AQA A-LEVEL CHEMISTRY · 7405

# CHEMISTRY



EXAM QUESTION PACK

## 3.3.8 ALDEHYDES AND KETONES

| SCORE | PERCENTAGE |
| --- | --- |
| /323 | % |

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## Q1.

Acyl chlorides are useful reagents in synthesis. They react with aromatic compounds and also with alcohols.

(a) $CH_3CH_2COCl$ reacts with benzene in the presence of $AlCl_3$ in an electrophilic substitution reaction.

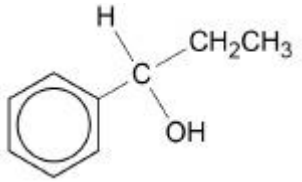
Give an equation for the reaction of $CH_3CH_2COCl$ with $AlCl_3$ to form the electrophile.
Outline a mechanism for the reaction of this electrophile with benzene.

Equation _____

Mechanism

**(4)**

(b) The organic product in **part (a)** can be converted into the alcohol shown.




Give the IUPAC name of the alcohol.
Give the reagent needed for this reaction and name the mechanism.

IUPAC name _____

Reagent _____

Name of mechanism _____

**(3)**

(c) The alcohol shown in **part (b)** reacts with ethanoyl chloride to form an ester.

Describe what would be observed when the alcohol reacts with ethanoyl chloride.
Name the mechanism for the reaction to form the ester.
Draw the structure of the ester.

Observation _____

_____

Name of mechanism _____



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Structure of ester

**(3)**
**(Total 10 marks)**

## Q2.

How many structural isomers with the molecular formula $C_5H_{10}O$ react with Tollens' reagent?

**A** 3 ⬭

**B** 4 ⬭

**C** 5 ⬭

**D** 6 ⬭

**(Total 1 mark)**

## Q3.

The aldehyde $CH_3CH_2CH_2CH_2CHO$ reacts with KCN followed by dilute acid to form a racemic mixture of the two stereoisomers of $CH_3CH_2CH_2CH_2CH(OH)CN$

(a) Give the IUPAC name of $CH_3CH_2CH_2CH_2CH(OH)CN$

_____

**(1)**

(b) Describe how you would distinguish between separate samples of the two stereoisomers of $CH_3CH_2CH_2CH_2CH(OH)CN$

_____

_____

_____

_____

_____

**(2)**

(c) Explain why the reaction produces a racemic mixture.

_____

_____

_____



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_____

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**(3)**

(d) An isomer of $CH_3CH_2CH_2CH_2CHO$ reacts with KCN followed by dilute acid to form a compound that does not show stereoisomerism.

Draw the structure of the compound formed and justify why it does not show stereoisomerism.

Structure

Justification

_____

_____

**(2)**

**(Total 8 marks)**

## Q4.

Which pair of compounds does **not** form a racemic mixture when the compounds react?

**A**  + HCl ⬭

**B**  + HCN ⬭

**C**  + HCl ⬭

**D**  + HCN ⬭

**(Total 1 mark)**

## Q5.

Which alcohol could **not** be produced by the reduction of an aldehyde or a ketone?

**A** 2,2-dimethylpropan-1-ol ⬭



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**B** 2-methylbutan-2-ol ○

**C** 3-methylbutan-2-ol ○

**D** pentan-3-ol ○

**(Total 1 mark)**

## Q6.

Which compound forms a racemic mixture when reacted with KCN followed by dilute acid?

**A** HCHO ○

**B** $CH_3CHO$ ○

**C** $CH_3COCH_3$ ○

**D** $(CH_3CH_2)_2CO$ ○

**(Total 1 mark)**

## Q7.

Butanone is reduced in a two-step reaction using $NaBH_4$ followed by dilute hydrochloric acid.

(a) Write an overall equation for the reduction of butanone using [H] to represent the reductant.

_____

**(1)**

(b) By considering the mechanism of the reaction, explain why the product has **no** effect on plane polarised light.

_____

_____

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**(6)**
**(Total 7 marks)**

## Q8.

Ethanol can be oxidised by acidified potassium dichromate(VI) to ethanoic acid in a two-step process.

ethanol ⟶ ethanal ⟶ ethanoic acid

(a) In order to ensure that the oxidation to ethanoic acid is complete, the reaction is carried out under reflux.

Describe what happens when a reaction mixture is refluxed and why it is necessary, in this case, for complete oxidation to ethanoic acid.

_____

_____

_____

_____

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_____

_____

_____

**(3)**

(b) Write a half-equation for the overall oxidation of ethanol into ethanoic acid.

_____

**(1)**

(c) The boiling points of the organic compounds in a reaction mixture are shown in the following table.

| **Compound** | ethanol | ethanal | ethanoic acid |
|---|---|---|---|
| **Boiling point / °C** | 78 | 21 | 118 |

Use these data to describe how you would obtain a sample of ethanal from a mixture of these three compounds. Include in your answer a description of the apparatus you would use and how you would minimise the loss of ethanal. Your description of the apparatus can be either a description in words or a labelled sketch.

