



Please write clearly in block capitals.

Centre number

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Candidate number

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Surname

StemUp Model Solution

Forename(s)

Candidate signature

I declare this is my own work.

A-level CHEMISTRY

Paper 3

Friday 20 June 2025

Morning

Time allowed: 2 hours

Materials

For this paper you must have:

- the Periodic Table/Data Booklet, provided as an insert (enclosed)
- a ruler with millimetre measurements
- a scientific calculator, which you are expected to use where appropriate.

Instructions

- Use black ink or black ball-point pen.
- Fill in the boxes at the top of this page.
- Answer **all** questions.
- You must answer the questions in the spaces provided. Do **not** write outside the box around each page or on blank pages.
- If you need extra space for your answer(s), use the lined pages at the end of this book. Write the question number against your answer(s).
- All working must be shown.
- Do **all** rough work in this book. Cross through any work you do not want to be marked.

Information

- The marks for questions are shown in brackets.
- The maximum mark for this paper is 90.

Advice

- You are advised to spend 70 minutes on **Section A** and 50 minutes on **Section B**.

For Examiner's Use	
Question	Mark
1	
2	
3	
4	
5	
6	
Section B	
TOTAL	



J U N 2 5 7 4 0 5 3 0 1

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7405/3

Section AAnswer **all** questions in this section.

0	1
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A value for enthalpy of solution can be determined from the results of a calorimetry experiment or by the application of Hess's law.

0	1	.	1
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Give the equation, including state symbols, for the process that occurs when the enthalpy of solution of potassium hydroxide is determined.

[1 mark]

0 1 2

A student does an experiment to determine a value for the enthalpy of solution of potassium hydroxide.

The student uses this method.

- Add some distilled water to a measuring cylinder and record the volume.
- Pour the water into a polystyrene cup.
- Record the temperature of the water in the cup.
- Add 3.00 g of solid KOH to the water in the cup.
- Stir the solution and record the maximum temperature reached.

Table 1 shows the student's results.

Table 1

Initial temperature / °C	18.1
Final temperature / °C	30.4

The student calculated the enthalpy of solution of potassium hydroxide to be $-51.3 \text{ kJ mol}^{-1}$

The specific heat capacity of the solution, $c = 4.18 \text{ J K}^{-1} \text{ g}^{-1}$

Assume the density of the solution = 1.00 g cm^{-3}

Calculate the volume of water that the student added to the cup.

[5 marks]

$$\text{Amount of KOH} = 3.00 \div 56.1 = 0.05348 \text{ mol}$$

$$q = \Delta H \times n = 51.3 \times 0.05348 = 2.743 \text{ kJ} = 2743 \text{ J}$$

$$\Delta T = 30.4 - 18.1 = 12.3 \text{ }^{\circ}\text{C} (= 12.3 \text{ K})$$

$$q = mc\Delta T \text{ so } m = q \div (c\Delta T) = 2743 \div (4.18 \times 12.3)$$

$$m = 53.4 \text{ g of solution}$$

$$\text{density} = 1.00 \text{ g cm}^{-3}, \text{ so volume of water} = 53.4 \text{ cm}^3$$

Volume of water 53.4 Units cm³

Turn over ►



0 1 . 3 Table 2 shows some enthalpy change data.

Table 2

	Enthalpy change / kJ mol ⁻¹
Enthalpy of lattice formation of potassium hydroxide	-789
Enthalpy of hydration of potassium ions	-322
Enthalpy of hydration of hydroxide ions	-528

Use the data in **Table 2** to calculate the enthalpy of solution for potassium hydroxide.
[2 marks]

$$\begin{aligned}
 \Delta H_{\text{soln}} &= \Delta H_{\text{hyd}}(\text{K}^+) + \Delta H_{\text{hyd}}(\text{OH}^-) - \Delta H_{\text{lattice formation}} \\
 &= (-322) + (-528) - (-789) \\
 &= -61 \text{ kJ mol}^{-1}
 \end{aligned}$$

Enthalpy of solution -61 kJ mol⁻¹

0 1 . 4 Suggest why the value calculated in Question **01.3** is different from the experimental value given in Question **01.2**.

[1 mark]

Heat is lost to the surroundings (through the cup and from the surface), or not all of the KOH dissolved, so the temperature rise measured is too small.



Turn over for the next question

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ANSWER IN THE SPACES PROVIDED**

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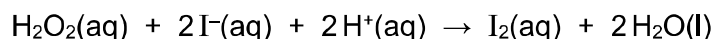


0 5

0 2

The effect of reactant concentrations on the rate of a reaction is investigated.

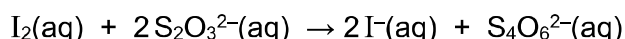
The equation for the reaction is shown.



The experiment is done five times in the presence of thiosulfate ions and starch.

The reactant concentrations are varied by changing the volume of each reactant in the mixture.

The thiosulfate ions react with the iodine as it is produced, until all the thiosulfate has reacted.



After all the thiosulfate has reacted, the iodine produced in the mixture then reacts with the starch and the solution turns blue-black.

The time taken for the solution to turn blue-black is recorded.

If the amount of thiosulfate is the same in each experiment, the time recorded represents the time to produce the same amount of iodine in each experiment.

Table 3 shows the results of the investigation.

Table 3

Experiment	$[\text{H}_2\text{O}_2]$ / mol dm^{-3}	$[\text{I}^-]$ / mol dm^{-3}	$[\text{H}^+]$ / mol dm^{-3}	Time for reaction mixture to turn blue-black / s
1	0.010	0.010	0.10	60
2	0.020	0.010	0.10	30
3	0.030	0.010	0.10	20
4	0.030	0.020	0.10	10
5	0.030	0.020	0.20	10



0 2 . 1

The amount of thiosulfate is the same in each experiment. To allow the order of reaction with respect to each reactant to be calculated, several other variables need to be controlled.

