} = (3 \times 1.14 \times 10^{-7})^3 \times (2 \times 1.14 \times 10^{-7})^2$ $K_{sp} = 2.079 \times 10^{-33}$ OR 2.08×10^{-33} OR 2.1×10^{-33} min 2sf M3: units = $mol^5 dm^{-15}$	3
5(d)(ii)	solubility decreases AND due to the common ion effect	1

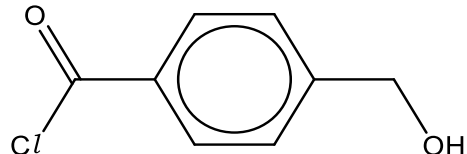
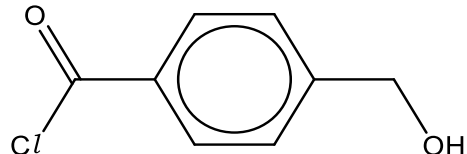
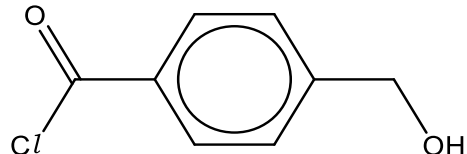
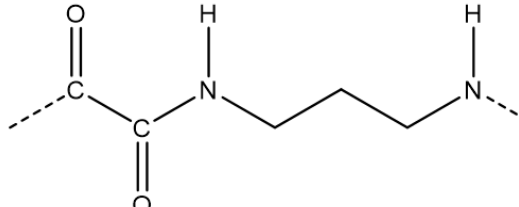
Question	Answer	Marks
6(a)(i)	$HNO_3 + H_2SO_4 \rightarrow NO_2^+ + HSO_4^- + H_2O$ OR $HNO_3 + 2H_2SO_4 \rightarrow NO_2^+ + 2HSO_4^- + H_3O^+$	1

Question	Answer	Marks
6(a)(ii)	 M1: curly arrow 1 from inside the hexagon to nitrogen of NO_2^+ M2: intermediate, usual rules for horseshoe and + charge M3: curly arrow 2 from C-H bond into hexagon AND H^+	3
6(b)	M1: p-orbital/lone pair from O / oxygen AND is delocalised / overlaps AND into the ring / π system M2: so electron density of the ring is increased OR polarises electrophiles / NO_2^+ better	2
6(c)		1
6(d)(i)	(hot) tin / Sn AND concentrated AND HCl	1
6(d)(ii)	reduction	1

Question	Answer	Marks								
6(d)(iii)		1								
6(d)(iv)	(nucleophilic) addition – elimination	1								
6(e)	M1: 2-nitrophenol > phenol > water > ethanol M2: weakens O—H bond / anion stabilised AND proton/H⁺ (more easily) lost / proton / H⁺ (more easily) donated M3: (for 2-nitrophenol / phenol), a p-orbital / lone pair from O / oxygen AND is delocalised / overlaps AND into the ring / π system (and increases electron density in the ring) M4: (for ethanol), positive inductive effect / electron donating of alkyl group	4								
6(f)(i)	<table border="1"><thead><tr><th>name of functional group</th><th>classification of functional group</th></tr></thead><tbody><tr><td>alcohol / hydroxyl</td><td>primary</td></tr><tr><td>alcohol / hydroxyl</td><td>secondary</td></tr><tr><td>amine</td><td>secondary</td></tr></tbody></table> Any three correct [1] all six correct [2]	name of functional group	classification of functional group	alcohol / hydroxyl	primary	alcohol / hydroxyl	secondary	amine	secondary	2
name of functional group	classification of functional group									
alcohol / hydroxyl	primary									
alcohol / hydroxyl	secondary									
amine	secondary									

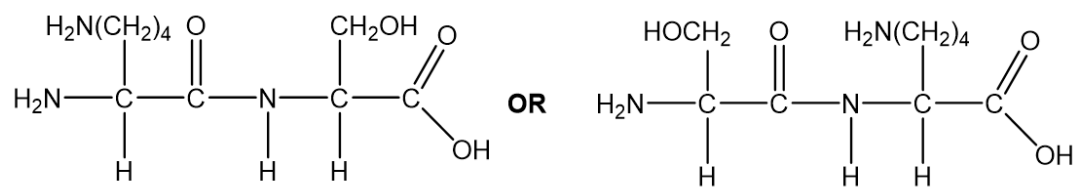
Question	Answer	Marks
6(f)(ii)		1
6(f)(iii)		1

Question	Answer	Marks
7(a)	rotate the plane of polarised light equally in the opposite direction	1
7(b)(i)	M1: nitrogen / N can accept / bond to a proton / H⁺ ion OR donation of lone pair/electrons on the nitrogen / N to a proton / H⁺ ion M2: CH₃CH(NH₂)CH₂CH₃ + H₂O → CH₃CH(NH₃⁺)CH₂CH₃ + OH⁻	2
7(b)(ii)	M1: more electron donating alkyl / ethyl group on N / NH (in diethylamine) OR two electron donating alkyl / R / alkane / ethyl group on N / NH (in diethylamine) M2: increase electron density on N OR make lone pair on N more available (to accept H⁺)	2
7(b)(iii)	(CH ₃ CH ₂) ₂ NH + CH ₃ COOH → (CH ₃ CH ₂) ₂ NH ₂ ⁺ CH ₃ COO ⁻	1

Question	Answer	Marks								
7(b)(iii)	$(\text{CH}_3\text{CH}_2)_2\text{NH} + \text{CH}_3\text{COOH} \rightarrow (\text{CH}_3\text{CH}_2)_2\text{NH}_2^+ \text{CH}_3\text{COO}^-$	1								
7(b)(iv)	$(\text{CH}_3\text{CH}_2)_2\text{NH} + \text{CH}_3\text{COCl} \rightarrow \text{CH}_3\text{CON}(\text{CH}_2\text{CH}_3)_2 + \text{HCl}$	1								
7(c)(i)	<table><tr><td>monomer</td><td>type of polymerisation</td></tr><tr><td>$\text{CH}_2\text{CHCH}_3\text{COOH}$</td><td>addition</td></tr><tr><td></td><td>condensation</td></tr><tr><td>$\text{CH}_3\text{CH}(\text{NH}_2)\text{COOH}$</td><td>condensation</td></tr></table>	monomer	type of polymerisation	$\text{CH}_2\text{CHCH}_3\text{COOH}$	addition		condensation	$\text{CH}_3\text{CH}(\text{NH}_2)\text{COOH}$	condensation	1
monomer	type of polymerisation									
$\text{CH}_2\text{CHCH}_3\text{COOH}$	addition									
	condensation									
$\text{CH}_3\text{CH}(\text{NH}_2)\text{COOH}$	condensation									
7(c)(ii)	 M1: correct displayed amide linkage with an adjacent C=O (to CO) and C(H₂) to NH M2: rest of the structure correct (only one repeat unit) with continuation bonds	2								
7(c)(iii)	they are chemically inert / difficult to hydrolyse / C-C bonds are non-polar	1								