_____

_____

_____



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_____

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**(5)**

(d) Use your knowledge of structure and bonding to explain why it is possible to separate ethanal in this way.

_____

_____

_____

_____

_____

**(2)**

(e) A student obtained a sample of a liquid using the apparatus in part **(c)**.

Describe how the student could use chemical tests to confirm that the liquid contained ethanal and did **not** contain ethanoic acid.

_____

_____

_____

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_____

_____

**(5)**

**(Total 16 marks)**

## Q9.

Which alcohol could **not** be produced by the reduction of an aldehyde or a ketone?

**A** 2-methylbutan-1-ol ⬭

**B** 2-methylbutan-2-ol ⬭



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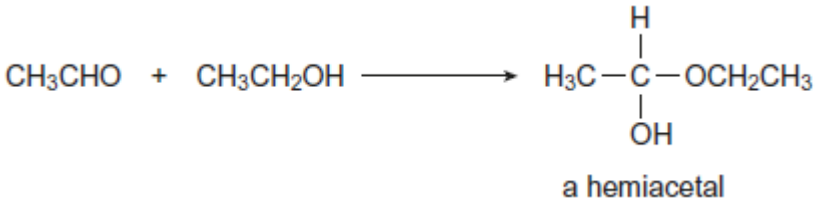
**C** 3-methylbutan-1-ol ⬭

**D** 3-methylbutan-2-ol ⬭

**(Total 1 mark)**

## Q10.

Hemiacetals and acetals are compounds formed by the reaction of aldehydes with alcohols, such as the reaction of ethanal with ethanol.




a hemiacetal

(a) (i) Use your knowledge of carbonyl mechanisms to suggest the name of the mechanism of this reaction.

_____

**(1)**

(ii) Outline how an ethanol molecule reacts with an ethanal molecule in the first step of this mechanism. Include two curly arrows to show the movement of electron pairs.

**(2)**

(b) The reaction produces a racemic mixture of chiral molecules.

(i) Explain the meaning of the term racemic mixture.

_____

_____

**(1)**

(ii) State the relationship between two chiral molecules with the same structural formula.

_____

_____

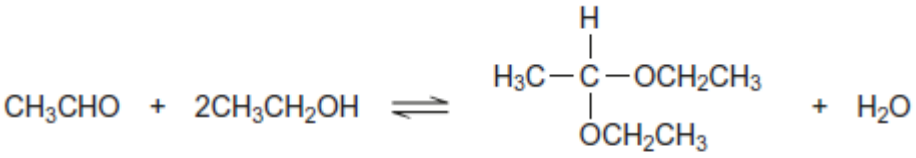
**(1)**

(c) In the presence of an acid catalyst such as dry hydrogen chloride, ethanal reacts with an excess of ethanol to form an acetal.



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The overall reaction of ethanal with an excess of ethanol forms an equilibrium mixture as shown. All reactants and products are liquids.



an acetal

A mixture of 0.75 mol of ethanal and 5.00 mol of ethanol was left to reach equilibrium in the presence of dry hydrogen chloride at a given temperature. The equilibrium mixture contained 0.42 mol of the acetal.

(i) Calculate the amount, in moles, of ethanal and of ethanol in this equilibrium mixture.

Amount of ethanal _____ mol

Amount of ethanol _____ mol

Space for working _____

_____

_____

_____

**(2)**

(ii) In a different experiment using the same reaction as in part **(c)**, an equilibrium mixture was established at a given temperature.
This mixture contained 0.58 mol of ethanal, 3.76 mol of ethanol, 0.37 mol of the acetal and 0.65 mol of water in a total volume of 310 $cm^3$.

Write an expression for the equilibrium constant $K_c$ for this reaction.
Calculate a value for $K_c$ at this temperature. Give units with your answer.

$K_c$ _____

_____

Calculation _____

_____

_____

_____

_____

_____

_____

**(4)**



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(d) Draw the structure of the acetal ($C_4H_8O_2$) formed by the reaction of ethanal with ethane-1,2-diol.

**(1)**

**(Total 12 marks)**

## Q11.

The carbonyl compound $CH_3CH_2CHO$ reacts very slowly with HCN

(a) Name and outline a mechanism for the reaction of $CH_3CH_2CHO$ with HCN

Name of mechanism _____

Mechanism

**(5)**

(b) The reaction in part (a) produces a pair of enantiomers.

(i) Draw the structure of each enantiomer to show how they are related to each other.

**(2)**

(ii) State and explain how you could distinguish between the two enantiomers.

_____

_____

_____

_____

**(2)**



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(c) Give the IUPAC name of the product of the reaction in part (a).

_____

**(1)**

(d) In practice, KCN rather than HCN is added to the carbonyl compound.

Given that $K_a$ for HCN = $4.0 \times 10^{-10}$ mol $dm^{-3}$, suggest why the reaction with HCN is very slow.