Identify **three** other control variables.

[3 marks]

Control variable 1 Temperature (of the reaction mixture)

Control variable 2 Volume / amount of starch solution added

Control variable 3 Total volume of the reaction mixture

0 2 . 2

Deduce the order of reaction with respect to each reactant.

Explain how you deduced the order of reaction with respect to iodide ions.

[3 marks]

Order with respect to H_2O_2 1 (first order)

Order with respect to I^- 1 (first order)

Order with respect to H^+ 0 (zero order)

Explanation Use experiments 3 and 4; only $[\text{I}^-]$ changes, $[\text{H}_2\text{O}_2]$ and $[\text{H}^+]$ stay the same. Doubling $[\text{I}^-]$ halves the time (20 s to 10 s), so the rate doubles, so it is first order with respect to I^- .

0 2 . 3

Give the rate equation for the reaction.

Determine the units of the rate constant.

[2 marks]

Rate equation

$\text{rate} = k[\text{H}_2\text{O}_2][\text{I}^-]$

Units of rate constant $\text{mol}^{-1} \text{dm}^3 \text{s}^{-1}$

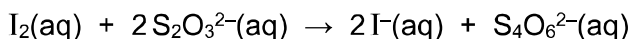
Question 2 continues on the next page

Turn over ►



0 2 . 4

In each experiment, 1.00 cm^3 of $1.00 \times 10^{-3} \text{ mol dm}^{-3}$ sodium thiosulfate solution is used to react with the iodine initially produced.



The total volume of the reaction mixture in experiment 3 is 40.0 cm^3

Calculate the initial rate of production of iodine, in $\text{mol dm}^{-3} \text{ s}^{-1}$, for experiment 3.

[3 marks]

$$n(\text{S}_2\text{O}_3^{2-}) = 1.00/1000 \times 1.00 \times 10^{-3} = 1.00 \times 10^{-6} \text{ mol}$$

$$n(\text{I}_2) = \frac{1}{2} \times 1.00 \times 10^{-6} = 5.00 \times 10^{-7} \text{ mol}$$

$$[\text{I}_2] = 5.00 \times 10^{-7} \div (40.0/1000) = 1.25 \times 10^{-5} \text{ mol dm}^{-3}$$

$$\text{time for experiment 3} = 20 \text{ s}$$

$$\text{rate} = 1.25 \times 10^{-5} \div 20 = 6.25 \times 10^{-7} \text{ mol dm}^{-3} \text{ s}^{-1}$$

Initial rate 6.25×10^{-7} $\text{mol dm}^{-3} \text{ s}^{-1}$

0 2 . 5

The total uncertainty in the times recorded for each experiment is $\pm 0.5 \text{ s}$ but the stopwatch used has a measurement uncertainty of $\pm 0.01 \text{ s}$

Suggest why, in each experiment, the uncertainty in the times recorded is greater than the measurement uncertainty of the stopwatch.

[1 mark]

The observer's reaction time in stopping the timer, as it is hard to judge exactly when the colour change (end point) occurs.

0 2 . 6

Calculate the percentage uncertainty in the time recorded in experiment 3.

[1 mark]

$$\% \text{ uncertainty} = 0.5 \div 20 \times 100$$

Percentage uncertainty 2.5%



0 2 . 7

The five experiments are repeated with the same stopwatch and the same concentrations and volumes of H_2O_2 , I^- and H^+

Suggest and explain how the method can be changed to decrease the percentage uncertainty in the times recorded.

[2 marks]

Use a greater concentration (or amount) of thiosulfate, so more iodine must form before the blue-black colour appears.

This increases the times, so the $\pm 0.5 \text{ s}$ is a smaller percentage of each time.

15

Turn over for the next question

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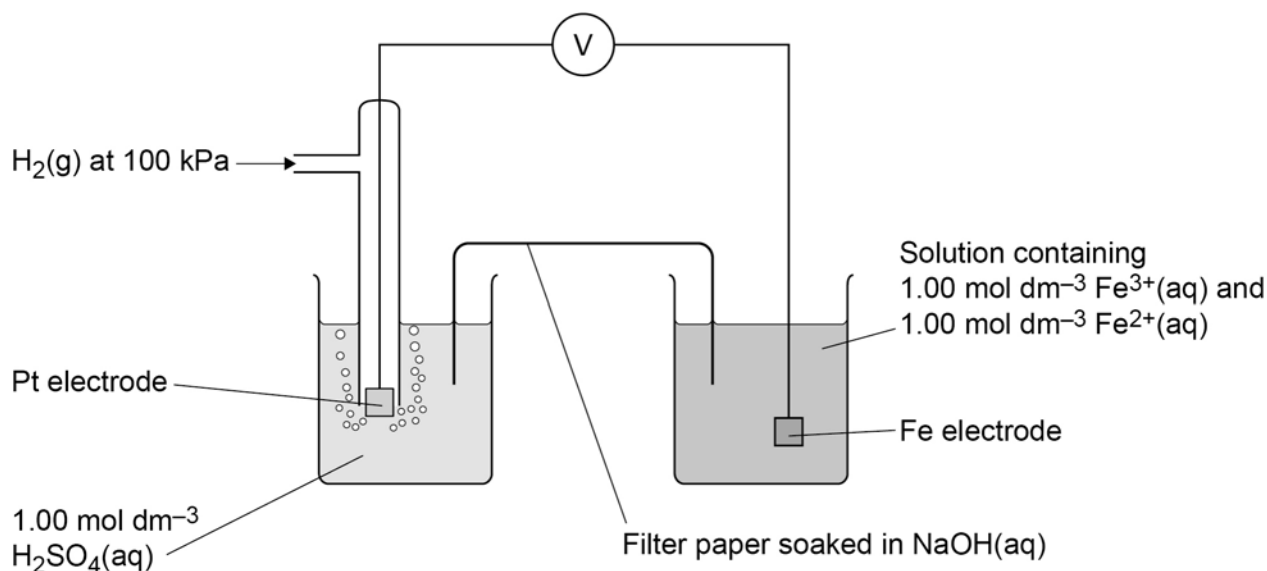


0 3

A student does an experiment to measure the standard electrode potential of an $\text{Fe}^{3+}/\text{Fe}^{2+}$ electrode at 298 K

Figure 1 shows the apparatus the student uses.