Question	Answer	Marks
8(a)	reference: TMS / tetramethylsilane / $(\text{CH}_3)_4\text{Si}$ AND solvent: D_2O / CDCl_3	1

Question	Answer				Marks
8(b)	six / 6 peaks				1
8(c)	chemical shift δ / ppm	splitting pattern	number of protons on adjacent carbon atoms	number of ^1H atoms responsible for the peak	4
	1.10	triplet	2	3	
	1.50	doublet	1	6	
	2.45	quartet / quadruplet	3	2	
	3.75	multiplet	6	1	
	Any three [1] any six [2] any nine [3] all eleven [4]				

Question	Answer	Marks
9(a)	 M1: correct displayed peptide bond with an adjacent C(H) (to CO) and C(H) to NH M2: rest of the dipeptide correct	2
9(b)(i)	pH at which a molecule has no overall charge / no net charge	1

Question	Answer	Marks
9(b)(ii)	M1: correct relative positions of the spots drawn M2: ser not moved AND as it is a zwitterion / neutral OR ser-lys / lys AND move towards negative (pole) as they are positively charged M3: lys moves the furthest / fastest AND as it has the lower M_r / smaller size (than ser-lys)	3
9(c)(i)	Q, $C_6H_{13}NO_3$	1
9(c)(ii)	M1: step 1 aqueous AND HCl / H_2SO_4 AND heat / reflux[1] M2: step 2 conc. H_2SO_4 catalyst AND high temp / heat	2

Cambridge International AS & A Level

CHEMISTRY**9701/43**

Paper 4 A Level Structured Questions

May/June 2025**MARK SCHEME**Maximum Mark: 100

Published

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge International will not enter into discussions about these mark schemes.

Cambridge International is publishing the mark schemes for the May/June 2025 series for most Cambridge IGCSE, Cambridge International A and AS Level components, and some Cambridge O Level components.

This document consists of **16** printed pages.

PUBLISHED**Generic Marking Principles**

These general marking principles must be applied by all examiners when marking candidate answers. They should be applied alongside the specific content of the mark scheme or generic level descriptions for a question. Each question paper and mark scheme will also comply with these marking principles.

GENERIC MARKING PRINCIPLE 1:

Marks must be awarded in line with:

- the specific content of the mark scheme or the generic level descriptors for the question
- the specific skills defined in the mark scheme or in the generic level descriptors for the question
- the standard of response required by a candidate as exemplified by the standardisation scripts.

GENERIC MARKING PRINCIPLE 2:

Marks awarded are always **whole marks** (not half marks, or other fractions).

GENERIC MARKING PRINCIPLE 3:

Marks must be awarded **positively**:

- marks are awarded for correct/valid answers, as defined in the mark scheme. However, credit is given for valid answers which go beyond the scope of the syllabus and mark scheme, referring to your Team Leader as appropriate
- marks are awarded when candidates clearly demonstrate what they know and can do
- marks are not deducted for errors
- marks are not deducted for omissions
- answers should only be judged on the quality of spelling, punctuation and grammar when these features are specifically assessed by the question as indicated by the mark scheme. The meaning, however, should be unambiguous.

GENERIC MARKING PRINCIPLE 4:

Rules must be applied consistently, e.g. in situations where candidates have not followed instructions or in the application of generic level descriptors.

GENERIC MARKING PRINCIPLE 5:

Marks should be awarded using the full range of marks defined in the mark scheme for the question (however; the use of the full mark range may be limited according to the quality of the candidate responses seen).

GENERIC MARKING PRINCIPLE 6:

Marks awarded are based solely on the requirements as defined in the mark scheme. Marks should not be awarded with grade thresholds or grade descriptors in mind.

Science-Specific Marking Principles

- 1 Examiners should consider the context and scientific use of any keywords when awarding marks. Although keywords may be present, marks should not be awarded if the keywords are used incorrectly.
- 2 The examiner should not choose between contradictory statements given in the same question part, and credit should not be awarded for any correct statement that is contradicted within the same question part. Wrong science that is irrelevant to the question should be ignored.
- 3 Although spellings do not have to be correct, spellings of syllabus terms must allow for clear and unambiguous separation from other syllabus terms with which they may be confused (e.g. ethane / ethene, glucagon / glycogen, refraction / reflection).
- 4 The error carried forward (ecf) principle should be applied, where appropriate. If an incorrect answer is subsequently used in a scientifically correct way, the candidate should be awarded these subsequent marking points. Further guidance will be included in the mark scheme where necessary and any exceptions to this general principle will be noted.
- 5 'List rule' guidance

For questions that require ***n*** responses (e.g. State **two** reasons ...):
 - The response should be read as continuous prose, even when numbered answer spaces are provided.
 - Any response marked *ignore* in the mark scheme should not count towards ***n***.
 - Incorrect responses should not be awarded credit but will still count towards ***n***.
 - Read the entire response to check for any responses that contradict those that would otherwise be credited. Credit should **not** be awarded for any responses that are contradicted within the rest of the response. Where two responses contradict one another, this should be treated as a single incorrect response.
 - Non-contradictory responses after the first ***n*** responses may be ignored even if they include incorrect science.

6 Calculation specific guidance

Correct answers to calculations should be given full credit even if there is no working or incorrect working, **unless** the question states 'show your working'.

For questions in which the number of significant figures required is not stated, credit should be awarded for correct answers when rounded by the examiner to the number of significant figures given in the mark scheme. This may not apply to measured values.

For answers given in standard form (e.g. $a \times 10^n$) in which the convention of restricting the value of the coefficient (a) to a value between 1 and 10 is not followed, credit may still be awarded if the answer can be converted to the answer given in the mark scheme.

Unless a separate mark is given for a unit, a missing or incorrect unit will normally mean that the final calculation mark is not awarded. Exceptions to this general principle will be noted in the mark scheme.

7 Guidance for chemical equations

Multiples / fractions of coefficients used in chemical equations are acceptable unless stated otherwise in the mark scheme.