_____

_____

_____

_____

**(2)**

(e) Acrylic fibres are used as a substitute for wool. Acrylics are copolymers of acrylonitrile with other compounds.

Acrylonitrile is the common name for the following compound.

$$H_2C = CH - C \equiv N$$

(i) Acrylonitrile can be formed from propene.

Write an equation for the reaction of propene with ammonia and oxygen to form acrylonitrile and one other product.

_____

**(1)**

(ii) The term copolymer is used to describe the product obtained when two or more different monomers form a polymer.

Draw the repeating unit of the acrylic copolymer that contains 75% acrylonitrile monomer and 25% chloroethene monomer.

**(1)**

(iii) Name the type of polymerisation involved in part (ii)

_____

**(1)**

**(Total 15 marks)**

## Q12.

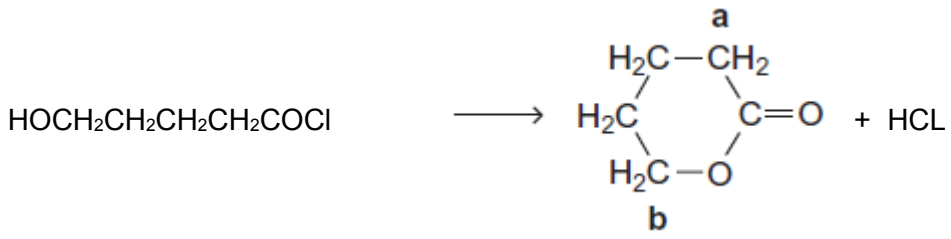


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This question is about some isomers of $C_5H_8O_2$

(a) Compound **H** is a cyclic ester that can be prepared as shown.

On the structure of **H**, two of the carbon atoms are labelled.




**H**

(i) Name and outline a mechanism for this reaction.

Use **Table C** on the Data Sheet to give the $^{13}C$ n.m.r. δ value for the carbon atom labelled **a** and the δ value for the carbon atom labelled **b**.

**(7)**

(ii) $HOCH_2CH_2CH_2CH_2COCl$ can also react to form a polyester in a mechanism similar to that in part (i).

Draw the repeating unit of the polyester and name the type of polymerisation



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involved.

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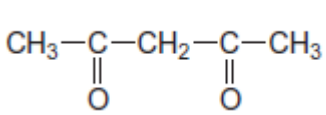
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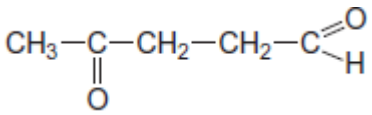
**(2)**

(b) State how you could distinguish between compounds **J** and **K** by a simple test-tube reaction.

State how you could distinguish between **J** and **K** by giving the number of peaks in the $^1$H n.m.r. spectrum of each compound.



**J**



**K**

_____

_____

_____

_____

_____

_____

_____

_____

**(5)**

(c) Draw the structure of each of the following isomers of $C_5H_8O_2$
Label each structure you draw with the correct letter **L**, **M**, **N**, **P** or **Q**.

**L** is methyl 2-methylpropenoate.

**M** is an ester that shows E-Z stereoisomerism.

**N** is a carboxylic acid with a branched carbon chain and does **not** show stereoisomerism.

**P** is an optically active carboxylic acid.

**Q** is a cyclic compound that contains a ketone group and has only two peaks in its $^1$H n.m.r. spectrum.



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**(5)**

**(Total 19 marks)**

## Q13.

Lactic acid, $CH_3CH(OH)COOH$, is formed in the human body during metabolism and exercise. This acid is also formed by the fermentation of carbohydrates such as sucrose, $C_{12}H_{22}O_{11}$.

(a) (i) Give the IUPAC name for lactic acid.

_____

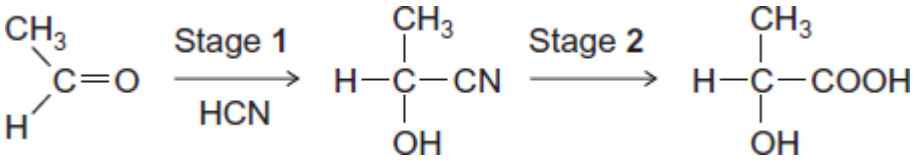
**(1)**

(ii) Write an equation for the formation of lactic acid from sucrose and water.

_____

**(1)**

(b) A molecule of lactic acid contains an asymmetric carbon atom.
The lactic acid in the body occurs as a single enantiomer.
A racemic mixture (racemate) of lactic acid can be formed in the following two-stage synthesis.




(i) Name and outline a mechanism for Stage **1**.

Name of mechanism _____

Mechanism

**(5)**

(ii) Give the meaning of the term *racemic mixture (racemate)*.

_____



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_____

_____

**(1)**

(iii) Explain how you could distinguish between a racemic mixture (racemate) of lactic acid and one of the enantiomers of lactic acid.