Figure 1



0 3 . 1

Figure 1 shows **three** mistakes made by the student.

Identify each mistake and the change needed to correct it.

[3 marks]

Mistake 1 $[\text{H}^+]$ is 2.00 mol dm^{-3} , not 1.00 , as H_2SO_4 is diprotic

Change 1 Use $0.500 \text{ mol dm}^{-3} \text{H}_2\text{SO}_4$ (or $1.00 \text{ mol dm}^{-3} \text{HCl}$)

Mistake 2 Salt bridge soaked in NaOH , which reacts with the acid

Change 2 Soak the filter paper in (saturated) $\text{KNO}_3(\text{aq})$ instead

Mistake 3 An Fe electrode is used in the $\text{Fe}^{3+}/\text{Fe}^{2+}$ half-cell

Change 3 Use a platinum (Pt) electrode instead of Fe



0 3 . 2 What feature of the voltmeter ensures that the voltmeter measures the cell EMF?

[1 mark]

Tick (✓) **one** box.

Low resistance

☐

Variable resistance

☐

High resistance

☒

0 3 . 3 The $\text{Fe}^{3+}/\text{Fe}^{2+}$ electrode is the positive electrode when the cell is set up correctly at 298 K

Give the redox equation for the overall reaction that occurs in the cell when the voltmeter is replaced with a lamp.

[1 mark]



Question 3 continues on the next page

Turn over ►



0 3 . 4 Table 4 shows some standard electrode potentials.

Table 4

Half-equation	$E^\ominus / \text{V}$
$\text{Fe}^{2+}(\text{aq}) + 2\text{e}^- \rightarrow \text{Fe}(\text{s})$	-0.44
$\text{Cu}^{2+}(\text{aq}) + \text{e}^- \rightarrow \text{Cu}^+(\text{aq})$	+0.15
$\text{Cu}^{2+}(\text{aq}) + 2\text{e}^- \rightarrow \text{Cu}(\text{s})$	+0.34
$\text{Cu}^+(\text{aq}) + \text{e}^- \rightarrow \text{Cu}(\text{s})$	+0.52
$\text{Fe}^{3+}(\text{aq}) + \text{e}^- \rightarrow \text{Fe}^{2+}(\text{aq})$	+0.77

Finely powdered iron filings are added to a solution of copper(II) sulfate and the mixture is stirred.

The solution gradually changes from blue to pale green and a brown solid forms.

Use the data in **Table 4** to explain these observations.

Give an ionic equation for the reaction that occurs.

Give the conventional representation of the cell that has the same overall reaction as in your equation.

Calculate the EMF, in volts, of this cell.

[6 marks]

Explanation *$E^\ominus \text{Fe}^{2+}/\text{Fe} (-0.44 \text{ V})$ is more negative than $E^\ominus \text{Cu}^{2+}/\text{Cu} (+0.34 \text{ V})$, so Fe is oxidised to Fe^{2+} and Cu^{2+} is reduced to Cu (EMF is positive).*

The brown solid formed is copper, Cu.

The solution changes from blue (Cu^{2+}) to pale green because $\text{Fe}^{2+}(\text{aq})$ ions form.

Equation



Conventional representation $Fe(s) | Fe^{2+}(aq) || Cu^{2+}(aq) | Cu(s)$

$$EMF = 0.34 - (-0.44) = +0.78 \text{ V}$$

EMF +0.78 V

11

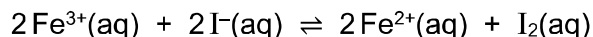
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Turn over ►



0 4

50.0 cm³ of 0.015 mol dm⁻³ Fe₂(SO₄)₃(aq) are mixed with 50.0 cm³ of 0.015 mol dm⁻³ KI(aq) at a temperature of 40 °C
The mixture is left for several hours until equilibrium is reached.



A sample of the equilibrium mixture is titrated to determine the amount of I⁻(aq) in the equilibrium mixture.

0 4 . 1

Give the expression for the equilibrium constant, K_c , for the equilibrium.

[1 mark]

$$K_c = \frac{[\text{Fe}^{2+}]^2 [\text{I}_2]}{[\text{Fe}^{3+}]^2 [\text{I}^{-}]^2}$$

0 4 . 2

Before the titration, the sample is 'quenched' by adding 50 cm³ of ice-cold water.

Suggest why the sample is 'quenched' in this way before the titration.
Explain your answer.

[2 marks]

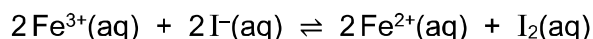
*Dilution and cooling make the reaction so slow that it effectively stops,
so the amounts in the sample stay as they were at equilibrium.
So during the titration the equilibrium does not shift and no more I⁻ is
formed, so the titre measures the true equilibrium amount of I⁻.*



0 4 . 3

The amount of $\text{I}^{-}(\text{aq})$ in the equilibrium mixture is found to be $3.00 \times 10^{-4} \text{ mol}$

Calculate the value of K_c at 40°C for the equilibrium.