State symbols given in an equation should be ignored unless asked for in the question or stated otherwise in the mark scheme.

Annotations guidance for centres

Examiners use a system of annotations as a shorthand for communicating their marking decisions to one another. Examiners are trained during the standardisation process on how and when to use annotations. The purpose of annotations is to inform the standardisation and monitoring processes and guide the supervising examiners when they are checking the work of examiners within their team. The meaning of annotations and how they are used is specific to each component and is understood by all examiners who mark the component.

We publish annotations in our mark schemes to help centres understand the annotations they may see on copies of scripts. Note that there may not be a direct correlation between the number of annotations on a script and the mark awarded. Similarly, the use of an annotation may not be an indication of the quality of the response.

The annotations listed below were available to examiners marking this component in this series.

Annotations

Annotation	Meaning
	Correct point or mark awarded
	Incorrect point or mark not awarded
	Unclear
	Information missing or insufficient for credit
	Benefit of the doubt given
	Contradiction in response otherwise markworthy, mark not given
	Part of the correct answer has been seen. Full credit has not been awarded.
	Error carried forward applied
	Incorrect or insufficient point ignored while marking the rest of the response
	Benefit of the doubt not applied in this instance

Annotation	Meaning
RE	Rounding error
REP	Repetition
SEEN	Blank page or part of script seen
SF	Error in number of significant figures
TE	Transcription error

Question	Answer	Marks
1(a)	$\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$ [1]	1
1(b)	M1: calcium carbonate AND calcium ion / Ca^{2+} smaller / ionic radius bigger down group M2: carbonate ion / anion / CO_3^{2-} is more distorted/polarised by Ca^{2+}	2
1(c)(i)	calcium hydroxide / $\text{Ca}(\text{OH})_2$	1
1(c)(ii)	M1: A / calcium hydroxide has more exothermic ΔH_{latt} AND ΔH_{hyd} than B / barium hydroxide M2: difference in ΔH_{latt} is greater for A / calcium hydroxide than for B / barium hydroxide M3: ΔH_{sol} is less exothermic for A / calcium hydroxide than for B / barium hydroxide	3

Question	Answer	Marks
2(a)(i)	$\text{mol}^{-2} \text{dm}^6 \text{s}^{-1}$	1
2(a)(ii)	M1: $2.57 \times 10^{-6} = 26.4 \times [\text{concentration}]^3$ M2: $[\text{concentration}] = 4.60 \times 10^{-3}$	2
2(a)(iii)	2.57×10^{-3}	1
2(b)(i)	overall 1st order / pseudo 1 st order OR 1st order with respect to Cl_2 AND concentration of NO doesn't change	1
2(b)(ii)	$\ln 2 / 105.6$ or $0.693 / 105.6$	1
2(b)(iii)	$0.0262 / 2.62 \times 10^{-2}$	1

Question	Answer	Marks
2(c)	M1: homogeneous catalyst OR catalyst because it is regenerated M2: $2\text{NO} + \text{O}_2 \rightarrow 2\text{NO}_2$ M3: $\text{NO}_2 + \text{SO}_2 \rightarrow \text{NO} + \text{SO}_3$	3

Question	Answer	Marks
3(a)(i)	M1: $K_{\text{sp}} = [\text{Cr}^{3+}][\text{OH}^-]^3$ M2: $\text{mol}^4 \text{dm}^{-12}$	2
3(a)(ii)	M1: $[\text{OH}^-] = 3[\text{Cr}^{3+}]$ M2: $2.47 \times 10^{-9} \text{mol dm}^{-3}$ M3: $2.54 \times 10^{-7} \text{g dm}^{-3}$	3
3(a)(iii)	common ion effect	1
3(b)(i)	M1: $[\text{H}^+] = 1.23 \times 10^{-3}$ M2: $\text{pH} = 2.91$	2
3(b)(ii)	13.0	1
3(b)(iii)	M1: $[\text{CH}_3\text{CH}_2\text{CH}_2\text{COOH}] = [\text{CH}_3\text{CH}_2\text{CH}_2\text{COO}^-]$ M2: $\text{pH} = 4.82$	2
3(c)	M1: 0.148 g is left in water M2: $K_{\text{pc}} = 3.01$ OR 0.333	2

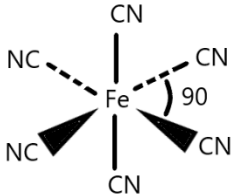
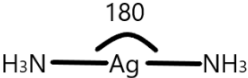
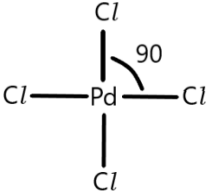
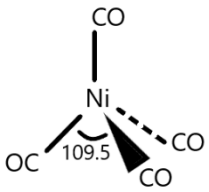
Question	Answer	Marks
4(a)(i)	energy change when one mole of gaseous ions dissolves in water	1
4(a)(ii)	M1: K^+ has larger radius AND K^+ has smaller charge M2: Ca^{2+} has larger attraction for H_2O	2
4(a)(iii)	energy change when one mole of ionic solid forms from gaseous ions	1
4(a)(iv)	M1: $-1650 - (2 \times 506) + 2602$ M2: -60	2
4(b)	M1: $\Delta G = \Delta H - T\Delta S$ AND $T = 298$ M2: -174.5 [1]	2

Question	Answer	Marks
5(a)(i)	(Ar) $3d^{10}$ AND (Ar) $3d^9$	1
5(a)(ii)	similar energy of 3d and 4s subshells	1
5(b)(i)	pale blue precipitate	1
5(b)(ii)	$[Cu(H_2O)_6]^{2+} + 2OH^- \rightarrow Cu(OH)_2(H_2O)_4 + 2H_2O$	1
5(c)(i)	deep blue solution	1
5(c)(ii)	$[Cu(H_2O)_6]^{2+} + 4NH_3 \rightarrow [Cu(NH_3)_4(H_2O)_2]^{2+} + 4H_2O$	1
5(c)(iii)	ligand exchange	1
5(d)(i)	M1: $E = E^\ominus + (0.059 / z) \log(\text{ox/red})$ M2: $E = 1.52 + (0.059 / 5) \times \log(0.002)$	2