_____

_____

_____

_____

**(2)**

(c) A mixture of lactic acid and its salt sodium lactate is used as an acidity regulator in some foods. An acidity regulator makes sure that there is little variation in the pH of food.

(i) Write an equation for the reaction of lactic acid with sodium hydroxide.

_____

**(1)**

(ii) The acid dissociation constant $K_a$ for lactic acid has the value $1.38 \times 10^{-4}$ mol $dm^{-3}$ at 298 K.

Calculate the pH of an equimolar solution of lactic acid and sodium lactate.

_____

_____

_____

_____

**(2)**

(iii) Suggest an alternative name for the term *acidity regulator*.
Explain how a mixture of lactic acid and sodium lactate can act as a regulator when natural processes increase the acidity in some foods.

Name _____

Explanation _____

_____

_____

_____

**(3)**

(d)



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The cup shown is made from PLA, poly(lactic acid).
PLA is the condensation polymer formed from lactic acid.

The polymer is described as 100% biodegradable and 100% compostable.

Compostable material breaks down slowly in contact with the moist air in a garden bin. This produces compost that can be used to improve soil.

The manufacturers stress that PLA cups differ from traditional plastic cups that are neither biodegradable nor compostable.

(i) Draw a section of PLA that shows **two** repeating units.

**(2)**

(ii) Name the type of condensation polymer in PLA.

_____

**(1)**

(iii) An intermediate in the production of PLA is a cyclic compound ($C_6H_8O_4$) that is formed from two PLA molecules.

Draw the structure of this cyclic compound.



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**(1)**

(iv) Traditional non-biodegradable plastic cups can be made from poly(phenylethene), commonly known as *polystyrene*.

Draw the repeating unit of poly(phenylethene).

**(1)**

(v) The manufacturers of PLA claim that the material will break down to compost in just 12 weeks.

Suggest **one** reason why PLA in landfill may take longer than 12 weeks to break down.

_____

_____

**(1)**

**(Total 22 marks)**

## Q14.

Suggest **one** reason why Tollens' reagent is used as the oxidising agent in the specific test for aldehydes rather than the less expensive acidified potassium dichromate(VI).

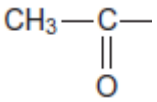
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**(Total 1 mark)**

## Q15.

The triiodomethane reaction is often used as a test for aldehydes and ketones that contain the $CH_3CO$ group shown.



The aldehyde or ketone is reacted with an alkaline solution of iodine. Triiodomethane ($CHI_3$) is formed as a precipitate. Compounds that contain a group that can be oxidised to the $CH_3CO$ group will also give a positive result in this test.

(a) State, with a reason, whether or not ethanol will give a positive result in the



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triiodomethane reaction.

_____

_____

_____

**(1)**

(b) The equation for the reaction of ethanal with an alkaline solution of iodine is

$CH_3CHO + 3I_2 + 4NaOH \longrightarrow CHI_3 + HCOONa + 3NaI + 3H_2O$

In an experiment using this reaction, the yield of triiodomethane ($CHI_3$) obtained by a student was 83.2%.

Calculate the minimum mass of iodine that this student would have used to form 10.0 g of triiodomethane.
Give your answer to the appropriate precision.
Show your working.

_____

_____

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**(5)**

(c) Triiodomethane can be separated from the reaction mixture by filtration.
State **one** reason why the solid residue is then washed with water after the filtration.

_____

_____

**(1)**

(d) State **one** reason, other than cost or availability, why water is suitable for washing this solid residue after the filtration.

_____

_____

**(1)**



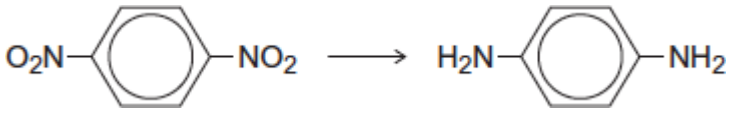
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**(Total 8 marks)**

## Q16.

Each of the following conversions involves reduction of the starting material.

(a) Consider the following conversion.



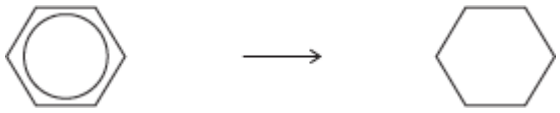

Identify a reducing agent for this conversion.

Write a balanced equation for the reaction using molecular formulae for the nitrogen-containing compounds and [H] for the reducing agent.

Draw the repeating unit of the polymer formed by the product of this reaction with benzene-1,4-dicarboxylic acid.

**(5)**

(b) Consider the following conversion.



Identify a reducing agent for this conversion.

State the empirical formula of the product.

State the bond angle between the carbon atoms in the starting material and the bond angle between the carbon atoms in the product.

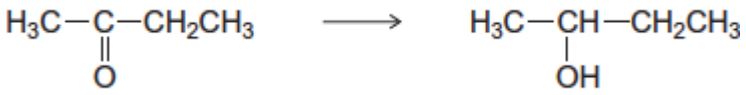


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_____

**(4)**

(c) The reducing agent in the following conversion is $NaBH_4$



(i) Name and outline a mechanism for the reaction.