Give the units of K_c

Assume that the total volume of the mixture remains constant at 100 cm^3

[6 marks]

$$\text{Initial } n(\text{Fe}^{3+}) = 2 \times 0.015 \times 50.0/1000 = 1.50 \times 10^{-3} \text{ mol}$$

$$\text{Initial } n(\text{I}^{-}) = 0.015 \times 50.0/1000 = 7.50 \times 10^{-4} \text{ mol}$$

$$n(\text{I}^{-}) \text{ reacted} = 7.50 \times 10^{-4} - 3.00 \times 10^{-4} = 4.50 \times 10^{-4} \text{ mol}$$

$$\text{At equilibrium: } n(\text{Fe}^{3+}) = 1.50 \times 10^{-3} - 4.50 \times 10^{-4} = 1.05 \times 10^{-3} \text{ mol}$$

$$n(\text{Fe}^{2+}) = 4.50 \times 10^{-4} \text{ mol}, \quad n(\text{I}_2) = 4.50 \times 10^{-4} \div 2 = 2.25 \times 10^{-4} \text{ mol}$$

$$\text{Divide by } 0.100 \text{ dm}^3: [\text{Fe}^{3+}] = 1.05 \times 10^{-2}, \quad [\text{I}^{-}] = 3.00 \times 10^{-3},$$

$$[\text{Fe}^{2+}] = 4.50 \times 10^{-3}, \quad [\text{I}_2] = 2.25 \times 10^{-3} \quad (\text{all in mol dm}^{-3})$$

$$K_c = \frac{(4.50 \times 10^{-3})^2 \times (2.25 \times 10^{-3})}{(1.05 \times 10^{-2})^2 \times (3.00 \times 10^{-3})^2} = 45.9$$

$$\text{Units} = (\text{mol dm}^{-3})^3 \div (\text{mol dm}^{-3})^4 = \text{mol}^{-1} \text{ dm}^3$$

$$K_c \quad 45.9$$

$$\text{Units of } K_c \quad \text{mol}^{-1} \text{ dm}^3$$

Turn over ►



0 5

This question is about oxidation.

0 5 . 1

A student has separate unlabelled samples of propan-1-ol and propan-2-ol but does not know which sample is which alcohol.

The student plans a series of steps to identify the alcohols.

Step 1 Oxidise each sample, distilling off the organic product as it is formed.

Step 2 Identify the organic products using a simple chemical test.

Step 3 Confirm the identities using the infrared spectrum of each original sample.

Give further details of each step in the plan. In your answer include

- an appropriate oxidising agent used in Step 1
- equations, using [O] to represent the oxidising agent, for each reaction in Step 1
- a suitable test and expected observations in Step 2
- how the infrared spectra can be used to confirm the identities.

[6 marks]

Step 1: Add each alcohol separately to acidified potassium dichromate(VI) ($K_2Cr_2O_7$ with dilute sulfuric acid), warm gently and distil off the organic product as it forms, so that the aldehyde is not oxidised further to a carboxylic acid. In both cases the dichromate turns from orange to green.

Propan-1-ol is a primary alcohol, so it is oxidised to the aldehyde propanal:



Propan-2-ol is a secondary alcohol, so it is oxidised to the ketone propanone:



Step 2: Warm a sample of each distillate with Tollens' reagent (or with Fehling's solution).

The product from propan-1-ol (propanal, an aldehyde) is oxidised and gives a silver mirror (with Fehling's solution, a brick-red precipitate).

The product from propan-2-ol (propanone, a ketone) gives no visible



change, as ketones are not oxidised.

So the sample whose product gives the silver mirror is propan-1-ol, and the sample whose product shows no change is propan-2-ol.

Step 3: Record the infrared spectrum of each original sample. Both spectra show a broad absorption at 3230 to 3550 cm^{-1} for the O-H bond in an alcohol (and C-O at 1000 to 1300 cm^{-1}), so this confirms both are alcohols but does not tell them apart.

Then compare the fingerprint region (below 1500 cm^{-1}) of each spectrum with the spectra of propan-1-ol and propan-2-ol in a database. The fingerprint region is unique to each compound, so an exact match confirms the identity of each alcohol.

Question 5 continues on the next page

Turn over ►



0 5 . 2

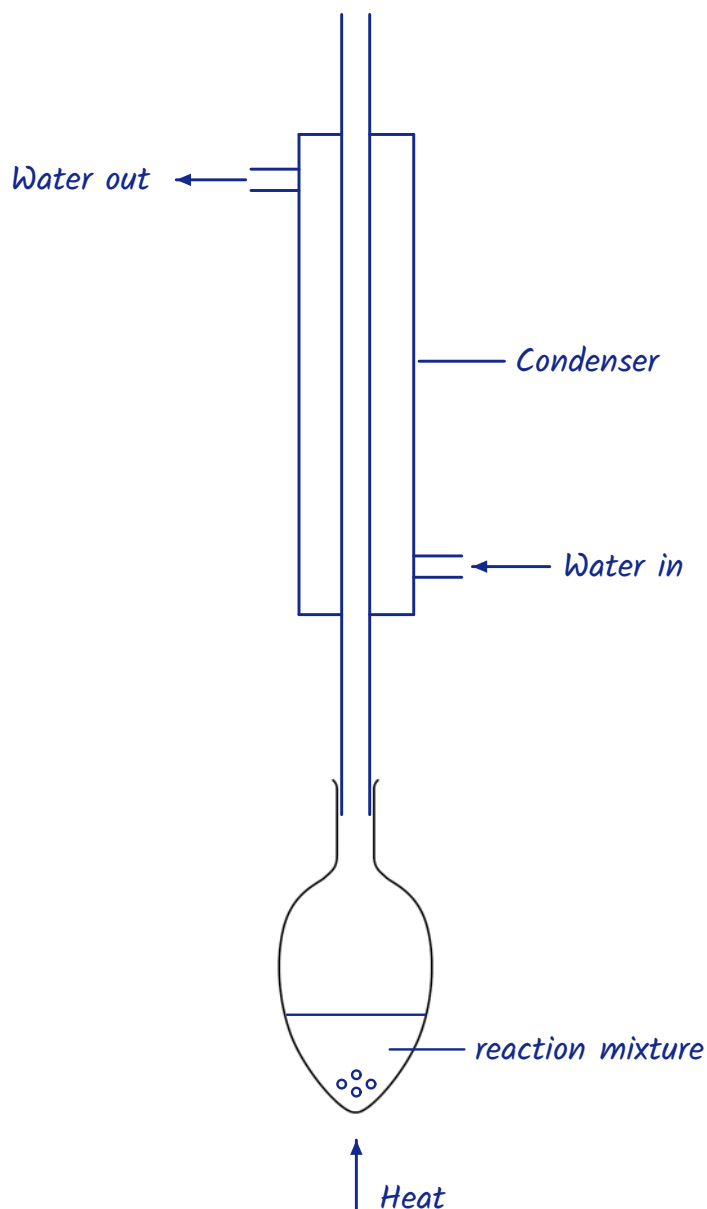
A primary alcohol can be oxidised to form a carboxylic acid using an excess of the oxidising agent in Question 05.1.