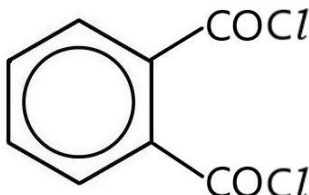
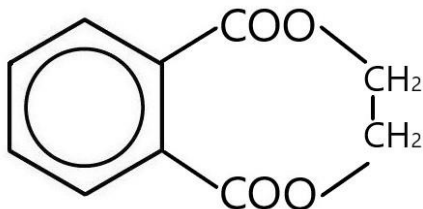
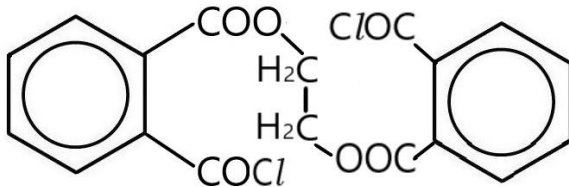
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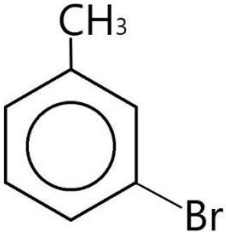
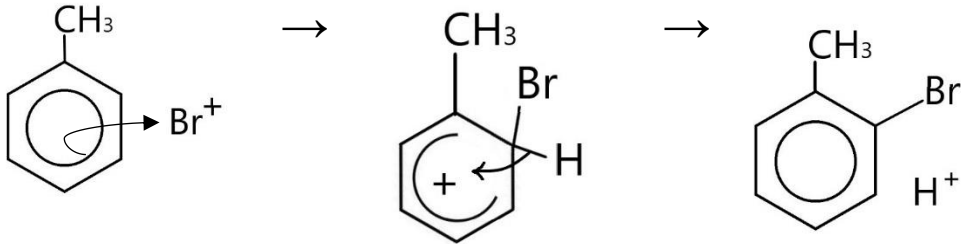
Question	Answer	Marks
5(d)(ii)	1.15	1
5(d)(iii)	$5\text{Cu} + 2\text{MnO}_4^- + 16\text{H}^+ \rightarrow 5\text{Cu}^{2+} + 2\text{Mn}^{2+} + 8\text{H}_2\text{O}$	1
5(d)(iv)	$\text{Cu}^{2+} / \text{Cu}$, $\text{Cu}^{2+} / \text{Cu}$, $\text{MnO}_4^- / \text{Mn}^{2+} / \text{platinum}$	1
5(e)	M1: moles of copper = 0.0120 mol M2: coulombs required = 2322 C M3: current required = 0.129 A	3

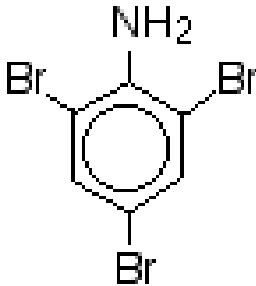
Question	Answer	Marks
6(a)(i)	$[\text{Fe}(\text{CN})_6]^{3-}$ AND $[\text{Fe}(\text{CN})_6]^{4-}$	1
6(a)(ii)	M1: different ΔE M2: absorption of different frequency of visible light OR wavelength of visible light	2
6(a)(iii)	dative covalent / coordinate	1

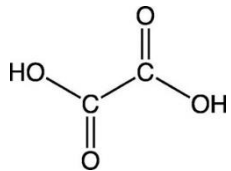
Question	Answer	Marks
6(b)		4
		
		
		
	Each correct structure = [1]	

Question	Answer	Marks																
7(a)	1,3-diethylbenzene	1																
7(b)(i)	CDCl_3	1																
7(b)(ii)	no difference AND no H that can be exchanged with D	1																
7(b)(iii)	<table><tr><th>δ / ppm</th><th>number of protons</th><th>group responsible</th><th>name of splitting pattern</th></tr><tr><td>1.3</td><td>6</td><td>CH_3</td><td>triplet</td></tr><tr><td>2.7</td><td>4</td><td>CH_2</td><td>quartet / quadruplet</td></tr><tr><td>7.1</td><td>4</td><td> </td><td> </td></tr></table> Each correct column = [1]	δ / ppm	number of protons	group responsible	name of splitting pattern	1.3	6	CH_3	triplet	2.7	4	CH_2	quartet / quadruplet	7.1	4			3
δ / ppm	number of protons	group responsible	name of splitting pattern															
1.3	6	CH_3	triplet															
2.7	4	CH_2	quartet / quadruplet															
7.1	4																	
7(b)(iv)	5 AND 6	1																
7(c)(i)	$\text{C}_{10}\text{H}_{14} + 12[\text{O}] \rightarrow \text{C}_8\text{H}_6\text{O}_4 + 2\text{CO}_2 + 4\text{H}_2\text{O}$	1																

Question	Answer		Marks
7(c)(ii)	compound G		3
	compound J		
	compound K		
	Each correct structure = [1]		

Question	Answer	Marks
8(a)(i)	$\text{Br}_2 + \text{AlBr}_3 \rightarrow \text{Br}^+ + \text{AlBr}_4^-$	1
8(a)(ii)	 - reference to methyl group having an effect - methyl group is 2,4,6 directing Correct structure and one correct point [1] correct structure and two correct points [2]	2
8(a)(iii)	 M1: first curly arrow from inside hexagon towards Br^+ M2: correct intermediate M3: second curly arrow from bond into hexagon AND organic product, H^+	3
8(b)(i)	chlorobutane AND white precipitate	1
8(b)(ii)	M1: $\text{C}_4\text{H}_9\text{Cl} + \text{H}_2\text{O} \rightarrow \text{C}_4\text{H}_9\text{OH} + \text{H}^+ + \text{Cl}^-$ M2: $\text{Cl}^- + \text{Ag}^+ \rightarrow \text{AgCl}$ [1]	2

Question	Answer	Marks
8(b)(iii)	M1: delocalisation of lone pair of Cl into benzene ring / π system M2: strengthens $C-Cl$ bond / causes $C-Cl$ bond to have partial double bond nature	2
9(a)(i)	- nitrobenzene / $C_6H_5NO_2$ Sn and HCl - concentrated HCl and heat / boil / reflux Two correct point [1] all three correct points [2]	2
9(a)(ii)	M1: propylamide / propenamide / $C_2H_5CONH_2$ OR propanenitrile / C_2H_5CN M2: $LiAlH_4$	2
9(a)(iii)	M1: bromine (aq) AND phenylamine / $C_6H_5NH_2$  M2: correct structure	2
9(b)	M1: phenylamine ammonia propylamine M2: a base has a lone pair that can accept a proton / H^+ M3: the lone pair on the nitrogen atom of phenylamine is delocalised into the benzene ring / π -system M4: propylamine has an electron donating alkyl group	4