Name of mechanism _____

Mechanism

**(5)**

(ii) By considering the mechanism of this reaction, explain why the product formed is optically inactive.

_____

_____

_____

_____

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**(3)**

**(Total 17 marks)**

## Q17.

Chemists have to design synthetic routes to convert one organic compound into another.

Propanone can be converted into 2-bromopropane by a three-step synthesis.

Step 1: propanone is reduced to compound **L**.
Step 2: compound **L** is converted into compound **M**.
Step 3: compound **M** reacts to form 2-bromopropane.



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Deduce the structure of compounds **L** and **M**.

For each of the three steps, suggest a reagent that could be used and name the mechanism.

Equations and curly arrow mechanisms are **not** required.

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_____

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**(Total 8 marks)**

## Q18.

(a) Propanoic acid can be made from propan-1-ol by oxidation using acidified potassium dichromate(VI). Propanal is formed as an intermediate during this oxidation.

(i) State the colour of the chromium species after the potassium dichromate(VI) has reacted.

_____

**(1)**

(ii) Describe the experimental conditions and the practical method used to ensure that the acid is obtained in a high yield. Draw a diagram of the assembled apparatus you would use.

Conditions _____

_____

Apparatus

**(4)**

(iii) Describe the different experimental conditions necessary to produce propanal



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in high yield rather than propanoic acid.

_____

_____

**(2)**

(b) Propan-1-ol is a volatile, flammable liquid.
Give **one** safety precaution that should be used during the reaction to minimise this hazard.

_____

**(1)**

(c) A student followed the progress of the oxidation of propan-1-ol to propanoic acid by extracting the organic compounds from one sample of reaction mixture.

(i) Give a chemical reagent which would enable the student to confirm the presence of propanal in the extracted compounds.
State what you would observe when propanal reacts with this reagent.

Reagent _____

Observation _____

_____

**(2)**

(ii) Give a chemical reagent that would enable the student to confirm the presence of propanoic acid in the extracted compounds.
State what you would observe when propanoic acid reacts with this reagent.

Reagent _____

Observation _____

_____

**(2)**

(d) Predict which **one** of the compounds, propan-1-ol, propanal and propanoic acid will have the highest boiling point. Explain your answer.

Prediction _____

Explanation _____

_____

_____

_____

_____

**(3)**

**(Total 15 marks)**



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## Q19.

Many synthetic routes need chemists to increase the number of carbon atoms in a molecule by forming new carbon–carbon bonds. This can be achieved in several ways including

- reaction of an aromatic compound with an acyl chloride
- reaction of an aldehyde with hydrogen cyanide.

(a) Consider the reaction of benzene with $CH_3CH_2COCl$

(i) Write an equation for this reaction and name the organic product.
Identify the catalyst required in this reaction.
Write equations to show how the catalyst is used to form a reactive intermediate and how the catalyst is reformed at the end of the reaction.

**(5)**

(ii) Name and outline a mechanism for the reaction of benzene with this reactive intermediate.

**(4)**

(b) Consider the reaction of propanal with HCN

(i) Write an equation for the reaction of propanal with HCN and name the product.



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**(2)**

(ii) Name and outline a mechanism for the reaction of propanal with HCN

**(5)**

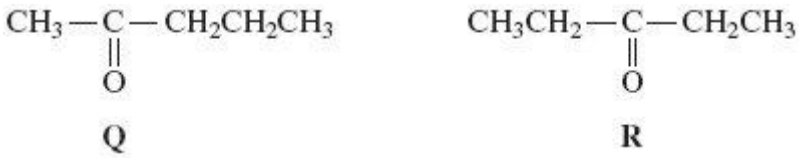
(iii) The rate-determining step in the mechanism in part (b) (ii) involves attack by the nucleophile.
Suggest how the rate of reaction of propanone with HCN would compare with the rate of reaction of propanal with HCN
Explain your answer.

**(2)**

**(Total 18 marks)**

## Q20.

Two isomeric ketones are shown below.



**Q** **R**



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(a) Name and outline a mechanism for the reaction of compound **Q** with HCN and name the product formed.

Name of mechanism _____

Mechanism

Name of product _____

**(6)**

(b) Some students were asked to suggest methods to distinguish between isomers **Q** and **R**.

One student suggested testing the optical activity of the products formed when **Q** and **R** were reacted separately with HCN.

By considering the optical activity of these products formed from **Q** and **R**, explain why this method would **not** distinguish between **Q** and **R**.