Give the condition needed for this reaction.

Complete the diagram, adding labels, to show how the apparatus would be set up.

[3 marks]

Condition Heat under reflux



0 5 . 3

During oxidation reactions, anti-bumping granules are often added to the pear-shaped flask containing the reactants.

State why anti-bumping granules are used.

[1 mark]

To produce small bubbles so the mixture boils smoothly, preventing large bubbles forming.

10

Turn over for the next question

Turn over ►

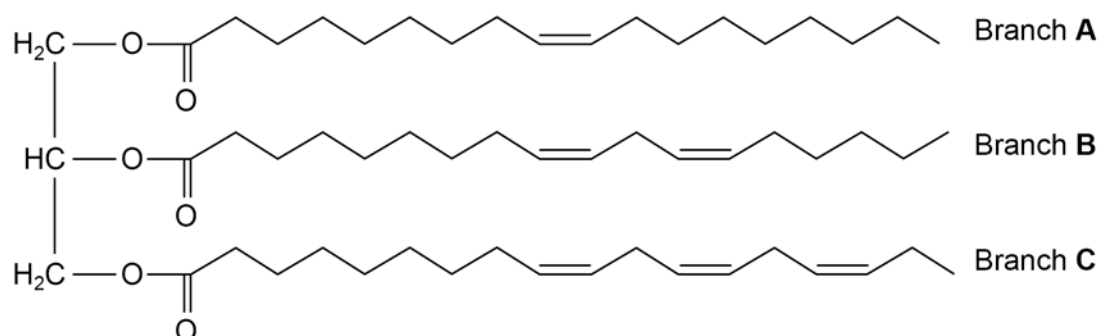


0 6

This question is about rapeseed oil.

Figure 2 shows the structure of compound **Q** found in rapeseed oil. Each branch in the molecule is formed from a different long-chain carboxylic acid.

Figure 2



0 6 . 1

State the type of stereoisomerism shown in Branch **A**.

[1 mark]

E-Z (cis-trans) stereoisomerism

0 6 . 2

Q can be reacted with methanol to form three different esters.

State the number used to identify the position of the C=C bond in the name of the ester formed from Branch **A**.

State the number of different stereoisomers of the ester formed from Branch **B**.

Give a use for these esters.

[3 marks]

Number used in the name of the ester 9

Number of stereoisomers 4 (four)

Use for esters Biodiesel (a biofuel)



0 6 . 3

Q also reacts with NaOHState the formula of the product, formed by Branch **B** in **Figure 2**, in this reaction.

State a use for this product.

[2 marks]Formula $C_{17}H_{31}COO^-Na^+$ Use for product As a soap (an anionic surfactant)

6

Turn over for Section B**Turn over ►**

Section B

Answer **all** questions in this section.Only **one** answer per question is **allowed**.For each question **completely fill** in the circle alongside the appropriate answer.

CORRECT METHOD



WRONG METHODS



If you want to change your answer you must cross out your original answer as shown.

If you wish to return to an answer previously crossed out, ring the answer you now wish to select as shown.

You may do your working in the **blank** space around each question but this **will** not be marked.
Do **not** use additional sheets for this working.

0 7

The reaction between methylamine (CH_3NH_2) and excess bromoethane ($\text{C}_2\text{H}_5\text{Br}$) produces a quaternary ammonium salt.

What is the formula for the cation in this salt?

[1 mark]

A $(\text{C}_2\text{H}_5)_4\text{N}^+$ ☐B $(\text{C}_2\text{H}_5)_3\text{N}^+\text{CH}_3$ ☒C $(\text{C}_2\text{H}_5)_2\text{N}^+(\text{CH}_3)_2$ ☐D $\text{C}_2\text{H}_5\text{N}^+(\text{CH}_3)_3$ ☐

0 8

Each of the conversions shown uses KOH as a reactant.

In which conversion does KOH **not** act as a base?

[1 mark]

A $(\text{CH}_3)_2\text{NH}_2^+\text{Br}^-$ to $(\text{CH}_3)_2\text{NH}$ ☐B CH_3COOH to $\text{CH}_3\text{COO}^-\text{K}^+$ ☐C $\text{CH}_3\text{CH}_2\text{Br}$ to $\text{H}_2\text{C}=\text{CH}_2$ ☐D CH_3Br to CH_3OH ☒

0 9

Which statement about the use of a catalyst in a reaction is correct?

[1 mark]

A The activation energy of the reverse reaction increases.

☐

B The equilibrium constant increases.

☐

C The rate of the reverse reaction increases.