Question	Answer	Marks												
10(a)		1												
10(b)	M1: reagent 1: Fehling's OR Tollens' reagent M2: acidified MnO_4^- <table border="1"> <thead> <tr> <th></th><th>observation when treated with reagent 1</th><th>observation when treated with reagent 2</th></tr> </thead> <tbody> <tr> <td>propanoic acid</td><td>no change</td><td>no change</td></tr> <tr> <td>methanoic acid</td><td>red ppt OR silver mirror</td><td>colourless</td></tr> <tr> <td>ethanedioic acid</td><td>no change</td><td>colourless</td></tr> </tbody> </table> M3 / M4: One completely correct column [1] both completely correct columns [2]		observation when treated with reagent 1	observation when treated with reagent 2	propanoic acid	no change	no change	methanoic acid	red ppt OR silver mirror	colourless	ethanedioic acid	no change	colourless	4
	observation when treated with reagent 1	observation when treated with reagent 2												
propanoic acid	no change	no change												
methanoic acid	red ppt OR silver mirror	colourless												
ethanedioic acid	no change	colourless												
10(c)	M1: chlorine atoms are electron withdrawing / electronegative M2: weakening O–H bond OR stabilising conjugate base	2												

Cambridge International AS & A Level

CHEMISTRY**9701/44**

Paper 4 A Level Structured Questions

May/June 2025**MARK SCHEME**Maximum Mark: 100

Published

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge International will not enter into discussions about these mark schemes.

Cambridge International is publishing the mark schemes for the May/June 2025 series for most Cambridge IGCSE, Cambridge International A and AS Level components, and some Cambridge O Level components.

This document consists of **17** printed pages.

PUBLISHED**Generic Marking Principles**

These general marking principles must be applied by all examiners when marking candidate answers. They should be applied alongside the specific content of the mark scheme or generic level descriptions for a question. Each question paper and mark scheme will also comply with these marking principles.

GENERIC MARKING PRINCIPLE 1:

Marks must be awarded in line with:

- the specific content of the mark scheme or the generic level descriptors for the question
- the specific skills defined in the mark scheme or in the generic level descriptors for the question
- the standard of response required by a candidate as exemplified by the standardisation scripts.

GENERIC MARKING PRINCIPLE 2:

Marks awarded are always **whole marks** (not half marks, or other fractions).

GENERIC MARKING PRINCIPLE 3:

Marks must be awarded **positively**:

- marks are awarded for correct/valid answers, as defined in the mark scheme. However, credit is given for valid answers which go beyond the scope of the syllabus and mark scheme, referring to your Team Leader as appropriate
- marks are awarded when candidates clearly demonstrate what they know and can do
- marks are not deducted for errors
- marks are not deducted for omissions
- answers should only be judged on the quality of spelling, punctuation and grammar when these features are specifically assessed by the question as indicated by the mark scheme. The meaning, however, should be unambiguous.

GENERIC MARKING PRINCIPLE 4:

Rules must be applied consistently, e.g. in situations where candidates have not followed instructions or in the application of generic level descriptors.

GENERIC MARKING PRINCIPLE 5:

Marks should be awarded using the full range of marks defined in the mark scheme for the question (however; the use of the full mark range may be limited according to the quality of the candidate responses seen).

GENERIC MARKING PRINCIPLE 6:

Marks awarded are based solely on the requirements as defined in the mark scheme. Marks should not be awarded with grade thresholds or grade descriptors in mind.

Science-Specific Marking Principles

- 1 Examiners should consider the context and scientific use of any keywords when awarding marks. Although keywords may be present, marks should not be awarded if the keywords are used incorrectly.
- 2 The examiner should not choose between contradictory statements given in the same question part, and credit should not be awarded for any correct statement that is contradicted within the same question part. Wrong science that is irrelevant to the question should be ignored.
- 3 Although spellings do not have to be correct, spellings of syllabus terms must allow for clear and unambiguous separation from other syllabus terms with which they may be confused (e.g. ethane / ethene, glucagon / glycogen, refraction / reflection).
- 4 The error carried forward (ecf) principle should be applied, where appropriate. If an incorrect answer is subsequently used in a scientifically correct way, the candidate should be awarded these subsequent marking points. Further guidance will be included in the mark scheme where necessary and any exceptions to this general principle will be noted.
- 5 'List rule' guidance

For questions that require ***n*** responses (e.g. State **two** reasons ...):
 - The response should be read as continuous prose, even when numbered answer spaces are provided.
 - Any response marked *ignore* in the mark scheme should not count towards ***n***.
 - Incorrect responses should not be awarded credit but will still count towards ***n***.
 - Read the entire response to check for any responses that contradict those that would otherwise be credited. Credit should **not** be awarded for any responses that are contradicted within the rest of the response. Where two responses contradict one another, this should be treated as a single incorrect response.
 - Non-contradictory responses after the first ***n*** responses may be ignored even if they include incorrect science.

6 Calculation specific guidance

Correct answers to calculations should be given full credit even if there is no working or incorrect working, **unless** the question states 'show your working'.

For questions in which the number of significant figures required is not stated, credit should be awarded for correct answers when rounded by the examiner to the number of significant figures given in the mark scheme. This may not apply to measured values.

For answers given in standard form (e.g. $a \times 10^n$) in which the convention of restricting the value of the coefficient (a) to a value between 1 and 10 is not followed, credit may still be awarded if the answer can be converted to the answer given in the mark scheme.

Unless a separate mark is given for a unit, a missing or incorrect unit will normally mean that the final calculation mark is not awarded. Exceptions to this general principle will be noted in the mark scheme.

7 Guidance for chemical equations

Multiples / fractions of coefficients used in chemical equations are acceptable unless stated otherwise in the mark scheme.

State symbols given in an equation should be ignored unless asked for in the question or stated otherwise in the mark scheme.

Annotations guidance for centres

Examiners use a system of annotations as a shorthand for communicating their marking decisions to one another. Examiners are trained during the standardisation process on how and when to use annotations. The purpose of annotations is to inform the standardisation and monitoring processes and guide the supervising examiners when they are checking the work of examiners within their team. The meaning of annotations and how they are used is specific to each component and is understood by all examiners who mark the component.

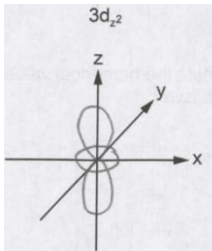
We publish annotations in our mark schemes to help centres understand the annotations they may see on copies of scripts. Note that there may not be a direct correlation between the number of annotations on a script and the mark awarded. Similarly, the use of an annotation may not be an indication of the quality of the response.