_____

_____

_____

_____

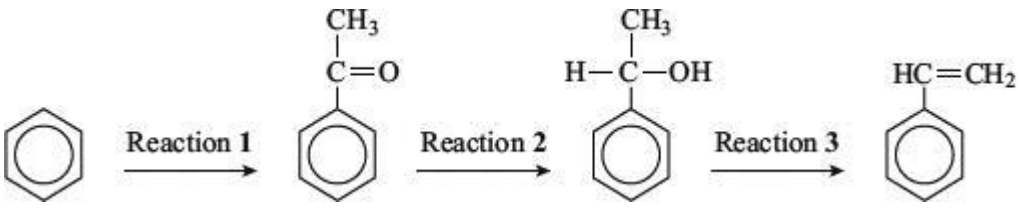
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**(3)**

**(Total 9 marks)**

## Q21.

A possible synthesis of phenylethene *(styrene)* is outlined below.




(a) In Reaction **1**, ethanoyl chloride and aluminium chloride are used to form a reactive species which then reacts with benzene.
Write an equation to show the formation of the reactive species.
Name and outline the mechanism by which this reactive species reacts with benzene.



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**(6)**

(b) $NaBH_4$ is a possible reagent for Reaction **2**.
Name and outline the mechanism for the reaction with $NaBH_4$ in Reaction **2**.
Name the product of Reaction **2**.

**(6)**

(c) Name the type of reaction involved in Reaction **3** and give a reagent for the reaction.



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_____

_____

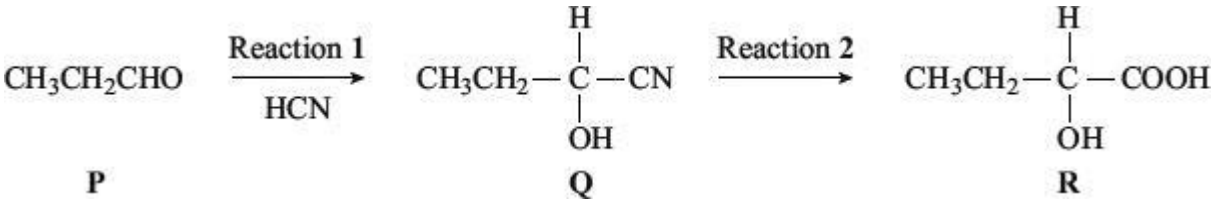
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**(2)**

**(Total 14 marks)**

## Q22.

Consider the sequence of reactions below.




(a) Name and outline a mechanism for Reaction **1**.

*Name of mechanism* _____

*Mechanism*

**(5)**

(b) (i) Name compound **Q**

_____

(ii) The molecular formula of **Q** is $C_4H_7NO$. Draw the structure of the isomer of **Q** which shows geometrical isomerism and is formed by the reaction of ammonia with an acyl chloride.

**(3)**

(c) Draw the structure of the main organic product formed in each case when **R** reacts separately with the following substances:

(i) methanol in the presence of a few drops of concentrated sulphuric acid;



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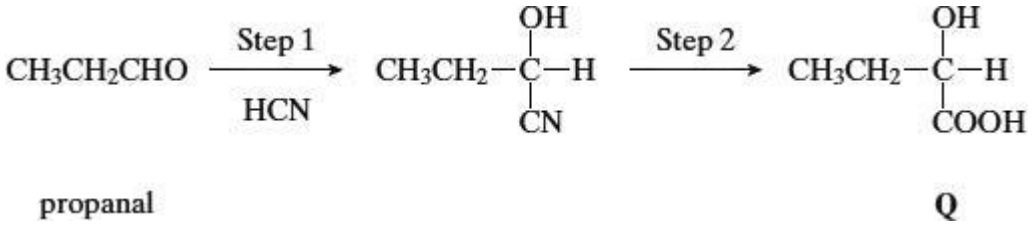
(ii) acidified potassium dichromate(VI);

(iii) concentrated sulphuric acid in an elimination reaction.

**(3)**
**(Total 11 marks)**

## Q23.

Consider the reaction sequence shown below.




propanal Q

(a) Name and outline a mechanism for the reaction in Step 1.

*Name of mechanism* _____

Mechanism



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**(5)**

(b) (i) Name compound **Q** formed in Step 2.

_____

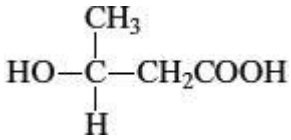
(ii) Two stereoisomers are formed by the dehydration of **Q**. Give the structures of these two isomers and name the type of stereoisomerism shown.

*Structures of isomers*

*Type of stereoisomerism* _____

**(4)**

(c) An isomer of **Q** which has the structure shown below is polymerised to form the biodegradeable polymer known as PHB.



(i) Draw the repeating unit of the polymer PHB.

(ii) Suggest a reason why the polymer is biodegradeable.

_____

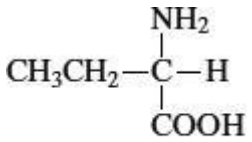
_____

**(2)**



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(d) The amino acid **R** is shown below.



**R**

(i) Draw the structure of the zwitterion formed by **R**.

(ii) Draw the structure of the major organic product formed when an excess of **R** is reacted with bromomethane.

(iii) Name the mechanism of the reaction which results in the formation of the product given in part (ii).