☒

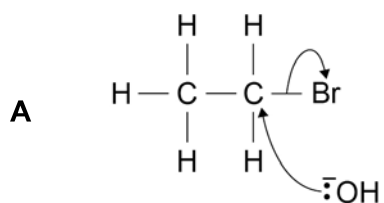
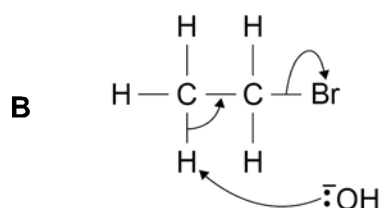
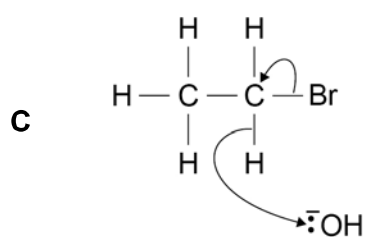
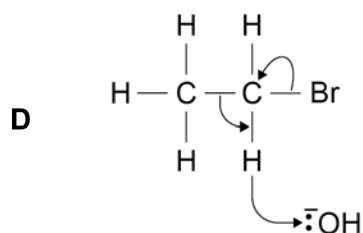
D The enthalpy change for the reaction decreases.

☐

1 0

Which diagram shows a correct step in the mechanism for an elimination reaction of bromoethane?

[1 mark]

☐☒☐☐

Turn over ►



1 1

The melting points of some Period 3 oxides are shown.

Oxide	Na ₂ O	MgO	Al ₂ O ₃	SiO ₂	P ₄ O ₁₀	SO ₃
Melting point / °C	920	2900	2040	1610	580	17

Which row is correct?

[1 mark]

	Giant ionic	Giant covalent	Simple molecular	
A	Na ₂ O MgO Al ₂ O ₃	SiO ₂	P ₄ O ₁₀ SO ₃	☒
B	Na ₂ O MgO Al ₂ O ₃	SiO ₂ P ₄ O ₁₀	SO ₃	☐
C	Na ₂ O MgO	Al ₂ O ₃ SiO ₂	P ₄ O ₁₀ SO ₃	☐
D	Na ₂ O	MgO Al ₂ O ₃ SiO ₃	P ₄ O ₁₀ SO ₃	☐

1 2

A sample of UF₆ consists of molecules containing ¹⁹F and either ²³⁵U or ²³⁸U atoms.

The relative molecular mass of the sample is 351.43

What is the percentage abundance of ²³⁵UF₆ in the sample?

[1 mark]

- A** 19 ☒
- B** 33 ☐
- C** 50 ☐
- D** 81 ☐



1 3

Compound **X** reacts with concentrated sulfuric acid to produce compound **Y**.

Compound **Y** reacts with an excess of water to produce an alcohol.

Which compound could be **X**?

[1 mark]

A Ethanoyl chloride

☐

B Methylbenzene

☐

C Propanone

☐

D Propene

☒**1 4**

A piece of strontium is added to an excess of water.

Which observation is correct and complete?

[1 mark]

A Colourless solution

☐

B Colourless solution and effervescence

☒

C Dense white precipitate

☐

D Dense white precipitate and effervescence

☐**1 5**

Which statement about trends from left to right across Period 3 is correct?

[1 mark]

A Atomic radius increases from Na to Si and then decreases

☐

B Atomic radius decreases

☒

C First ionisation energy decreases

☐

D Electronegativity decreases

☐**Turn over ►**

1 6

Which statement about ionisation energies in Period 3 is **not** correct?

[1 mark]

A The first ionisation energy of Mg is greater than that of Na

☐

B The first ionisation energy of Al is less than that of Mg

☐

C The first ionisation energy of P is less than that of S

☒

D Ar has the highest first ionisation energy in the period.

☐

1 7

Which type(s) of stereoisomerism is/are shown by the complex ion $[\text{Cr}(\text{H}_2\text{NCH}_2\text{CH}_2\text{NH}_2)_3]^{3+}$?

[1 mark]

A none

☐B *cis-trans* only☐

C optical only

☒D *cis-trans* and optical☐

1 8

What is the correct structural formula for 4-chloro-2-methylpent-2-enoic acid?

[1 mark]

A $\text{CH}_3\text{CCl}=\text{CHCH}(\text{CH}_3)\text{COOH}$ ☐B $(\text{CH}_3)_2\text{C}=\text{CHCHClCOOH}$ ☐C $\text{CH}_3\text{CHClCH}=\text{C}(\text{CH}_3)\text{COOH}$ ☒D $(\text{CH}_3)_2\text{CHCH}=\text{CClCOOH}$ ☐

1 9

0.100 mol of each of the oxides listed is added to separate 1.00 dm³ samples of water.

Which oxide produces the solution with the lowest pH?

[1 mark]**A** MgO☐**B** Al₂O₃☐**C** SO₂☐**D** SO₃☒**2 0**

Which property of the Group 2 elements increases with increasing atomic number?

[1 mark]**A** Atomic radius☒**B** Electronegativity☐**C** First ionisation energy☐**D** Melting point☐**2 1**

Which electron configuration represents an element that can have a maximum oxidation state of +5 in its compounds?

[1 mark]**A** 1s² 2s² 2p¹☐**B** 1s² 2s² 2p⁶ 3s² 3p⁵☐**C** 1s² 2s² 2p⁶ 3s² 3p⁶ 4s² 3d³☒**D** 1s² 2s² 2p⁶ 3s² 3p⁶ 4s² 3d⁵☐**Turn over ►**

2 2

Which statement is correct?

[1 mark]**A** Tollens' reagent contains linear $[\text{Ag}(\text{NH}_3)_2]^+$ ions.☒**B** Tollens' reagent contains octahedral $[\text{Ag}(\text{NH}_3)_6]^+$ ions.☐**C** Fehling's solution contains linear $[\text{Ag}(\text{NH}_3)_2]^+$ ions.☐**D** Fehling's solution contains octahedral $[\text{Ag}(\text{NH}_3)_6]^+$ ions.☐**2 3**

Which reaction produces an octahedral complex ion?