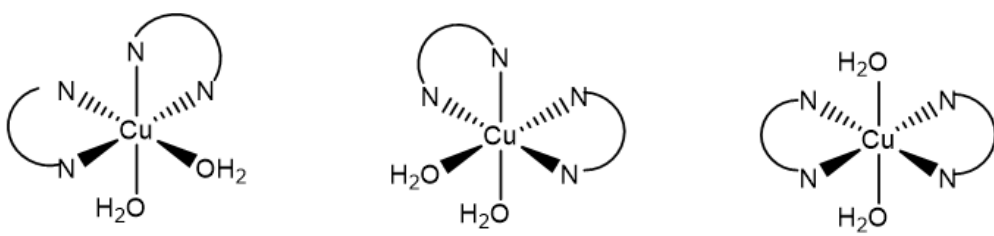
The annotations listed below were available to examiners marking this component in this series.

Annotations

Annotation	Meaning
	Correct point or mark awarded
	Incorrect point or mark not awarded
	Unclear
	Information missing or insufficient for credit
	Benefit of the doubt given
	Contradiction in response otherwise markworthy, mark not given
	Part of the correct answer has been seen. Full credit has not been awarded.
	Error carried forward applied
	Incorrect or insufficient point ignored while marking the rest of the response
	Benefit of the doubt not applied in this instance

Annotation	Meaning
RE	Rounding error
REP	Repetition
SEEN	Blank page or part of script seen
SF	Error in number of significant figures
TE	Transcription error

Question	Answer	Marks									
1(a)(i)	M1: (thermal stability) increases (down the group <i>or</i> from Mg to Ba) OR carbonates become more stable (down the group <i>or</i> from Mg to Ba) M2: ion / cation size / radius increases OR charge density on ion / cation decreases M3: less polarised $\text{CO}_3^{(2)-}$ / anion / carbonate <u>ion</u> OR weakens the C-O / C=O / covalent bonds less OR weakens bonds in the anion less OR lattice enthalpy of MO falls (less exothermic) faster than MCO_3 (due to differing sizes of anions)	3									
1(a)(ii)	$\text{CuCO}_3 \longrightarrow \text{CuO} + \text{CO}_2$	1									
1(b)(i)	[Ar] <table><tr><td>1↓</td><td>1↓</td><td>1↓</td><td>1↓</td><td>1</td><td></td><td></td><td></td><td></td></tr></table>	1↓	1↓	1↓	1↓	1					1
1↓	1↓	1↓	1↓	1							
1(b)(ii)		1									
1(b)(iii)	the (3)d and (4)s are close / similar in energy	1									
1(c)(i)	M1: species with / has two lone pairs of electrons M2: that form dative (covalent) / co-ordinate bond(s) to a (central) transition-element / metal AND atom / ion OR donates electron pair(s) to a (central) transition-element / metal AND atom / ion	2									

Question	Answer	Marks				
1(c)(ii)	 cis-1 cis-2 trans each correct structure = [1]	3				
1(c)(iii)	optical AND geometrical / cis-trans	1				
1(c)(iv)	identification of either cis isomer AND dipoles / polar bonds / partial charges do not cancel OR identification of either cis isomer AND it is asymmetric and has polar bonds	1				
1(d)(i)	<table><tr><td>complex ion in Na₃Al F₆</td><td>Al F₆³⁻</td></tr><tr><td>ligand in Na₃Al F₆</td><td>F⁻</td></tr></table>	complex ion in Na ₃ Al F ₆	Al F ₆ ³⁻	ligand in Na ₃ Al F ₆	F ⁻	1
complex ion in Na ₃ Al F ₆	Al F ₆ ³⁻					
ligand in Na ₃ Al F ₆	F ⁻					
1(d)(ii)	M1: charge passed = 1.5 × 30 × 60 OR 2700 (C) M2: n(e⁻) = 2700 / 96500 OR 2.80 × 10⁻² (mol) M3: n(Al) = 2.80 × 10⁻² ÷ 3 OR 9.33 × 10⁻³ (mol) M4: mass (Al) = (27 × 9.33 × 10⁻³) = 0.25 (g)	4				

Question	Answer	Marks
2(a)(i)	white / steamy fumes	1
2(a)(ii)	number of (possible) arrangements of particles AND energy in a system	1
2(a)(iii)	due to the large number increase in gas molecules formed (in the reaction) OR there are much more gaseous molecules in the products OR six gas moles are produced	1
2(a)(iv)	M1: $\Delta H_r = (-859) + (2 \times -297) + (4 \times -92) - (-1460) - (2 \times -246)$ OR $+131 \text{ kJ mol}^{-1}$ M2: use of $\Delta G = \Delta H - T\Delta S$ AND use of 298 (or 273+25) for T M3: $\Delta G = 131 - (298 \times 0.768) = -97.9 \text{ (kJ mol}^{-1}\text{)}$	3
2(b)(i)	$2\text{Ba}^{2+}(\text{aq}) + \text{Cr}_2\text{O}_7^{2-}(\text{aq}) + \text{H}_2\text{O}(\text{l}) \longrightarrow 2\text{BaCrO}_4(\text{s}) + 2\text{H}^+(\text{aq})$	1
2(b)(ii)	M1: d orbital(s) of two different energies / d-d splitting occurs OR d orbital(s) / d (sub)-shell splits / d-d gap OR (inferred from movement of an electron) from lower d to higher d orbital M2: electron(s) promoted / excited OR electron(s) moves to higher (d-) orbital OR electron(s) jumps up (to d- orbital) / jumps to higher (d-orbital) M3: wavelength / frequency / light / photon / $h\nu$ / hf absorbed OR radiation / energy from <u>visible</u> (region) absorbed AND colour (seen) is complementary OR wavelength / frequency / colour / light not absorbed is transmitted / reflected / seen	3
2(c)	M1: ΔH_{latt} and ΔH_{hyd} decrease / both become less exothermic / less negative M2: ΔH_{hyd} changes more / dominant factor / changes faster OR ΔH_{latt} changes less / becomes less exothermic by a smaller extent M3: ΔH_{sol} becomes less exothermic / less negative OR ΔH_{sol} becomes (more) endothermic / (more) positive	3