_____

**(3)**

**(Total 14 marks)**

## Q24.

Which one of the following reactions will produce an organic compound that has optical isomers?

**A** dehydration of butan-2-ol by heating with concentrated sulphuric acid

**B** reduction of pentan-3-one by warming with $NaBH_4$

**C** addition of $Br_2$ to 3-bromopropene

**D** reduction of 2,3-dimethylpent-2-ene with $H_2$ in the presence of a nickel catalyst

**(Total 1 mark)**

## Q25.

(a) Describe how propanal, $CH_3CH_2CHO$, and propanone, $CH_3COCH_3$, can be



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distinguished using

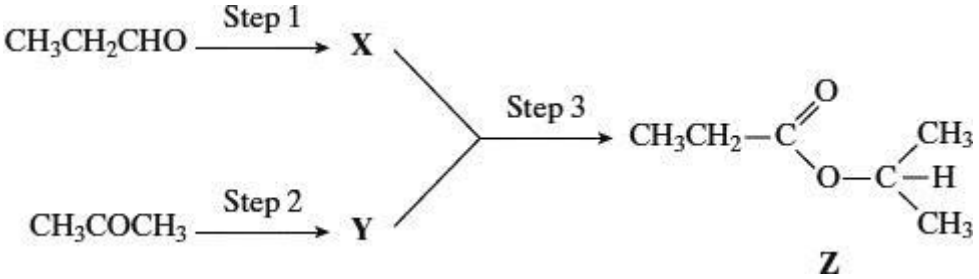
(i) a chemical test and

**(3)**

(ii) the number of peaks in their proton n.m.r. spectra.

**(2)**

(b) Compound **Z** can be produced by the reaction of compound **X** with compound **Y** as shown in the synthesis outlined below.




**Z**

Identify compounds **X** and **Y**.
For each of the three steps in the synthesis, name the type of reaction involved and give reagents and conditions. Equations are **not** required.



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**(10)**

**(Total 15 marks)**

## Q26.

(a) Addition reactions to both alkenes and carbonyl compounds can result in the formation of isomeric compounds.

(i) Choose an alkene with molecular formula $C_4H_8$ which reacts with HBr to form two structural isomers. Give the structures of these two isomers and name the type of structural isomerism shown.

Outline a mechanism for the formation of the major product.



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**(7)**

(ii) Using HCN and a suitable carbonyl compound with molecular formula $C_3H_6O$, outline a mechanism for an addition reaction in which two isomers are produced.
Give the structures of the two isomers formed and state the type of isomerism shown.

**(7)**

(b) Explain why ethanoyl chloride reacts readily with nucleophiles.
Write an equation for one nucleophilic addition–elimination reaction of ethanoyl



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chloride.
(A mechanism is not required.)

_____

_____

_____

_____

_____

_____

_____

_____

**(4)**

**(Total 18 marks)**

## Q27.

In which one of the following mixtures does a redox reaction occur?

**A** ethanal and Tollens' reagent

**B** ethanoyl chloride and ethanol

**C** ethanal and hydrogen cyanide

**D** ethanoic acid and sodium hydroxide

**(Total 1 mark)**

## Q28.

Propanone can be reduced to form an alcohol. A functional group isomer of the alcohol formed is

**A** $CH_3CH_2CH_2OH$

**B** $CH_3CH_2CHO$

**C** $CH_3OCH_2CH_3$

**D** $CH_3COCH_3$

**(Total 1 mark)**

## Q29.

(a) (i) Give a suitable reagent and state the necessary conditions for the conversion of propan-2-ol into propanone. Name the type of reaction.

*Reagent* _____

*Conditions* _____



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*Type of reaction* _____

(ii) Propanone can be converted back into propan-2-ol. Give a suitable reagent and write an equation for this reaction.
(Use [H] to represent the reagent in your equation.)

*Reagent* _____

*Equation*

_____

**(5)**

(b) Propanal is an isomer of propanone.

(i) Draw the structure of propanal.

(ii) A chemical test can be used to distinguish between separate samples of propanone and propanal. Give a suitable reagent for the test and describe what you would observe with propanone and with propanal.

*Test reagent* _____

*Observation with propanone* _____

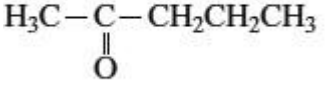
*Observation with propanone* _____

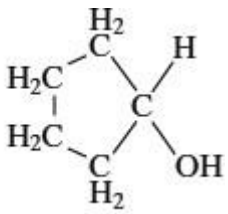
**(4)**

**(Total 9 marks)**

## Q30.

Compounds **C** and **D**, shown below, are isomers of $C_5H_{10}O$







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**C** **D**

(a) Name compound **C**.

_____

**(1)**

(b) Use **Table 2** on the Data Sheet to help you to answer this question.

(i) Suggest the wavenumber of an absorption which is present in the infra-red spectrum of **C** but not in that of **D**.