[1 mark]**A** The addition of excess aqueous ammonia to a solution containing $[\text{Co}(\text{H}_2\text{O})_6]^{2+}$ ions.☒**B** The addition of excess aqueous sodium carbonate to a solution containing $[\text{Fe}(\text{H}_2\text{O})_6]^{2+}$ ions.☐**C** The addition of excess concentrated hydrochloric acid to a solution containing $[\text{Cu}(\text{H}_2\text{O})_6]^{2+}$ ions.☐**D** The addition of excess aqueous sodium carbonate to a solution containing $[\text{Fe}(\text{H}_2\text{O})_6]^{3+}$ ions.☐**2 4**

Ethanol biofuel can be produced by fermentation.

Which statement about this process is correct?

[1 mark]**A** It is aerobic and the ethanol is separated by filtration.☐**B** It is aerobic and the ethanol is separated by fractional distillation.☐**C** It is anaerobic and the ethanol is separated by filtration.☐**D** It is anaerobic and the ethanol is separated by fractional distillation.☒

2 5

Which element is in the p block in the Periodic Table?

[1 mark]**A** Radium☐**B** Radon☒**C** Rhenium☐**D** Roentgenium☐**2 6**

Which statement is correct?

[1 mark]**A** A minimum of 0.5 mol of water is needed to react completely with 1 mol of Na_2O ☐**B** A solution of hydrochloric acid containing 1 mol of HCl reacts with 1 mol of MgO to form a solution with pH 7☐**C** A solution of hydrochloric acid containing 6 mol of HCl reacts completely with 1 mol of Al_2O_3 ☒**D** A minimum of 4 mol of water is needed to react completely with 1 mol of P_4O_{10} ☐**2 7**

Which species has a tetrahedral shape?

[1 mark]**A** SF_4 ☐**B** FeCl_4^- ☒**C** XeF_4 ☐**D** ICl_4^- ☐

Turn over ►



2	8
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Alkanes are obtained from crude oil by fractional distillation.

Which statement is correct?

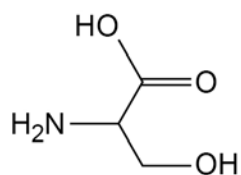
[1 mark]

- A** Alkanes obtained are converted to alkenes by thermal cracking.
- B** Alkanes with longer chains rise to the top of the fractionating column.
- C** The diesel fraction contains sulfur compounds that react to form SO_3 in car exhausts.
- D** The boiling points of alkane isomers increase as the carbon chains become more branched.

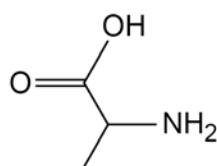
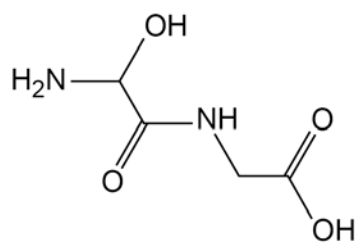
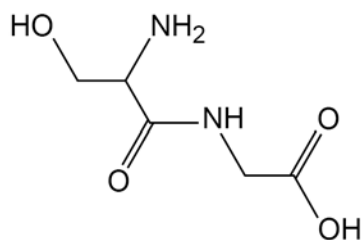
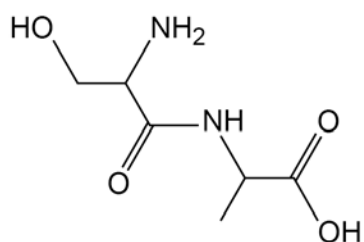
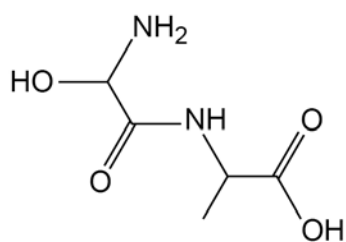
☒☐☐☐

2 9

What is the structure of a dipeptide formed from the two amino acids shown?



and

**[1 mark]****A****B****C****D****Turn over for the next question****Turn over ►**

3 0

Equal amounts of three Period 3 elements are added to separate 100 cm³ samples of water.

Which list is correct when the solutions formed from these elements are arranged in order of decreasing pH?

[1 mark]

- | | | | | |
|----------|-----------|-----------|-----------|----------------------------------|
| A | chlorine | magnesium | sodium | ☐ |
| B | chlorine | sodium | magnesium | ☐ |
| C | magnesium | sodium | chlorine | ☐ |
| D | sodium | magnesium | chlorine | ☒ |

3 1

The oxides P₄O₁₀ and SO₂ both react with water to form acid molecules.

Which row gives the correct shape around the central atom for each acid molecule that is formed?

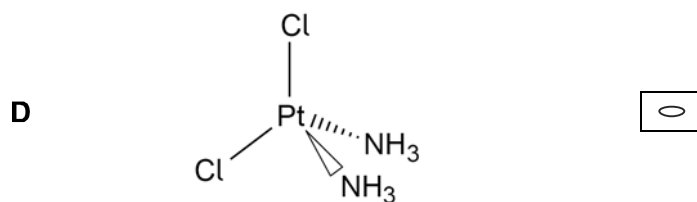
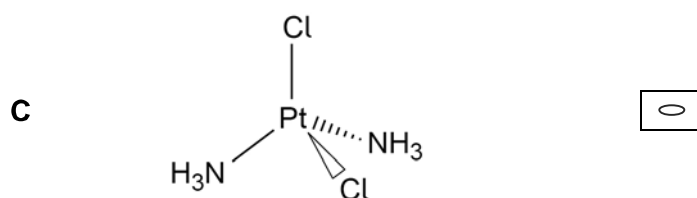
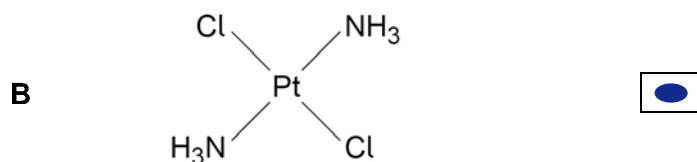
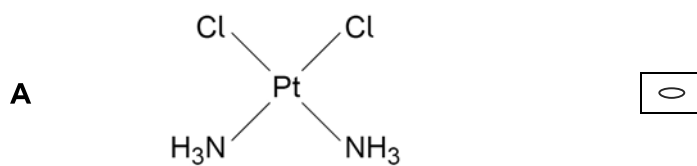
[1 mark]