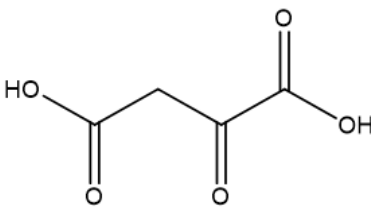
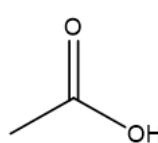
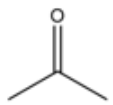
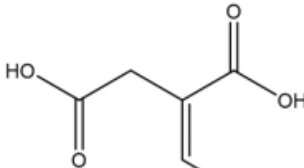
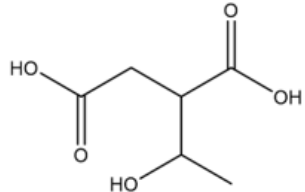
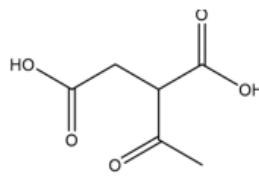
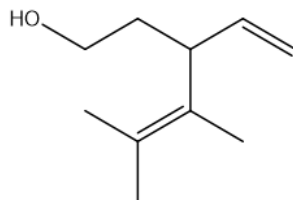
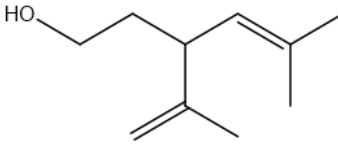
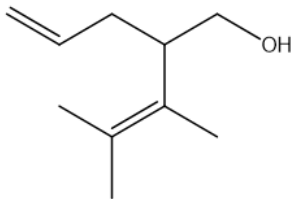
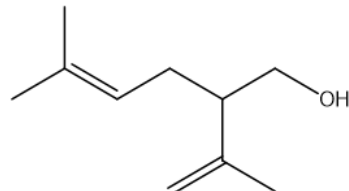
Question	Answer	Marks
3(a)(i)	M1: $K_{sp} = [\text{Ni}^{2+}][\text{IO}_3^-]^2$ M2: units: $\text{mol}^3 \text{dm}^{-9}$	2
3(a)(ii)	$[\text{Ni}^{2+}] = 2.3 \times 10^{-2}$; $[\text{IO}_3^-] = 4.6 \times 10^{-2}$ $K_{sp} = (2.3 \times 10^{-2}) \times (4.6 \times 10^{-2})^2 = 4.87 \times 10^{-5}$	1
3(b)(i)	$E^\circ_{\text{cell}} = 1.19 - (0.34) = (+)0.85 \text{ (V)}$ AND (positive electrode is) right hand side / nickel electrode / platinum / iodate	1
3(b)(ii)	(E_{cell} would be) less positive / more negative AND as $[\text{IO}_3^-]$ is less than 1.0 mol dm^{-3} / has a lower / smaller concentration	1
3(b)(iii)	M1: E_{cell} is less positive ticked M2: $[\text{IO}_3^-(\text{aq})]$ is lowered due to the common ion effect OR $\text{Ni}(\text{IO}_3)_2$ precipitating or shown by equation OR solubility of $\text{Ni}(\text{IO}_3)_2$ decreasing	2
3(c)(i)	$[\text{H}^+] = 10^{-\text{pH}} = 10^{-0.47} = 0.34 / 0.339 \text{ (mol dm}^{-3}\text{) min 2sf}$	1
3(c)(ii)	$[\text{HIO}_3]_{\text{eqm}} = 1.0 - 0.34 = 0.66 \text{ (mol dm}^{-3}\text{) min 2sf}$ AND $[\text{H}^+]_{\text{eqm}} = [\text{IO}_3^-]_{\text{eqm}} = 0.34 \text{ (mol dm}^{-3}\text{) min 2sf}$	1
3(c)(iii)	$K_a = (0.34)^2 / 0.66 = 0.17 - 0.18 \text{ (mol dm}^{-3}\text{) min 2sf}$	1
3(d)(i)	M1: opposes / resists change in pH / controls pH / pH kept within a small range M2: when small amount of acid / H^+ or alkali / base / OH^- is added	2
3(d)(ii)	M1: $k = \{2.1 \times 10^{-2}\} / \{0.5 \times (1 \times 10^{-3})^2 \times 0.01^2\} = 4.20 \times 10^8 \text{ min 2sf}$ M2: $\text{mol}^{-4} \text{dm}^{12} \text{s}^{-1}$	2
3(d)(iii)	rate will be $(0.03 / 0.01)^2 = 9$ times as fast, so rate = $9 \times 2.1 \times 10^{-2} = 0.19 \text{ (0.189) (mol dm}^{-3} \text{s}^{-1}\text{) min 2sf}$	1

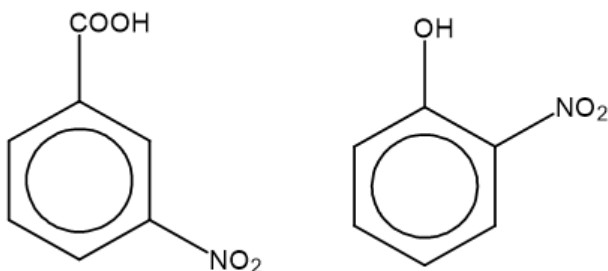
Question	Answer			Marks
4(a)	polymer	type of polymerisation	structure of monomer	4
	X	condensation	H ₂ (CH ₂) ₅ CO ₂ H OR NH ₂ (CH ₂) ₅ COC <i>l</i>	
	Y	addition	CH ₂ =CHCO ₂ CH ₃	
	Z	condensation	CH ₃ CH(OH)CO ₂ H OR CH ₃ CH(OH)COC <i>l</i>	
	Type of polymerisation (all three) = [1] Structure of monomers = [1] for each structure			
4(b)	pH (where) the species is a zwitterion (is the dominant form) OR pH (where) the species is (electrically) neutral OR pH (where) the species has a (net overall) charge of zero			1
4(c)(i)	</			