_____

(ii) Suggest the wavenumber of an absorption which is present in the infra-red spectrum of **D** but not in that of **C**.

_____

**(2)**

(c) Deduce the number of peaks in the proton n.m.r. spectrum of **C**.

_____

**(1)**

(d) Identify a reagent that you could use to distinguish between **C** and **D**. For each of **C** and **D**, state what you would observe when the compound is treated with this reagent.

*Reagent* _____

*Observation with* **C** _____

*Observation with* **D** _____

**(3)**

(e) Compound **E**, $CH_3CH_2CH_2CH_2CHO$, is also an isomer of $C_5H_{10}O$

Identify a reagent which will react with **E** but not with **C** or **D**. State what you would observe when **E** is treated with this reagent.

*Reagent* _____

*Observation with* **E** _____

**(2)**

**(Total 9 marks)**

## Q31.

Which one of the following is **not** a correct general formula for the non-cyclic compounds listed?

**A** alcohols $C_nH_{2n+2}O$

**B** aldehydes $C_nH_{2n+1}O$



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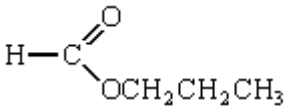
**C** esters $C_nH_{2n}O_2$

**D** primary amines $C_nH_{2n+3}N$

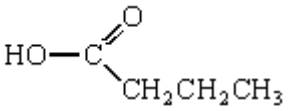
**(Total 1 mark)**

## Q32.

(a) Consider the following pair of isomers.



**C**



**D**

(i) Name compound **C**.

_____

(ii) Identify a reagent which could be used in a test-tube reaction to distinguish between **C** and **D**. In each case, state what you would observe.

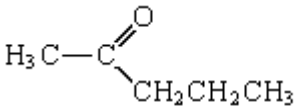
*Reagent* _____

*Observation with **C*** _____

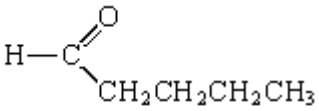
*Observation with **D***_____

**(4)**

(b) Consider the following pair of isomers.



**E**



**F**

(i) Name compound **E**.

_____

(ii) Identify a reagent which could be used in a test-tube reaction to distinguish between **E** and **F**. In each case, state what you would observe.

*Reagent* _____

*Observation with **E*** _____

*Observation with **F***_____

**(4)**

(c) Draw the structure of the chain isomer of **F** which shows optical isomerism.



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**(1)**

**(Total 9 marks)**

## Q33.

The three compounds $CH_3CH_2CH_2CH_2OH$, $(CH_3)_3COH$ and $CH_3CH_2CH_2CHO$ can be distinguished by use of the following three reagents

1. potassium dichromate(VI) acidified with dilute sulphuric acid
2. Tollens' reagent
3. ethanoic acid, together with a small amount of concentrated sulphuric acid.

(a) Identify which of these three organic compounds would reduce acidified potassium dichromate(VI). Give the structures of the organic products formed. Write a half-equation for the reduction of dichromate(VI) ions in acidic solution.

**(6)**

(b) Identify which one of these three organic compounds would reduce Tollens' reagent. Give the structure of the organic product formed. Write a half-equation for the reduction of Tollens' reagent.

**(3)**



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(c) Identify which of these three organic compounds would react with ethanoic acid in the presence of concentrated sulphuric acid. In each case, give the structure of the organic product formed.

_____

_____

_____

_____

_____

_____

_____

_____

**(4)**

(d) State the number of peaks in the proton n.m.r. spectra of $CH_3CH_2CH_2CH_2OH$ and of $(CH_3)_3COH$. (Analysis of peak splitting is not required.)

_____

_____

_____

_____

**(2)**

**(Total 15 marks)**

## Q34.

(a) Outline a mechanism for the reaction of $CH_3CH_2CH_2CHO$ with HCN and name the product.

*Mechanism*

*Name of product* _____

**(5)**

(b) Outline a mechanism for the reaction of $CH_3OH$ with $CH_3CH_2COCl$ and name the organic product.



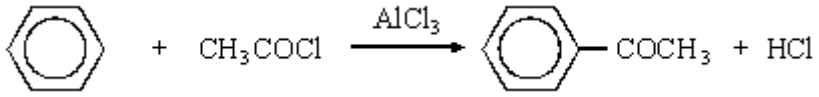
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*Mechanism*

*Name of organic product* _____

**(5)**

(c) An equation for the formation of phenylethanone is shown below. In this reaction a reactive intermediate is formed from ethanoyl chloride. This intermediate then reacts with benzene.




(i) Give the formula of the reactive intermediate.

_____

(ii) Outline a mechanism for the reaction of this intermediate with benzene to form phenylethanone.

**(4)**
**(Total 14 marks)**

## Q35.

Which one of the following would **not** reduce an acidified aqueous solution of potassium dichromate(VI)?

**A** $CH_3COOH$

**B** Zn

**C** $CH_3CHO$



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**D** $Fe^{2+}(aq)$

**(Total 1 mark)**



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