	Molecule formed from P ₄ O ₁₀	Molecule formed from SO ₂	
A	pyramidal	pyramidal	☐
B	pyramidal	tetrahedral	☐
C	tetrahedral	pyramidal	☒
D	tetrahedral	tetrahedral	☐



3 2

Which structure is correct for transplatin?

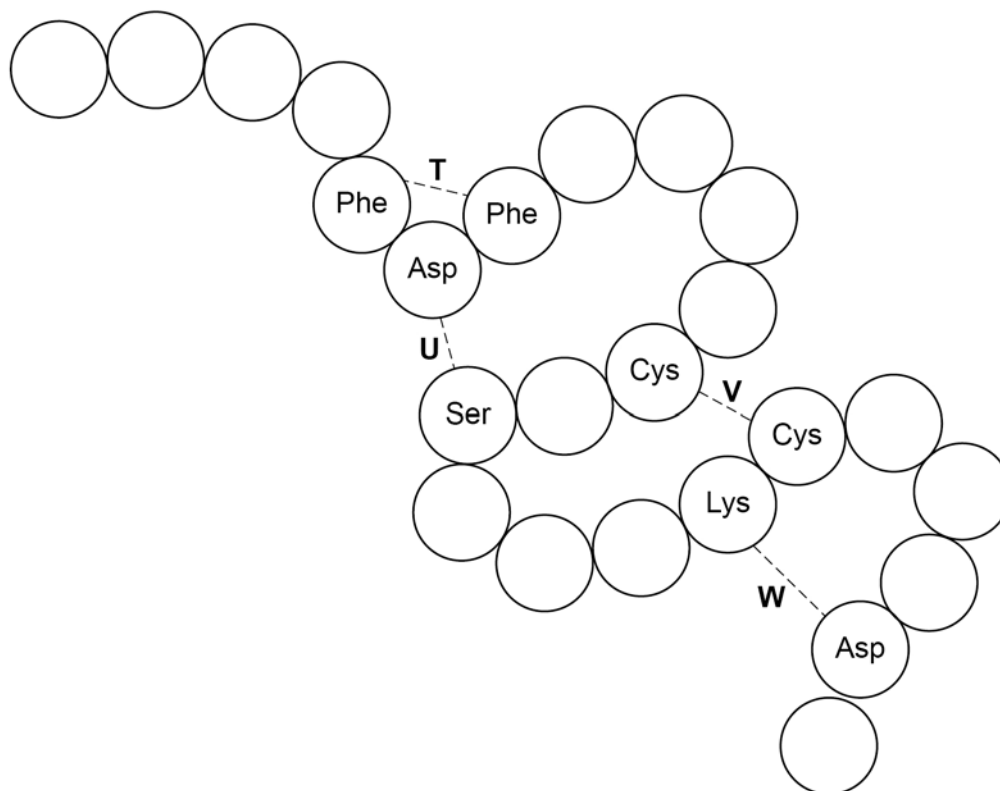
[1 mark]**Turn over for the next question****Turn over ►**

3 3

The diagram shows part of the tertiary structure for a protein.
Some of the amino acids in this structure are shown using the first three letters of their names.

The general formula of an amino acid is $\text{RCH}(\text{NH}_2)\text{COOH}$

Interactions between amino acid R groups are shown at positions **T**, **U**, **V** and **W**.



Which two R group interactions are a type of intermolecular force and a covalent bond?

Use the Data Booklet to help you answer this question.

[1 mark]

- | | | |
|----------|----------------|-------------------------------------|
| A | T and U | ☐ |
| B | V and W | ☐ |
| C | U and V | ☒ |
| D | U and W | ☐ |



3 4

Which statement about enzymes is correct?

[1 mark]

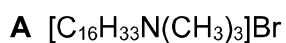
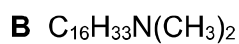
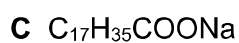
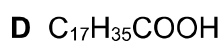
- A Enzymes are a type of nucleotide.
- B Enzymes are polymers.
- C Enzymes are **not** affected by changes of pH.
- D Enzymes act as catalysts by increasing activation energy.

☐☒☐☐

3 5

Which compound can be used as a cationic surfactant?

[1 mark]

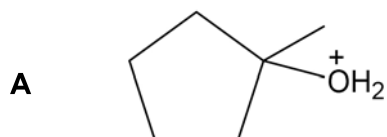
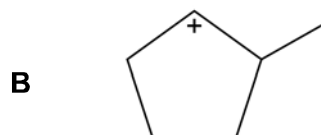
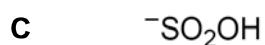
☒☐☐☐

3 6

The hydration of 1-methylcyclopent-1-ene uses a catalyst of sulfuric acid and produces a mixture of products.

Which ion is formed during the mechanism for this reaction to give the **major** product?

[1 mark]

☒☐☐☐

30

END OF QUESTIONS



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4 0



2 5 6 A 7 4 0 5 / 3