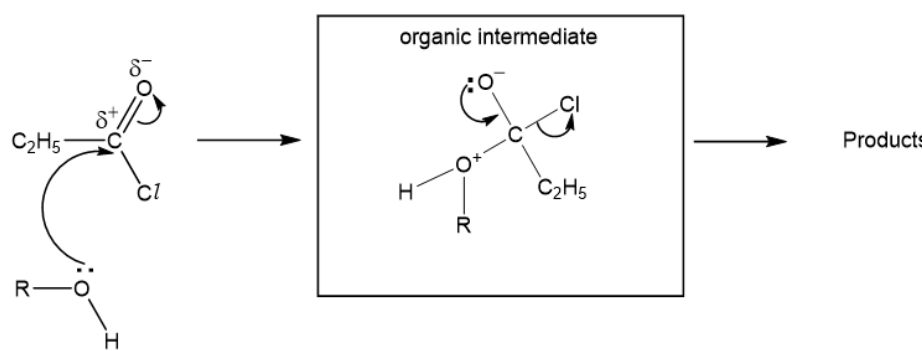
Question	Answer	Marks
5(a)(i)		1
5(a)(ii)	the OH / NH groups / O atoms AND can hydrogen bond with water	1
5(a)(iii)	$C_{33}H_{35}N_2O_5F$	1
5(a)(iv)		1
5(a)(v)	different / better biological activity	1
5(b)(i)	16 / sixteen	1
5(b)(ii)	M1: 4 / four M2: proton exchange between NH / OH with D	2

Question	Answer	Marks
5(c)	groups reacting • carboxylic acid • amide groups forming • (primary) alcohol • amine Any two [1] all four [2]	2

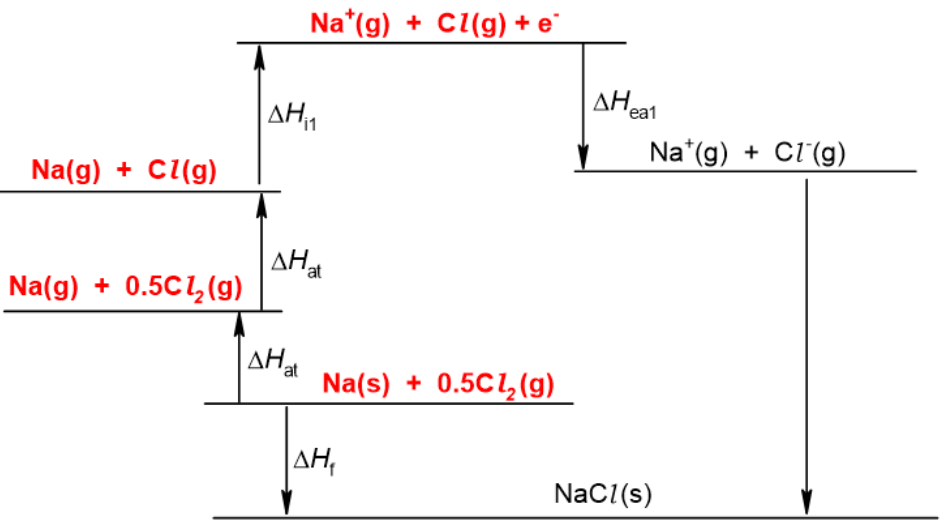
Question	Answer	Marks																		
6(a)	<table><tr><th colspan="6">test</th></tr><tr><td>sodium metal</td><td>Na₂CO₃(aq)</td><td>2,4-DNPH</td><td>I₂(aq) + OH⁻(aq)</td><td>warm with Fehling's solution</td><td>Br₂(aq)</td></tr><tr><td>alcohol / carboxylic acid</td><td>carboxylic acid</td><td>carbonyl / ketone / aldehyde</td><td>(methyl) ketone / (methyl) alcohol</td><td>aldehyde</td><td>alkene</td></tr></table> Any two [1] any three [2] any five [3] all six [4]	test						sodium metal	Na ₂ CO ₃ (aq)	2,4-DNPH	I ₂ (aq) + OH ⁻ (aq)	warm with Fehling's solution	Br ₂ (aq)	alcohol / carboxylic acid	carboxylic acid	carbonyl / ketone / aldehyde	(methyl) ketone / (methyl) alcohol	aldehyde	alkene	4
test																				
sodium metal	Na ₂ CO ₃ (aq)	2,4-DNPH	I ₂ (aq) + OH ⁻ (aq)	warm with Fehling's solution	Br ₂ (aq)															
alcohol / carboxylic acid	carboxylic acid	carbonyl / ketone / aldehyde	(methyl) ketone / (methyl) alcohol	aldehyde	alkene															
6(b)(i)	(primary) alcohol	1																		
6(b)(ii)	M1: A → B oxidation/dehydrogenation M2: C → E reduction/(nucleophilic) addition M3: E → F elimination/dehydration	3																		

Question	Answer	Marks
6(b)(iii)	G  H  each correct structure [1]	2
6(b)(iv)	D  F  E  C  A any one of the following     each correct structure [1]	5

Question	Answer	Marks
7(a)	M1: benzoic acid > phenol > ethanol M2 / M3: any two [1] any three [2] - correct link of acidity once AND weakens O—H / OH bond / hydroxyl bond / -O-H OR H⁺ more easily lost / (carboxylate) anion stabilised u / c (benzoic acid) due to negative inductive effect / electron withdrawing effect AND of C=O / COOH / carboxyl (phenol) as lone pair / p-orbital (electrons) on oxygen (on phenol) / AND overlap / delocalised into the ring / π-system (ethanol) alkyl / ethyl / R group AND is electron donating / positive inductive effect	3
7(b)	 Each correct structure = [1]	2
7(c)(i)	ethyl propanoate	1

Question	Answer	Marks
7(c)(ii)	 M1 / M2: any two [1] all four [2] • lone pair on O • correct arrow from (lone pair) O to C (of C=O) • dipole on C=O • correct arrow on C=O M3: correct intermediate M4: arrow from lone pair or charge on O⁻ to C-O bond AND arrow from C-Cl to Cl	4
7(c)(iii)	(nucleophilic) addition – elimination	1

Question	Answer	Marks
8(a)(i)	M1: energy change / released when 1 mole of a (ionic) solid / lattice / crystal / compound is formed M2: from gas (phase) ions / gaseous ions (under standard conditions)	2
8(a)(ii)	(enthalpy change when) 1 mole of a substance / solid / solute / molecule AND dissolves in water (to give a solution of infinite dilution)	1
8(b)	$\Delta H_{\text{sol}}(\text{NaCl}) = \Delta H_{\text{hyd}}(\text{Na}^+) + \Delta H_{\text{hyd}}(\text{Cl}^-) - \Delta H_{\text{latt}}(\text{NaCl})$	1

Question	Answer	Marks
8(c)	 M1 / M2: any two [1] all four [2] M3: all state symbols for the formula are present and correct	3
8(d)	M1: (more exothermic because) Cl^- OR NO_3^- because its (ionic) radius / size is smaller M2: (more exothermic because) (ion-dipole) attraction / bond between it and water is stronger OR M1: NO_3^- because it has lone pairs on O / more lone pairs M2: which can form hydrogen bonds with water	2